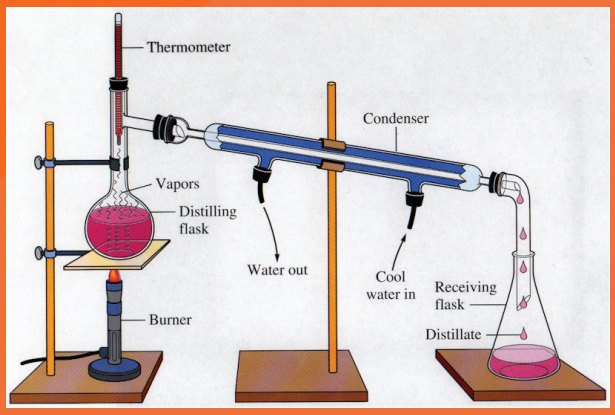
General Science

Student Textbook



Grade 7

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General Science

Student Textbook

Grade 7

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Acknowledgement

Above all, Ato Zelalem Mulatu, AAEB Head, should receive the most acknowledgements for his outstanding leadership from the outset to the end of the Textbook and Teacher's guide preparation. Just to mention his chief roles but a few: he generated valuable ideas, shared his vast experience during most panels, initiated and convinced his conviction to all stakeholders that the Addis Ababa City Government School teachers have to take the lion's share in the Textbook and Teacher Guide development. His unabated inspiration and urge to the team for diligence, deep sense of patriotism, synergy and true professional ethics has been energy to all actors partaking in the task.

The next bottom-heart gratitude has to be extended to the following management members of the bureau: Ato Admasu Dechasa, Deputy Head of the Curriculum Division, Ato Dagnew Gebru, Deputy Head of the Education Technology, Ato Samson Melese, deputy Head of Teacher Development Division, W/ro Abebech Negash, Bureau Head Advisor, Ato Desta Mersha, Bureau Technical Advisor and Ato Sisay Endale, Head of Education Bureau Office. Members of the AAEB management, have to be commended for their painstaking efforts in addressing instantly each issue of challenge, reviewing all drafts and providing immediate feedbacks. Without their unreserved devotion, the timely and successful realization of this huge work would not have been possible.

The Last deepest acknowledgement needs to go to the school principals for allowing the textbook writers to be free of their regular job of teaching and to focus on the material preparation. Moreover, their earnest and consistent moral support deserves special words of praise.

Unit 1 BASIC CONCEPTS OF SCIENCE

Learning Outcomes

At the end of this unit, you will be able to:

- define science as a body of knowledge and the processes and practices used to add to that body of knowledge;
- describe the main branches of science and explain their relationship
- relate how science and technology affect one's beliefs, practices and ways of thinking;
- appreciate the contributions of outstanding scientists to science and technology;
- discuss the importance of scientific values in decision making and problem solving;
- identify the significant contributions of Ethiopian Scientists in science and technology;
- identify different laboratory tools;
- demonstrate safe ways of using apparatus in the laboratory;
- practice precautionary measures in the laboratory;
- exhibit knowledge of laboratory safety rules and procedures;
- identify potential hazards and implement appropriate safety procedures when working in the laboratory.

Main Contents

1.1. The Nature of Science and its Branches
- Definition of science
- Branches of science
- Science and technology
- Scientists and ethical discipline

1.2. Common Laboratory Equipment, Uses, Safety Rules and Procedures in Science Laboratories
- Common Laboratory Apparatus
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- Steps to write Laboratory report

1.1 The Nature of Science and its Branches

By the end of this section you should be able to:

- Define science;
- Distinguish between the Indigenous Science and Conventional Science;
- Describe the main branches of science and explain their relationship;
- Relate how science and technology affect one's beliefs, practices, and ways of thinking;
- Appreciate the contributions of outstanding Ethiopian scientists to science and technology;
- Discuss the importance of ethical disciplines in scientific investigations;
- Solve the issue of environmental problems in their school compound and its surroundings.

Introduction

In the lower grades, you have learnt about science in general. For example, environmental science. In this and next grade you will learn about general science which deals with things related to your day to day life.

Activity 1.1

Form a group and discuss the following questions. Then share your ideas to the classmates

i. Describe science by your own words
ii. Investigate the ways in which the major areas of science are further divided. You can use reference books and the internet to augment your current ideas
iii. Differentiate Conventional Science and Indigenous Science

The word science comes from the Latin word ***'Scientia'***, which means ***'Knowledge'.*** But science is not just about having knowledge: Science is a systematic method of gaining knowledge about the physical and natural world and the social aspect of human society. It provides an ordered way of learning about the nature of things, based on observation and evidence. Science can be indigenous or conventional.

Indigenous science is process by which indigenous people build their empirical knowledge of their natural environment. It is knowledge based on the social, physical and spiritual understandings.

Conventional science is the system of knowledge which relies on certain laws that have been established through the application of the scientific method to phenomena in the world around us.

Indigenous Science incorporating local people's knowledge and indigenous perspectives, while conventional scientific approaches are commonly recognized as Western science.

Activity 1.2
Perform the following activities.
Find some practical indigenous knowledge in your community that solves community problems and present your finding to your class

Why Do You Learn Indigenous Knowledge In Science?

There are two main reasons to include Indigenous Knowledge in the science: firstly, to increase awareness of original culture and identity and secondly, to integrate indigenous knowledge with western science. Ethiopia is one of the countries where a wide variety indigenous knowledge practiced for a long time to solve practical problem that exist in different areas like:

- extractions of medicinal chemicals from plants to treat disease and fight infections. The common medicinal plants used for treating and curing various disease are: Hagenia Abyssinica (Kosso tree) , Eucalyptus globulus (bahrzaf), and Ocimum lamiifolium Hochst (Damakese)etc
- preserving meat by adding a salt and smoke drying.

1.1.2. Branches of Science

Activity 1.3
Form a group and discuss the following questions. Then share your ideas to the class
1.What are the major branches of science?
2. Give short descriptions of physics, Chemistry and Biology.

Science has two major categories, which are natural science and social science.

Natural science is the study of nature and natural laws. It includes fields such as Chemistry, Biology and Physics.

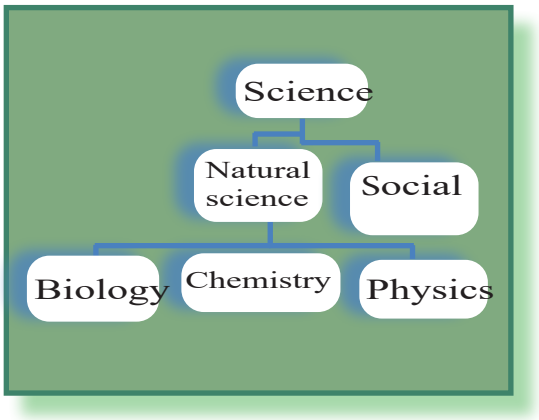


Figure 1.1. Branches of science

- Biology is a branch of natural science which studies about living things.
- Chemistry is a branch of natural science which deals with the properties, composition, structure and transformation of substances.
- Physics is the branch of natural science. It is the study of the nature of matter, energy and their interactions.

There is no clear boarder line between the different branches of natural sciences. Knowledge of natural sciences overlaps with each other. For example, Chemistry and Physics knowledge are studied as a subject called physical science/physical chemistry. It is the study of properties of materials and their interaction.

◊ **Biopysics:** a combination of Biology and Physics. It is the study of physical phenomena and physical processes in living things, on scales spanning molecules, cells, tissues and organisms.

◊ **Biochemistry:** combination of biology and Chemistry. It is the branch of science that explores the chemical processes within and related to living organisms. It involves the study of chemical reaction in living things.

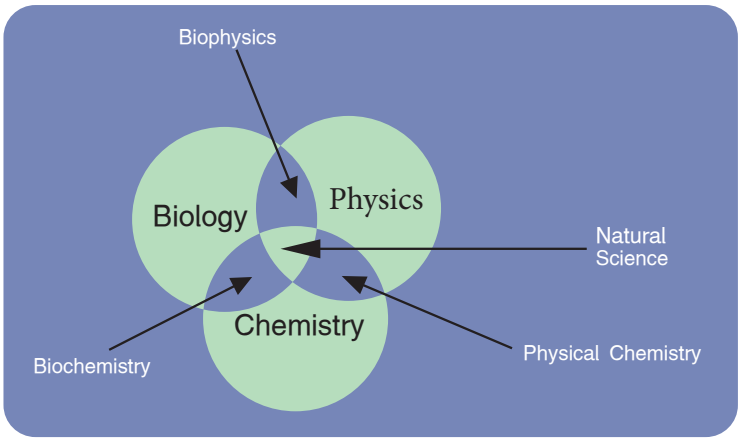


Figure1.2.The relationships between some fields of Natural Science

Exercise 1.1

I. Give short answer for the followings questions

1. What is science?
2. List the three branches of natural science
3. Which field of science studies about matter and energy?
4. What is the difference betwen biophysics and biochemistry?

1.1.3. Science and Technology

You have already discussed what science is. Now, you will see what a technology is.

Technology is the use of scientific knowledge to help human beings work easier and live better as well as enjoy their environment more. It includes the use of materials, tools, techniques, and sources of power to make life easier, more pleasant and work more productive.

Things such as automobiles, TV sets, radio, bulb, microchip, computer, airplane and home tools (appliances) are the products of technology.

Key words

Science is a systematic method of gaining knowledge about the physical and natural world.

Technology is the use of scientific knowledge to help human bings work easier and live better or putting scientific knowledge intopractice.

A person who studies technology is called a technologist.
Technologists apply Science and mathematical knowledge and skills to produce a very useful tool.

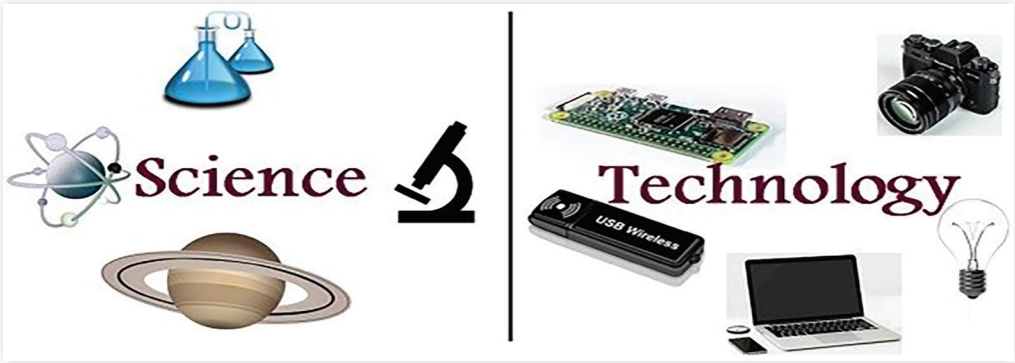


Figure 1.3 Relationship of science and technology

Science and technology is integrally connected; but they are different fields driven by different concepts and processes. Science generates knowledge for its own sake, in order to propose and test explanations. Technology, on the other hand, develops human-made solutions to real-world problems. Of course, when science uses technology to generate knowledge, technology uses scientific knowledge to generate solutions.

Uses of Science and Technology for the Society

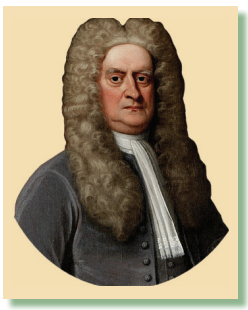
Science and technology plays an important role in our daily life. They mainly concerned with the production of new materials of desirable properties and qualities to satisfy social needs.They play an important role in agriculture, in production of medicines and drugs, in environment and population control, in construction industry, in manufacturing various products such as cosmetics, textiles, dyes, soaps and detergents, plastics, rubber and a variety of metals, non-metals, alcoholic beverages, dry cells and car batteries, etc.

1.1.4. Scientists and Ethical Discipline

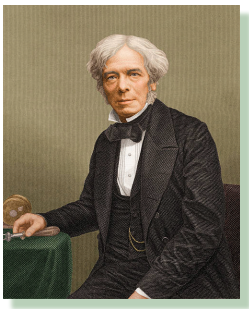
Famous Scientists in the World and Ethiopia

A scientist is someone who systematically gathers and uses research and evidence, to make hypothesis and test them, to gain and share understanding and knowledge. Some of the world and Ethiopian scientists and their contribution are listed below.

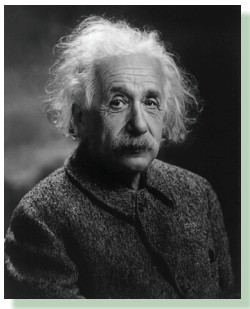
Marie Curie (1867-1934), won the Nobel Prize for the discovery of the elements polonium and radium



Isaac Newton (1643-1727) discovered the laws of motion and law of gravity.



Michael Faraday (1791-1867),discovered the generation of electricity from magnetism. He built the 1st dynamo.



Albert Einstein (1879-1955) Developed special relativity theory

Fig1.4 Some World known scientist and their works

Activities 1.4

Form a group and do the following activity

Choose one of the scientists and create a role-play for the press release following the news of his/her discovery. Various roles to consider would include: the scientist; media; fellow scientists; and the general public. Alternatively, you could choose an Ethiopian scientist and create a cartoon strip showing their discovery.

Dr.Aklilu Lemma(1935-1997) made his most important scientific discovery very early in his career, in 1964, when he discovered a natural treatment to schistosomiasis,also known as snail fever disease or bilharzia. He found that berries from the endod plant, which is commonly used to make soap and shampoos in many parts of Africa, is a potent, inexpensive and safe molluscicide, to prevent the spread of the parasitic worm. This discovery made the plant an object of scientific research in many parts of the world.

Prof.Yalemtsehay Mekonnen was born in Asela, Ethiopia on May 30, 1955. She work on human physiology, the impact of pesticides on human health, the use of plants as medicinal against human and animal disease and as the first female professor of Addis Ababa university.

Dr. Tewolde Berhan Gebre Egziabher is an Ethiopian scientist, who has worked to ensure biodiversity and the rights of communities to their genetic resources.

Gebisa Ejeta (born 1950) is an Ethiopian plant breeder, geneticist and Professor at Purdue University. In 2009, he won the World Food Prize for his major contributions in the production of sorghum.

Fig1.5 Some Ethiopian known scientist and their works

Ethical Principle in Science

Activity 1.5
Form a group and discuss on the following questions then present your opinion to the class
i. Do you think ethical discipline is important for science?
ii. List down some ethical disciplines in science

Ethics is an integral part of science. Like science, it requires to be consistent and empirically justified in our interpretation the action of scientists. Things are always get in front of us either right or wrong, good or bad, but we have to decide that what we actually want to do through our ethical point of view. The following lists are some of ethical principles that various codes address in science:

- Honesty
- Responsibility
- Objectivity
- Openness
- Competence
- Legality
- Non-Discrimination
- Carefulness.

Exercise 1.2

Choose the best answers for the following questions.

1. The natural science disciplines are
   A. interacting C. interrelated
   B. overlapping D. all of the above
2. The branches of natural science studying the composition of compounds and the processes taking place in organisms, respectively, are:
   A. Chemistry and Biology C. Biology and Physics
   B. Physics and Geology D. Biology and Geology
3. Which one of the following is true about Science?
   A. It is the study of physical and natural world.
   B. It comes from Latin word "Scientia meaning 'knowledge'
   C. It is the system of acquiring knowledge based on scientific method.
   D. All are correct
4. Which one of the following is true about Indigenous knowledge?
   A. It is based on scientific method.
   B. The knowledge derived from western countries.
   C. It incorporates local people's knowledge.
   D. It has universal perspective and commonly recognized as western science.
5. Which one of the following is the World famous scientists who discover law of motion and gravity?
   A. Michael Faraday B. Marie Curie
   C. Isaac Newton D. Albert Einstein

1.2 Common laboratory Equipment, Uses, Safety Rules and Procedures in Science laboratories

At the end of this topic, students will be able to:-

- Identify different laboratory tools (such as Balance, Beaker; Tongs, Bunsen burner, Test tubes, Petri dishes, etc.), and describe their uses;
- Prepare some laboratory equipment/tools from locally available materials;
- Demonstrate knowledge of lab safety rules and procedures;
- Practice precautionary measures in the laboratory;
- Identify potential hazards when working in the laboratory;
- Implement appropriate safety procedures when working in the laboratory;
- Demonstrate the appropriate use of personal protective equipment for a given laboratory activity.

1.2.1. Common Laboratory Apparatus

Activities 1.6

Form a group and perform the following task. From locally available materials produce laboratory tools such as beaker, measuring cylinder, balance, tong, etc. and present its use to the class.

Laboratory equipment comprises different sets of apparatus, which are designed to perform various tasks in the laboratory by students, teachers and scientists. The students can conduct laboratory work smoothly and more efficiently only when they are familiar with the apparatus commonly used in the laboratory. Some laboratory apparatus are shown in table 1 below.

Table 1 Different laboratory tools and their use

Name	Picture	Uses
Triple-beam balance		Obtaining the mass of an object
Beaker		Holding water (also used to heat liquid
Tongs		Transporting a hot beaker; and removing lid from crucible
Thermometer		Used to measure temperature
Test tubes		Holds small amounts of liquids for mixing or heating.
Petri dish		To grow and count bacteria
Graduated cylinder		Marked with milliliter (ml) scale and is used to measure volume
Bunsen burner		Heating (flame-safe) contents in the lab

Key words

Laboratory equipment refers to the various tools and equipment used by students, teachers and scientists working in a laboratory.

Laboratory report explains what you did in experiment, what you learnt and what the result mean

Project work
Prepare laboratory tools
Dear students, prepare some laboratory equipment's or tools such as beaker, measuring cylinder, balance, tongs, etc. from locally available materials

1.2.2 Laboratory Safety Rules

Activities 1.7
i. List down some laboratory safety rules
ii. Discuss hazard symbols on chemical bottles, electrical gadgets and other materials found in the laboratory

Laboratory can be considered as a place of discovery and learning. However, by the very nature of laboratory work, it can be a place of danger if proper scientfic precautions are not taken. Follow the followings laboratory safety rules precautions when you perform an activity in laboratory.

- Dress appropriately (goggles, gloves, shoes and laboratory coats).
- Tie back loose hair.
- Know the locations of safety equipment like fire extinguisher.
- Know what to do in case of an accident.
- Do not taste or smell chemicals.
- Do not eat or drink in laboratory.
- Never add water to concentrated acid solutions. Always add acid into water. Follow the Amharic AW! Not WA! (A-acid, W-water)
- Carry out only the experiments assigned by your teacher.
- Dispose of all chemical wastes properly.

1.2.3 Science Laboratory Safety Symbols and Hazard Signs, Meanings

Depending upon the scientific investigation being conducted, a lab can be filled with dangerous chemicals, Biological specimen, sharp instrument, breakable objects. In order to safe workplace and avoid accidents, lab safety symbols and signs need to be posted throughout the workplace. The following laboratory safety symbols warn of possible dangerous in laboratory user to help keep safe and informed.

Figure: 1. 6 A) hazard signs **B) Laboratory safety symbols**

1.2.4 Writig a Laboratory Report

Laboratory is a place where experiments in science is performed; therefore, it helps students learn and develop their power of observation, skill in handling apparatus varies kinds and independent thinking.

A laboratory report is a written composition of the results of an experiment. It should be written precisely and clearly, using good grammar and punctuation. Each report must include: *date*, *title, objective, theory, materials or equipment, chemical used, procedure, observation, result, discussion, and conclusion.*

Exercise 1.3

Choose the best answers for the following questions

1. Which one of the following is NOT allowed in science laboratory?
   A. Knowing the hazards of the materials being used
   B. Reading the labels on the reagent bottle carefully
   C. Wearing any type of cloth and shoes
   D. Not using laboratory glassware for eating or drinking purposes.
2. Which of the following laboratory tool is used for the approximate measurements of volume of liquids
   A. Test tube
   B. Thermometer
   C. Measuring Cylinder
   D. Dropper
3. The type of laboratory equipment categorized under measuring equipment is
   A. Bunsen burner B. Triple-beam balance
   C. Tongs D. Stand and clamp

Key terms

Conventional Science
Ethics
Hazard Signs
Indigenous Science
Laboratory
Laboratory Equipment
Natural Science
Safety Rules
Science
Scientist
Technologist
Technology

Unit Summary

- Science is a systematic method of gaining knowledge about the physical and natural world and the social aspect of human society.
- Indigenous science is process by which Indigenous people build their empirical knowledge of their natural environment
- Conventional science is the system of knowledge which relies on certain laws that have been established through the application of the scientific method to phenomena in the world around us.
- Science has two major categories, which are natural science and social science.
- Natural science has three branches which are Biology, chemistry and physics.
- Technology is the use of scientific knowledge to help human beings work easier and live better as well as enjoy their environment more.
- A person who studies technology is called a technologist.
- Science and technology plays an important role in our daily life
- Some famous scientists in Ethiopia are Dr. Aklilu Lemma, Engineer Kitew Ejigu,Dr.Gebisa Ejeta, and Prof.Yalemtsehay Mekonnen
- Famous scientists from the world are Albert Einstein, Michael Faraday. Marie Curie and Isaac Newton etc.
- Laboratory equipment comprises different sets of apparatus, which are designed to perform various tasks in the laboratory.
- Knowing Laboratory safety rule is very important to reduce risks faced during laboratory investigation.

Review Exercise

Part I. Write 'True' for the correct statements and 'False' for the wrong statements.

1. Natural science is the study of nature and natural laws.
2. Biology, Physics and chemistry do not share common areas of study.
3. Technology makes life easier or more pleasant and work more productive.

Part II: Choose the best answers for the following questions

1. The study of living things is the concern of
   A. Chemistry C. Biology
   B. Physics D. Geology
2. _____ is a branch of natural science which studies the nature of matter, energy and their interaction.
   A. Chemistry C. Biology
   B. Physics D. Geology
3. The famous Ethiopian scientist who discovered a natural treatment to Schistosomiasis or bilharzia disease.
   A. Eng. Kitew Ejigu C. Dr. Gebisa Ejeta
   B. Dr. Aklilu Lemma D. Prof. Yalemtsehay Mekonen
4. Which of the following is NOT a laboratory safety rule?
   A. You should tie back loose hair.
   B. You should add water to Acid.
   C. Do not suck solution in the pipette by mouth.
   D. When lighting a Bunsen burner, you should light the match stick before turning on the gas.

5._____is a branch of Knowledge and the systematic study of universe and its all encompasses, one that based upon facts, observation and experiments.

A. Theory B. Natural law C. Dogma D. Science

Part III Match the items in column 'A' with items in column 'B'

A	B
1. Biochemistry	A. Combination of Biology and Physics
2. Physical Chemistry	B. Combination of Geology and Physics
3. Biophysics	C. Combination of Biology and Chemistry
4. Geo-physics	D. Combination of Chemistry and Physics

Part IV; Fill in the blanks with appropriate terms.

1._____is the place where experiments in science is performed.

2. A person who study about technology is called _____.

3.Who is the famous Ethiopian scientist involved in development of African commercial hybrid strains of sorghum_____.

Part V: Give short answer to the following questions.

1. Define technology
2. What is the difference between science and technology?
3. Mention the steps to write laboratory report.
4. Why ethics in science is important?

Unit 2 MATTER IN OUR SURROUNDING

Learning Outcome

At the end of this unit, you will be able to:

- use particles theory's postulates to explain properties and behaviour of materials;
- classify matter as an element, compound, homogeneous mixture, or heterogeneous mixture with regard to its physical properties;
- describe the structure of solids, liquids and gases in terms of particle separation, arrangement and types of motion.
- differentiate between physical and chemical properties and changes of matter;
- appreciate that matter can be classified based on physical or chemical properties;
- use properties of matter to identify substances and to separate them;
- demonstrate scientific inquiry skills along this unit: observing, classifying, comparing and contrasting, making mode, inferring, communicating, asking questions, designing experiments, drawing conclusions, applying concepts.

Main Contents

Introduction

The object around us, called matter, exist in three physical forms or states. These are solids, liquids and gases. For example, water can exist as ice (solid), water (liquid) and steam (gas).The physical state of a given sample of matter depend on temperature and pressure. Energy must be added or removed to change one form or state of substance into another.

The idea that matter is made up of tiny particles is called the Particulate nature of matter.

Most of the changes that occur in our surrounding are either physical or chemical. A physical change is a change in the form of matter but not in its chemical identity. A chemical change, or chemical reaction, is a change in which one or more kinds of matter are transformed into a new kind of matter.

There are two principal ways of classifying matter: by its physical state as a solid, liquid or gas and by its chemical constitution as pure substance and mixture. Mixtures can be separated using a variety of techniques. Some of the methods used to separate mixtures are separation by hand, sieving, filtration, evaporation, magnetic separation, decantation and distillation.

2.1. Characteristics and Nature of Matter

After completing this section, you will be able to

- define matter with examples from day today life;
- demonstrate that matter is made up of tiny particles;
- state the postulates of the particle theory of matter;
- infer the particulate nature of matter from demonstration /investigation;
- apply particle nature of matter in explaining diffusion and every day effect of diffusion;

- describe and/or make a representation of the arrangement, relative spacing, and relative motion of the particles in each of the three states of matter;
- describe and explain compression in terms of distance between particles;
- use the terms melting, evaporating, condensing, and freezing/ solidification to describe changes of state;
- use the particulate nature of matter to explain: melting, freezing/ Solidification, Evaporation, Condensation.

Activity 2.1

Form a group and discuss the following questions and share your ideas with the rest of the class.

1. Describe matter by your own words?
2. Consider the followings: air, light, soil, plant, water, sound, table and heat. Try to classify them as matter and non-matter.

2.1.1. Meaning and Properties of Matter

What is matter? Matter is anything that has mass and occupies space. The term mass refers to the amount of matter present in a sample. Matter includes all things both living and nonliving that can be seen (such as plants, water, soil, rocks, table and even this book), as well as things that cannot be seen by our naked eye (such as air and bacteria). Unlike matter, energy is known and recognized by its effect. It cannot be seen, touched, smelt or weighed.

Therefore, various forms of energy such as heat, light, and sound are not considered to be matter.

Exercise 2.1

1. Classify each of the following as matter or energy (non-matter).

a. Air	e. Gold	i. Silver
b. Pizza	f. Virus	j. Cake
c. Sound	g. heat	k. Water
d. Light	h. Bacteria	l. Magnesium

2.1.2 Particulate Nature of Matter.

Activity 2.2

Form a group and perform the following activity. Then present your finding to the class.

1. Inflate a balloon and observe its shape in the class room
2. Make observations while wind blowing leaves, or dust in your surroundings. Based on the above activity, work on the given question below
   a. What do you think that matter is made of?
   b. How do the particles move around in space

The particle model of matter states that all matter is made up of tiny, moving particles with spaces between them. Matter is made of particles too small to be seen that move freely around in space. The inflation and shape of balloon indicates that it is filled with a small particle of gas such as helium, hydrogen, nitrous oxide, oxygen, or air. On other hand, from the effect of wind blowing leaves or dust, it is possible to understand the particle matter is in continuous motion. The idea that matter is made up of tiny particles is called the Particulate nature of matter.

Activity 2.3

Perform the following activities
Fill in the blank by using the following words
{Increase, less, faster, cold, temperature, water, particles, moving, more, energy}

1. Everything is made of _____.
2. Particles are always_____.
3. An increase in _____makes particles move _____.

4. An increase in_____ is the same thing as an _____ in energy.
5. The particles in hot water have _____energy than _____water.
6. The particles in ice move _____ than particles in _____.

2.1.3 Particle Theory of Matter (Particle Model of Matter)

Particulate nature of matter means that all matter is made up of discrete tiny particles. Many years later, scientists came back to Democritus' idea and added to it. The theory they developed is called the particle model of matter.

The followings are main ideas (postulate) in the particle model of matter:

1. All matter is made up of tiny particles.
2. The particles of matter move continuously.
3. The particles have spaces between them.
4. Adding heat to matter makes the particles move faster.
5. There are forces between the particles.
6. Particles of one substance differ from the particles of other substance.

Exercise 2.2

I. Give short answers

1. List the postulates of particle theory.
2. Describe the particulate nature of matter.

2.1.4. Diffusion

Experiment: 2.1

Title: simple experiment on diffusion
Objective: To discover what is meant by diffusion
Materials and Chemicals: Perfume, ink, beaker, pipette, Water (H_2O)
Procedure

1. Take a bottle of perfume and open it in one corner of the room and record how long it takes to reach to different students at different distances to smell it.
2. Add 2 or 3 drop ink into a beaker of water using a pipette and watch the ink diffuse to color the water.
3. Record your observation for the above to experment

Observation and analysis

Write your result/conclusion for the above experiments

The mixing and spreading out of a substance with another substance due to the movement or motion of its particles is called **diffusion.** It is also defined as the net movement of particles from an area of high concentration to an area of low concentration.

Diffusion in gases is very fast. This is because the particles move very quickly in all direction. Example: The smell of hot sizzling food reaches us even when we are at considerable distance.

Diffusion in liquids is slower that in gas, because the particles in liquids move slower as compared to particles in gases. Example: If a drop of ink is put into a beaker of water, then the color of ink spreads into the whole water of the beaker.

Diffusion in solids is very very slow process because the particles of solids are highly restricted to motion.

Diffusion in Daily Life

Diffusion is everywhere around us in our everyday life. The followings are some common effect of diffusion in day to day activities.

Tea: A tea bag placed in a cup of hot water will diffuse into the water.
Perfume: When perfume is produced in one part of a room, it spreads to the rest through diffusion. There are fewer of the scent-producing chemicals in the further parts of the room, so the molecules naturally spread out.
Food Coloring: A drop of food coloring in a glass of water colors the water through diffusion. The dye molecules slowly spread evenly through the liquid, creating one particular shade.
Soda: Leave a soda bottle open and the carbon dioxide bubble will diffuse and leave it flat. Air has a lower concentration of that bubbly carbon dioxide than the drink does, so the CO_2 molecules depart the beverage and spread into the air.

Exercise 2.3

1. Complete the blank space from the word box

Diffusion	low	scent	high

You can smell deodorant like axe in the classroom after someone sprays it in the hallway because the _____moves from_____ concentration in the hallway to___ concentration in the classroom. This is an example of_____

2. Arrange in an increasing order of the rate of diffusion of solids, liquids and gases.

2.1.5 Properties of Solids, Liquids and Gases

Activity 2.4

Copy the table in your exercise book and complete it using objects around you. Discuss your reasons for each decision with your group.

Substance	solids, liquids and gases	I know this is be cause….
Water	liquid	I can pour it.

According to Kinetic (particle) theory, all matter is composed of tiny particles (atoms, molecule, and ions).These particles are arranged differently in solids, liquids and gases.

Solids

In solids, the particles are arranged in fixed pattern. The particles held together strongly and are tightly packed. Particles in solid can vibrate, but, they stay in the same place. Solids have definite shape and definite volume. Examples of Solids are Stones, wood, metals etc.

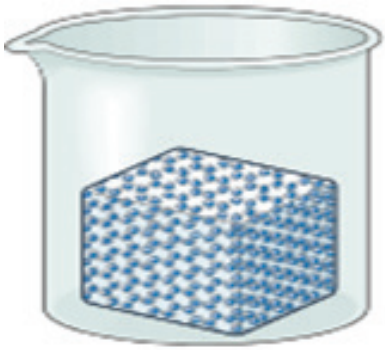


Figure 2.1: pattern of Solids

Liquids

The particles in a liquid are separated by spaces that are large enough to allow the particles to slide past each other. It takes the shape of its container because the particles can move around more freely than they can in a solid. At room temperature water, ethanol, benzene, oil are liquids.

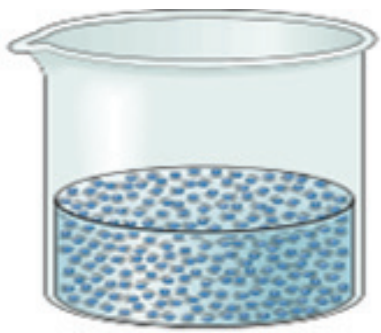


Figure 2.2: pattern of liquids

Gases

The particles in a gas are separated by much larger spaces than the particles in a liquid or a solid. Therefore, a gas is mostly empty space. For example, air, hydrogen, oxygen, carbon dioxide and nitrogen are gases.

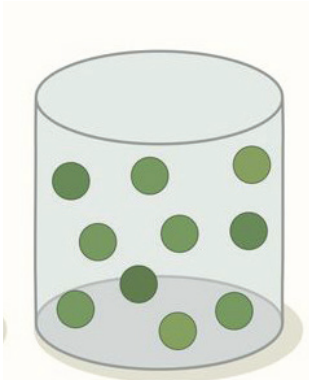
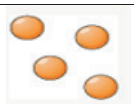
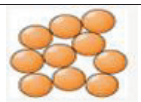
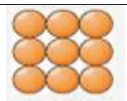


Figure 2.3: pattern of gas

Table 2.1 Properties of solids, liquids and gases

Properties	Physical states		
	Gases	Liquids	Solids
Arrangement of particles	-disorderly arranged -particles are very far apart -almost no attractive force between particles	•Less orderly arranged particles are relatively close to each other -Have relative attractive force between particles	•Orderly arranged(regular pattern) •Particles are very close to each other
Motion of particles	-flow(move) freely -flow together in random motion -are known as fluids	flow together in random motion -are known as fluids	-Do not flow or move -Vibrate in a fixed position
compressibility	Highly and easily compressible	Compressible to a very small extent	Not compressible
Volume and shape	-Have no definite Shape and volume -Assume the shape of the container and entirely fill it.	-Have no definite Shape -Assume the shape of the container - have definite volume	-Have definite Shape and volume
Density	Have very low density than liquids and solids	-Have low density than solids -Condensed state compared to gases	-Denser than all -Condensed state than all
Pressure	Exert pressure equally in all direction	Exert pressure towards depth.	Exert pressure towards gravity
Diffusion	Diffuse spontaneously in all directions with random motion	Diffuse very slowly in random motion	Difficult to diffuse
Diagram			

Exercise 2.4

I. Give short answers

1. What are the three states of matter?
2. List the properties of solids
3. Name a property of liquids that do not share with solids
4. Name a property of gas that do not share with liquids
5. Give a characteristic that is the same for liquids and solids
6. Give a characteristic that is the same for gases and liquids
7. Which state of matter can not be poured?
8. Which state of matter can be compressed easily?

II. Choose the correct answer for the following questions.

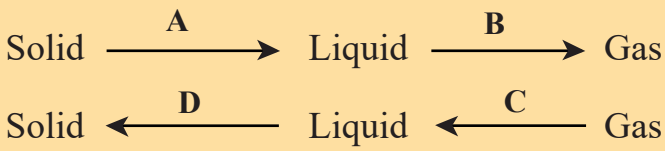
1. Which state of matter is fluid?
   A. solid B. Liquid C. Gas D. B and C
2. In which state of matter are particles close together?
   A. solid B. Liquid C. Gas D.all

2.1.6. Changes in State

Activity 2.5

Copy the flow chart. The arrows represent the process involved when matter change state.

Write the name for each process on the arrow to your flow chart.



A change of state is the change of a substance from one physical form of matter to another. A change in physical state is the most common type of physical change. Melting, freezing, evaporation, and condensation are all changes of state. The three states of matter can be interconverted without changing the composition of the substance. To change a substance from one state to another, energy must be added or removed.

How do solids and liquids change state?

When a solid is warmed, its particles gain energy and speed up, and the attraction between them decreases. Eventually they slide past one another. The process in which a solid substance changes into a liquid on heating is called **melting (fusion).**

The process of changing a liquid into a solid is called **freezing (solidification).** When a liquid is cooled, its particles have less energy, become slow down, and lock into the fixed arrangement of a solid. The temperature at which a liquid substance changes into a solid is the liquid's freezing point.

How do liquids and gases change state?

As a liquid is warmed, its particles gain energy. Some particles gain enough energy that they escape from the surface of the liquid and become a gas. The change from a liquid to a gas is called **evaporation.** The temperature at which a liquid substance changes into a gas is the liquid's boiling point.

As a gas is cooled, its particles lose energy. The attraction between particles overcomes the speed of their motion, and a liquid forms. The change of state from a gas to a liquid is called **condensation.**

How do solids and gases change state?

Some solids and gases can change state without ever becoming a liquid. The change from a solid state directly into a gas is called **sublimation.**

Some common substance undergo sublimation are: Iodine, ammonium chloride and solid carbon dioxide (dry ice).

Deposition is the change in state from a gas directly to a solid.

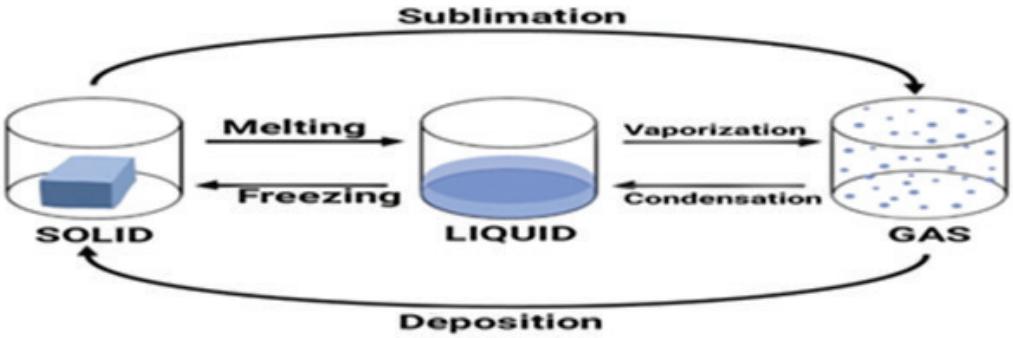


Figure 2.4: Interconversion process of the three state

2.2. Physical and Chemical Properties of Matter

After completing this section, you will be able to

- describe physical Properties;
- use physical properties of matter to identify substances;
- conduct experiments to identify properties of substances and make group report;
- identify chemical properties;
- distinguish between physical and chemical properties.

2.1.4 Physical properties of matter

Activity 2.6

Perform the following tasks in groups and present your conclusions to the class. Given the following physical properties of substances: odor, color, taste, melting point, boiling point and density.

1 Which of these physical properties have constant values under specific condition, such as temperature?

2 Which physical properties can be recognized directly by our sense organs?

3 Which of these properties are measured using instruments?
4 What will happen to ice kept in a cup in the classroom?
5. Which sense organs help us to detect color, odor and taste?
6. How do you describe the taste of lemon?

Substances are identified by their properties as well as by their composition. A physical property can be measured and observed without changing the composition or identity of a substance. For example, we can measure the melting point of ice by heating a block of ice and recording the temperature at which the ice is converted to water. Water differs from ice only in appearance and not in composition, so this is a physical change; we can freeze the water to recover the original ice. Therefore, the melting point of a substance is a physical property. Other examples of physical property is gold is a shiny yellow metal, lead has a high density. Observations of these characteristics do not change the composition.

There are two kinds of physical properties, namely, extensive and intensive physical properties.

Extensive physical properties are the properties, which depend on the amount or quantity of sample and they vary from sample to sample. Examples: length, diameter, mass, and volume

Intensive physical properties are properties which do not depend on the amount of a substance present. Examples: density, color, melting point, and hardness.

Intensive properties are useful in distinguishing between different substances because they do not vary from sample to sample.

Some Physical Properties of Substances are Listed Below

1. Physical Properties Detected by Sense Organs

Color: The color of a substance results from its interaction with light. Substances can be identified by their colors. For example, chalk is white, water is colorless, and gold is yellow and so on.

Odor: refers to the property of a substance perceived by the sense of smell. Terms commonly used to describe the odor of a substance are pungent, fragrant, spicy, fruity and odorless. For example water is odorless, flowers are fragrant, and orange smells fruity.

Caution!

Care has to be taken in smelling substance as they may be harmful

Taste: refers to physical properties that can be perceived by the taste buds of the tongue. The taste of a substance is usually described by terms like sweet, bitter, sour, salty, and tasteless. For example honey is sweet, lemon is sour and table salt is salty.

Caution!

Tasting can be used to identify substances only if the substance to be tasted is not harmful.

Activity 2.7

You are allowed to taste some acids in the forms of citric acid that are found in lemon and orange or acetic acid in the form of vinegar at home but you have never been allowed to taste any kind of acids in the laboratory. What is the reason?

Discuss your finding in group and present to the class.

2.Physical State: Physical state is the form in which a substance is found under a given conditions such as temperature and pressure. The three physical states of matter are solid, liquid and gas. The same substance may exist in different states at different conditions. For example, water exists in three physical state form as a solid below 0 ^{o}C, as a liquid between 0 ^{o}C and 100 ^{o}C, and as a vapor or gas above 100 ^{o}C.

3. Measurable Physical Properties

Measurable physical properties are the properties of a substance that can be measured using an appropriate apparatus. These physical properties have constant values under specific conditions like melting point, boiling point, density and electrical conductivity.

Melting Point: is the temperature at which a solid substance changes to its liquid state. For example, ice is the solid form of water. Ice melts to liquid (water) at 0°C. Therefore, the melting point of ice is 0°C.

Boiling Point: is the temperature at which the vapor pressure of the liquid equals the surrounding atmospheric pressure. At sea level water boils at 100°C.

Density: is defined as the mass per unit volume of a substance. It is expressed mathematically as:

$$\text{Density} = \frac{\text{Mass of substance}}{\text{Volume of substance}} \quad \text{or} \quad d = \frac{m}{V}$$

Units of density are kilogram per cubic meter (kg/m^3).

Electrical Conductivity: Electrical conductivity is the ability of a substance to conduct electricity. This is a physical property mostly characteristic of metallic substances such as copper, aluminum, iron, silver and zinc.

2.1.5 Chemical Properties of Matter

A chemical property is a characteristic of a substance that describes the way the substance undergoes or resists change to form a new substance. Chemical properties cannot be determined just by viewing or touching the substance; therefore, the substance's internal structure must be affected for its chemical properties to be investigated.

Flammability is one example of a chemical property. Reactivity between two substances is also another chemical property of matter.

Table 2.2: Comparison between Physical and chemical properties.

Physical properties	Chemical properties
Properties can be measured or observed without changing the chemical nature of the substance.	Properties that describe how a substance changes (or resists change) to form a new substance.
Easily identified.	Cannot be determined just by viewing or touching the substance.
The composition or identity of a substance not change.	The substance's internal structure must be affected for its chemical properties to be investigated.
Examples: color, density, volume, melting. boiling. Conductivity.	Examples :Flammability and reactivity

Exercise 2.5

1. Classify each of the following properties as a physical property or a chemical property.
   a. Iron metal rusts in an atmosphere of moist air.
   b. Mercury metal is a liquid at room temperature.
   c. Nickel metal dissolves in acid to produce a light green solution.
   d. Potassium metal has a melting point of 63°C.
   e. Copper metal possesses a reddish brown color.
   f. Titanium metal can be drawn into thin wires.
   g. Beryllium metal, when inhaled in a finely divided form, can produce serious lung disease.
   h. Silver metal shows no sign of reaction when placed in hydrochloric acid.
   i. Lead is denser than aluminum.
   j. Flammability of plastics.
2. Classify each of the following properties as an intensive property or extensive property.
   a. boiling point b. length c. mass d. Color e. density
   f. volume g. melting point

3. Categorize the following physical properties as physical properties recognized by our sense organs or measurable physical properties

a. Density b. odor c. taste d. melting point
e. color f. conductivity

2.3 Classification of Substances

After completing this section, you will be able to

- use the particle theory to describe the difference between pure substances and mixtures;
- differentiate between elements and compounds;
- classify common elements into metals and non-metals;
- investigate the properties of metals and non-metals and compile a list of general properties;
- investigate the properties of non-metals and compile a list of general properties;
- describe and classify mixtures as homogeneous and heterogeneous.
- use models/ particles diagrams to show differences between homogenous and heterogeneous;
- describe the relationship among elements, compounds, mixtures, homogenous mixture and heterogeneous mixtures.

Activity 2.8

Perform the following tasks in groups and present your findings to the rest of the class.

1 Conseder the following substances: chalk, bronze, sugar solution, iron, water, milk, oxygen, copper, gold, sugar, table salt, cooking oil, sulfur, air, silver, hydrogen, ink, chlorine and soil.

Classify each of them under pure substance or mixture.
Among pure substances, state whether it is an element or a compound

In addition to its classification by physical state, matter can also be classified in terms of its chemical composition into two broad categories: pure substances and mixtures.

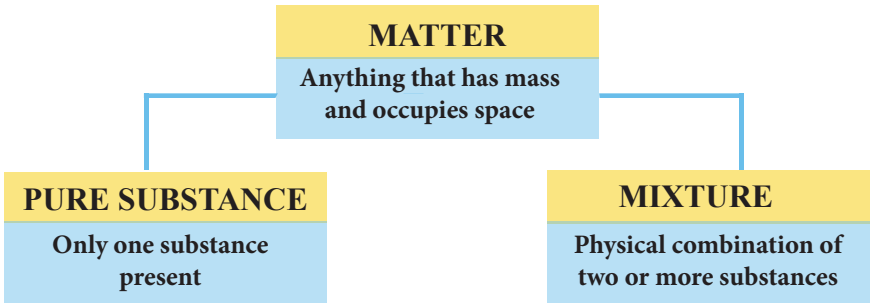


Figure 2.5 Matter falls into two basic classes: pure substances and mixtures.

2.3.1. A Pure Substance

A pure substance is a single kind of matter that cannot be separated into other kinds of matter by any physical means. All samples of a pure substance contain only that substance and nothing else. Pure water is water and nothing else. A pure substance always has a definite and constant composition. Some other common examples of pure substances are oxygen, sulfur, copper, silver, gold, sugar, table salt, water and carbon dioxide. Pure substances are classified as elements and compounds.

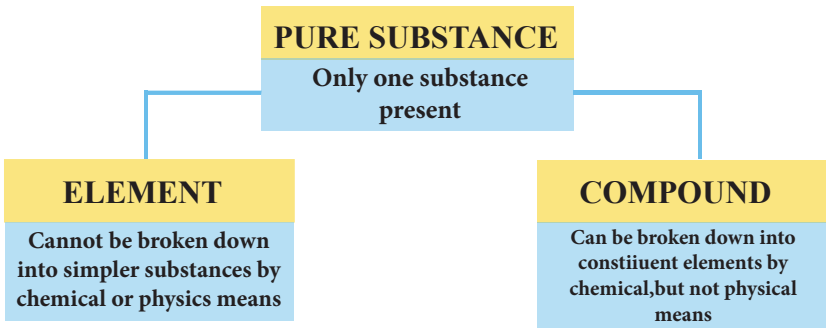


Figure 2.6 a pure substance can be either an element or a compound.

2.3.2 Elements and compounds

Elements: an element is a pure substance that cannot be broken down into simpler substances by ordinary chemical means. It is composed of only one kind of particle (atoms), which is the smallest particle of an element. At present, 118 elements are known. Among these elements, 92 of them occur naturally on earth while the rest are man-made or artificial elements. Elements might be divided into metals and non-metals.

Metals: Many chemical elements are referred to as metals. Some examples of metals are gold, iron, silver, copper, aluminum, sodium and lead.

Metals are characterized by the following physical properties

- They are shiny (lustrous) in nature.
- They are good conductor of heat and electricity .
- Their density and melting point is high.
- Moldable (Malleable): malleability is the ability of a substance to be pressed into sheets when hammered.
- Ductile: ductility is the ability to be drawn into thin wire.
- Are solid at room temperature except mercury which is found in liquid state.

Figure 2.7: Image of copper, silver & gold

Non-metals: non-metal is a chemical element that does not have metallic properties and they are few in number as compared to metals. Carbon, oxygen, sulfur, fluorine and phosphorous are some common examples of nonmetals.

Activity 2.9

Perform the following activity in groups.

Which non-metal is essential for our life?

Non metals are characterized by the following physical properties:

- they exist in two of the three states of matter at room temperature except bromine which exists as a liquid state.
- they are not shiny (dull appearance), and are non-conductors of heat and electricity
- they have relatively, low melting points and boiling points.

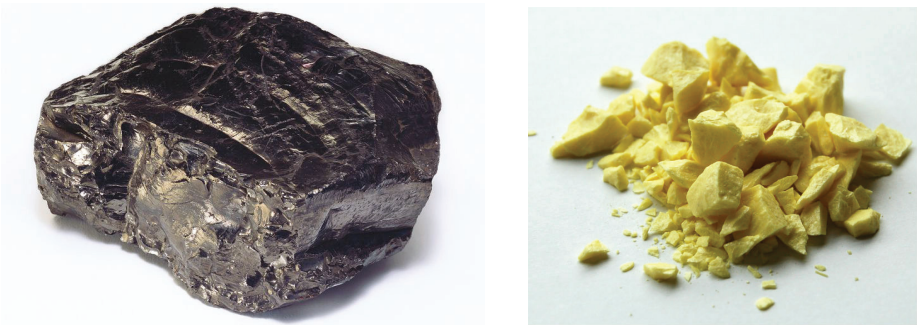


Figure 2.8: Images of carbon and sulfur

Table 2.3 Comparison between metals and non-metals.

Metals	**Non-metals**
These are solids at room temperature except mercury	These exist in all three states
These are very hard except sodium and potassium	These are soft except diamond
These are malleable and ductile	These are brittle and can break down into pieces
These are shiny	These are non-lustrous except iodine
Electropositive in nature	Electronegative in nature
Have high densities.	Have low density

Compounds

A compound is a pure substance that is made up of more than one type of atom bonded together. A compound can be broken into two or more elements by a chemical means. For example, Water is a compound. By means of an electric current, water can be broken down into the gases hydrogen and oxygen, both of which are elements. The ultimate breakdown products for any compound are elements. Elements can combine with other elements to form compounds. Sodium chloride is formed by the combination of sodium and chlorine elements. Such types of compounds that are formed by the combination of two different elements are called binary compounds.

What distinguishes an element from a compound?

A compound's properties are always different from those of its component elements, because the elements are chemically rather than physically combined in the compound.

Experiment 2.2

Title: Distinguishing compounds and mixtures.

Objective: To investigate the difference between a compound and a mixture.

Materials Required: Small bar magnet, iron filings, powdered sulfur, test tube, Bunsen burner, magnifying glass, test tube tong, sand, beam balance, watch glass and test tube made from soda glass tube.

Procedure

Part I

1. Prepare a mixture containing iron powder and sulfur powder in the ratio 7:4 by mass. Do this by weighing out 7 g of iron powder and 4 g of finely powdered sulfur onto separate pieces of filter paper (or use weighing boats).
2. Mix the two powders by pouring repeatedly from one piece of paper to the other until a homogeneous mixture (by appearance) is obtained.

Note the appearance of the pure elements and the mixture.

3. Demonstrate that iron can be separated from the mixture by physical means. Do this by bringing one end of a magnet close to the mixture as shown in figure 2.9

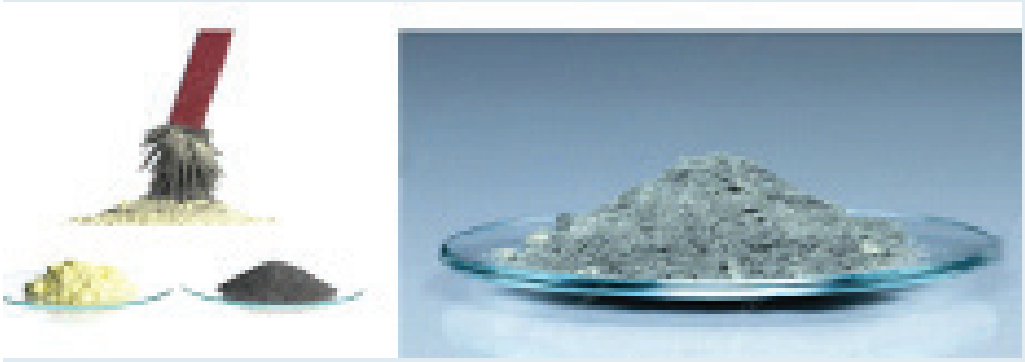


Figure 2.9 Separating iron from a mixture of iron and sulfur

Questions: i .What did you observe as you bring the magnet close to the mixture?

ii. What did you observe under the magnifying glass?

Part II 1. Place about 2 g of the mixture into a soda glass tube

2. Insert a plug of mineral wool (mineral fiber) into the mouth of the test tube. Clamp the test tube as shown in the diagram

3.Heat the powder mixture at the base of the test tube gently at first and then more strongly (use a blue flame throughout). Heat until an orange glow is seen inside the test tube. Immediately stop heating. Let the students see that the glow continues and moves steadily through the mixture.

4. Allow the test tube to cool down.

5. Once cool, it is possible to break open the test tube to show the appearance of the product, iron (II) sulfide. The test tube can be broken open using a pestle and mortar. It is advisable to wear protective gloves.
6. Take the product formed and powder it. Examine the product under a magnifying glass. Bring a magnet over it.

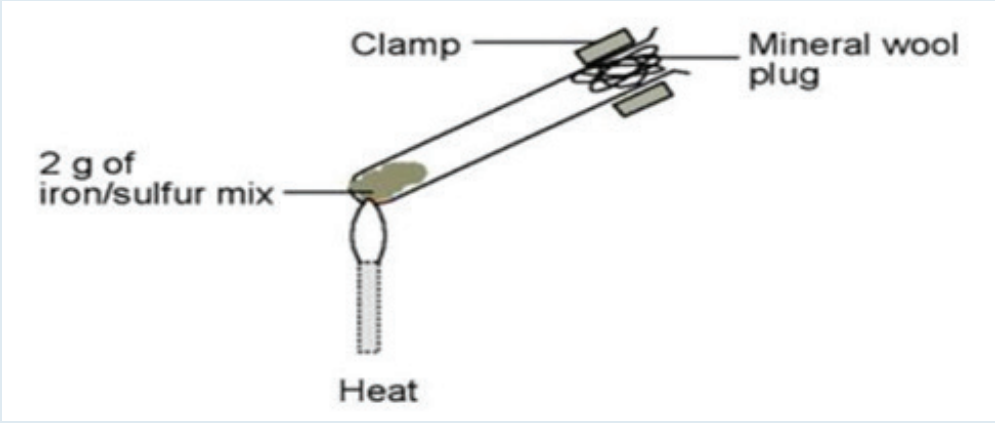


figure 2.10 The reaction between iron and sulfur

2.3.3. Mixtures

Activities 2.10

Discuss the following ideas in groups and present your opinion to the rest of the class.

1. Suppose a teaspoon of magnesium filings and a teaspoon of powdered sulfur are placed together in a metal beaker. Would this constitute a mixture or a pure substance? Suppose the magnesium filings and sulfur are heated so they react with each other, forming magnesium sulfide. Would this still be a "mixture"? Why or why not?
2. What is the difference between pure water and a solution of sodium chloride in water?
3. Do you think air is a pure substance or a mixture? Why?

A mixture is a physical combination of two or more pure substances in which each substance retains its own properties. Components of a mixture retain their identity because they are physically mixed rather than chemically combined. Consider a mixture of small rock salt crystals and ordinary sand. Mixing these two substances changes neither the salt nor the sand in any way. Common mixtures include:
Soil- a mixture of different sized particles and plant material.
Cooking oil – a mixture of vegetable oils.
Ink- contains a mixture of dyes, dissolved in alcohol and water.
Milk- contains proteins, carbohydrates, fats, water, minerals.
Air- contain oxygen, nitrogen , carbon dioxide.
Mixtures are sub classified as heterogeneous and homogeneous.
Homogeneous Mixture

Activity 2.11

Perform the following tasks in groups and present your findings to the rest of the class.
The following substances are given: air, milk, soil, salt solution, brass, chalk, water, cooking oil, gold, silver, sugar solution, Pepsi. Identify which of them are homogeneous mixtures.

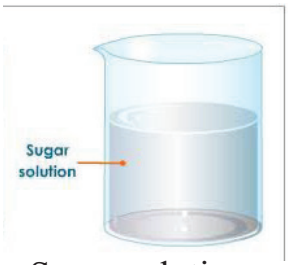
Homogenous mixtures are a combination of two or more substances that has the same composition throughout and has no visible boundary. A homogeneous mixture also called solution. For example, a mixture of table salt and water (salt solution) is a homogeneous mixture because all the parts of the solution have the same salt-water composition. Homogeneous mixtures (solutions) may exist in one of the three states i.e. solid, liquids and gas.

Table 2.4: state and type of homogeneous mixture (solution)

Type of homogeneous mixture(solution)		Common examples
Gaseous	Gas in gas	Air, mixture of oxygen and nitrogen
Liquids	Gas in liquid	Soft drinks(Pepsi, Miranda, coca cola),beer
	Liquid in liquid	Alcohol in water
	Solid in liquid	Salt solution, sugar solution
Solids	Solid in a solid	brass (Zn/Cu), Bronze (cu/Sn)

Vinegar

Steel



Sugar solution

Figure 2.11 Some common examples of homogeneous mixture

Heterogeneous Mixtures

Activity 2.12

Perform the following tasks in groups and present your findings to the rest of the class.

Consider the following substances: ethanol alcohol, bronze, sugar solution, iron, water, milk, oxygen, copper, gold, sugar, table salt, cooking oil, sulfur, air, silver, charcoal, ink, chlorine and soil.

Then identify among the list of substances which are heterogeneous mixture

Heterogeneous mixtures are a combination of two or more substances that has no uniform composition throughout and contains one or more visible boundaries between the components. The components of a heterogeneous mixture can be identified by our naked eyes or with the help of a microscope or a magnifying glass. For example, a mixture of sulfur and iron filings is a heterogeneous mixture. This is because the sulfur and iron particles remain visible and physically separated. Other Examples of heterogeneous mixture are blood, milk, mixture of sand and water, river water, muddy water, benzene and water, oil and water, dusty air, soil etc.

Soup

Soil

Blood

Figure 2.12 Some common examples of heterogeneous mixtures

Table 2.5: Differences between homogenous and heterogeneous mixtures.

Homogeneous mixture	Heterogeneous mixture
It has a uniform composition	It has a non-uniform composition
It has only one phase	There are two or more phase
The constituent cannot be seen easily.	The constituent can be seen easily
'Homo' means the same	'Hetero' means different
E.g. sugar solution, soft drinks, salt solution,	Milk, soil. sand and water, oil and water

Exercise 2.6

I Answer the following questions correctly

1. Classify each of the following as a mixture or a pure substance.
   a.Water b. uranium c. blood d .alcohol e. the oceans
   f. iron g. table salt h. brass i. hydrogen j. gold
   k. sugar. l. milk m. honey n. benzene

Of the pure substances, which are elements and which are compounds?

2. Define and give four examples illustrating each of the following terms.
   a. element b. compound c. homogeneous mixture
   d. heterogeneous mixture

II. Choose the best answers for the following questions

1. Which of the following is metallic liquid element at room temperature?
   A. bromine B. mercury C. sodium D. iron
2. Substance 'Y' is hard, lustrous solid which readily conduct heat and electricity. 'Y' is likely to be:
   A. Salt. B. Metalloid C. Metal D. Non-metal
3. Which of the following substance make a homogeneous mixture with water?
   A. benzene B. oil C. sugar D. sulfur
4. Which of the following substances is not a mixture?
   A. air B. Sea water C. Pure water D. Brass

2.4. Change around Us: Physical and Chemical Changes

After completing this section, you will be able to

- describe physical and chemical change;
- distinguish the physical and chemical changes using their characteristic;
- conduct some simple activities to show physical and chemical changes and write group report;
- observe and describe physical and chemical changes that are important in everyday life;
- identify useful and harmful physical and chemical changes.

Activity 2.13

The followings are day to day activities in your home. Copy and complete the table, by identifying which activities represent "physical change" and which one represent "Chemical change" by giving reasons for your choice

Name of activities	Physical changes	Chemical changes	Reasons
Burning of charcoal			
Melts of ice			
Dissolving sugar in water			
Fermentation			
Rusting of nail			
Evaporation of water			
Spoilage of food			
Burning candle			

Change is happening around us all the time. Changes are classified as either physical or chemical changes.

2.4.1. Physical Change

A physical change is a process in which a substance changes its physical appearance but not its chemical composition. A new substance is never formed as a result of a physical change.

Figure 2.13 melting of ice

Melting of ice, Grinding salt, tearing paper into small pieces, Making an iron bar magnetic, evaporation of water, dissolving sugar in water and breaking a stick are common examples of physical changes.

2.4.2. Chemical Change

A chemical change is a process in which a substance undergoes a change in chemical composition. Some examples of chemical changes are: Iron nail going rusty, heating magnesium ribbon, burning candle. photosynthesis,fermentation,etc

Figure 2.14 burning of candle

Experiment 2.3

Title: Rusting of iron.

Objective: To investigate the type of change that occurs during rusting of iron

Materials and Chemicals: test tube rack, iron nail, test

Procedure:

1. Put a few lean shiny iron nails into a test tube containing some fresh tap water. The water contains dissolved air.
2. Set the test tube in a rack. After a few days observe the change that has taken place.

Figure: 2.15 rusting of iron

Observation and analysis

1. What color do you observe on the iron nail?
2. Is the change physical or chemical? Why?

2.4.3 Characterstics of physical and chemical changes

Activity 2.14

Discuss the following ideas in groups and present your opinion to the rest of the class

1. List the characteristics of physical and chemical changes you know
2. Compare and contrast the characteristics of physical and chemical changes

Characteristics of Physical Change

- No new substance is formed.
- The composition of substance is not altered.
- It is easily reversed by physical means.
- Energy changes are not necessarily.
- It is a change in physical property.

Characteristics of Chemical change

- New substances with new properties are formed
- The composition of substance altered
- It is accompanied by Energy changes
- The change is not easily reversed
- It is a change in chemical property

2.4.4 Useful and Harmful physical and Chemical Changes

Activity 2.15

Perform the following tasks in groups and present your findings to the rest of the class.

List the important and harmful physical and chemical changes that encounter in our life.

a. Important physical changes
b. Important chemical changes
c. Harmful physical changes
d. Harmful chemical changes

Useful effect of Physical changes

Physical changes are useful in the following ways:

- Evaporation and condensation create water cycle
- Freezing preserves food, medicine, and other materials
- Melting, cutting, bending and mould different tools and accessories
- To get substances in the form, shape or size we want
- To mix two or more substances together
- To separate substances from their mixtures

Harmful effect of Physical changes

Even many physical changes are useful, it may also be harmful in several way: like cutting tree, bad weather condition, oil spills, etc.

Useful effect of chemical changes

Chemical Changes are useful in the following ways:

-Photosynthesis: chemical changes which occur in plants (photosynthesis) produce substances which enable plants to grow We depend on plants for our food. The change which occur in the food we consume are chemical changes.

-Energy production: Most of the energy used nowadays, with the exception of wind, water and nuclear energy, is chemical energy. This energy released as heat or electricity when certain chemical change takes place.

Food and medicine production, food digestion, fermentation, food cooking, etc. are also some important chemical change in our life to produce new substance.

Harmful effect of Chemical changes

In contrast to its usefulness, some chemical change has negative impact. For example rusting (rusting of cars, bridges, and ships), souring food, burning of fuel, smoke emission, plastic disposal, dumping of chemicals, etc. are harmful chemical changes in our life.

Exercise2.7

I Answer the following questions correctly

1.Classify the following as physical changes or chemical changes.

a. the cutting of wood
b. interaction of food with saliva and digestive enzymes
c. The vigorous reaction of potassium metal with water to produce hydrogen gas and potassium hydroxide.
d. Straightening a bent piece of iron with a hammer.
e. The ignition and burning of a match.
f. photosynthesis
g. boiling of an egg.
h. boiling of water
i. dissolution of salt

II. Choose the best answers for the following questions

1. Change in size, shape and state of a substance is a____

A. chemical change B. physical change
C. cyclic change D. none

2. Which of the following statements is correct?

A. Evaporation is a chemical change
B. Digestion of food is a chemical change
C. Burning of paper is physical change
D. all

3. Among the following which one is a physical change?

A. Burning of candle B. Fermentation
C. making an iron bar magnetic D. all

2.5. Separation of Mixtures and its Application

After completing this section, you will be able to

- list methods of separation of mixtures;
- give some specific examples of mixtures that can be eparated by filtration, decantation, simple distillation, magnetic separation and using separator funnel;
- name apparatuses used in decantation, filtration, simple;

- assemble apparatuses used in decantation, filtration, simple distillation, separator funnel;
- conduct and report on an investigation that uses physical means such as particle size, density, boiling point, solubility and magnetism to separation;
- perform simple activities in group, to carry out the separation of mixtures using local materials and write a group report;
- compare and evaluate the different ways of separating mixtures from products in community.

Activity 2.16

Discuss the following questions in groups and present your conclusion to the class

1. Write the common separation methods you know for the following common mixtures from your daily life experiences

A. A mixture of Teff and peas
B. A mixture of iron filings and sulfur powder
C. A mixture of chalk particles in water
D. A mixture of cooking oil and water
E. Salt solution
F. A mixture of alcohol and water
G. A mixture of salt and sand
H. A mixture of orange, banana and mango

2.5.1. Separation Techniques of Mixture

Most of the substances around us exist in the form of mixtures. However, these mixtures can be separated into pure substances using various separation techniques. The process of separating the constituent substances of a mixture by physical methods, taking advantage of the differences in their physical properties is called **separation process.** Some of the methods used to separate mixtures are Separation by hand, sieving filtration, evaporation, magnetic separation, decantation and distillation.

Note that the methods for the separation of mixtures into their components depend on the differences in the size, magnetic property, melting point, boiling point, solubility, etc. of the components. We will discuss some of the methods that are used to separate the components of mixtures.

I. Magnetic Separation

Magnetic separation is used to separate magnetic and non-magnetic substances in a mixture. For example, if sand is mixed with iron filings the mixture is heterogeneous. To separate the iron filings from the sand, you can use a magnet. The iron filings (magnetic component) are attracted by the magnet, while the sand is not attracted.

a b

Figure 2.16 (a) the mixture contains iron fillings and sand. (b) A magnet separates the iron fillings from the mixtures.

Experiment 2.4

Title: Separation of mixture using bar magnet

Objective: To separate a mixture of iron fillings and sand

Apparatus and chemicals

Magnetic bar, Iron fillings, Sand, Petri dish/ plastic plate/bowl
Plastic bag/wrapper, Spatula

Experimental Procedure

1. Mix the sand with the iron filings in the plastic plate.
2. Wrap the plastic bag around the bar magnet

3. Suspend the bar magnet over the plate
4. The iron would be collected / attracted to the surface of themagnetic bar
5. Carefully remove the plastic bag around the magnetic bar and scrape off the iron filings

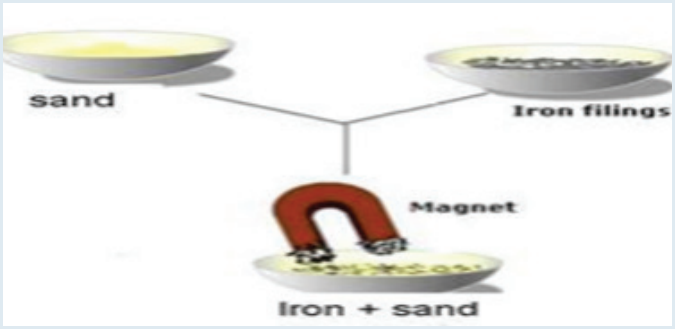


Figure 2.17 mixture of sand and iron fillings

Observation and analysis

1. Why was the sand not attracted by the magnet? What can you conclude from this experiment?

II.Decantation

What type of separation method is used to get a cup of clear coffee as it is poured from coffee pot ("jebena") as shown in Figure 2.18?

Figure 2.18: Separation by decantation

Decantation is the process of separation of liquid from solid and other immiscible (non-mixing) liquids, by removing the liquid layer at the top from the layer of solid or liquid below. The process can be carried out by tilting the mixture after pouring out the top layer. This process can also be used to separate two liquids that do not mix with each other like cooking oil and water. When we leave the mixture of cooking oil and water, two separate layers are formed, where water at the bottom and oil, being lighter, at the top. We can remove the oil layer from the top by pouring it into another vessel, which leaves us with the water layer at the bottom.

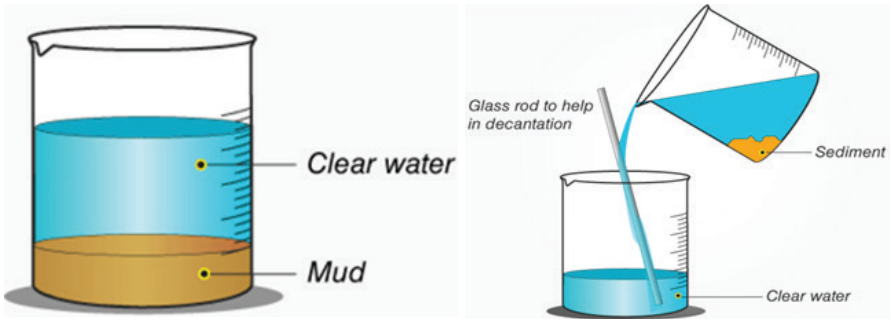


Figure 2.19: Decanation of solid from a solid-liquid mixture

Separating funnel: When two liquids do not mix, they form two separate layers and are known as immiscible liquids. These two liquids can be separated by using a separating funnel. A separating funnel is a special type of glass funnel, which has a stop-cock in its stem to regulate the flow of liquid. It will separate the immiscible liquids into two distinct layers depending on their densities. The heavier liquid forms the lower layer while the lighter one forms the upper layer. Remove the stopper and open the tap to run the lower layer into a beaker. You will be left behind with just the upper layer in thefunnel. Collect this liquid into

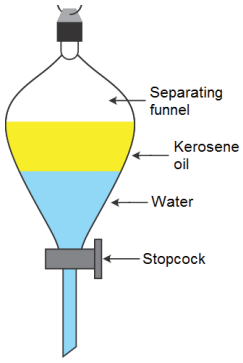


Figure 2.20: Separation of immiscible liquids using separating funnel

another beaker. Examples: Kerosene and water mixture is separated by using separating funnel method. This method is also used to separate oil and water.

III. Filtration:

Filtration is a process by which insoluble solids can be removed from a liquid by using a filter paper. A filter paper is a special type of paper which has pores that are tiny enough to let only liquids pass through it. If you pass a solution through filter paper, any undissolved solid particles will get left behind on the paper whereas the liquid will filter through. The liquid that passes through is called the filtrateand the undissolved solid particles are called residue. Example: A mixture of chalk powder and water, soil and water, sand and salt solution, etc. can be separated by this method. In practical application, filtration is a key step in the purification of the tap water you drink.

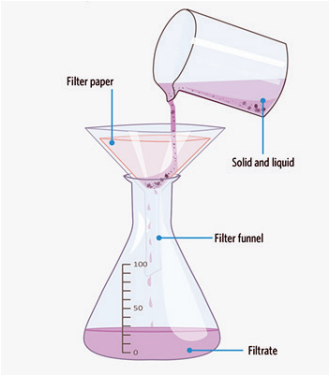


Figure 2.21: Filtration

Iv. Evaporation

Activity 2.17

Perform the following activity.
Dissolve sodium chloride (or any other soluble salt) and water to form a homogeneous mixture (solution). How can you recover the salt again?

Figure 2.22: Evaporation of a solution

Evaporation is a method used to separate a soluble solid from a liquid in a solution or the process. of vaporizing the solvent to obtain the solute. It is used to separate a mixture containing a non-volatile, soluble solid from its volatile, liquid solvent. We can separate salt from a solution by evaporating the water from the solution.

.

V. Distillation:

This method is used for the separation of a mixture containing two miscible liquids that boil without decomposing and have a large difference between their boiling points. It is also used in obtaining pure water from salt solution. Process of conversion of a liquid into vapor by boiling, and then re-condensing the vapor into liquid is called distillation. In simple distillation, a mixture is heated and the most volatile component vaporizes at the lowest temperature. The vapor passes through a cooled tube (a condenser), where it condenses back into its liquid state. The condensate that is collected is called distillate. Figure 2.23 show the simple distillation set up

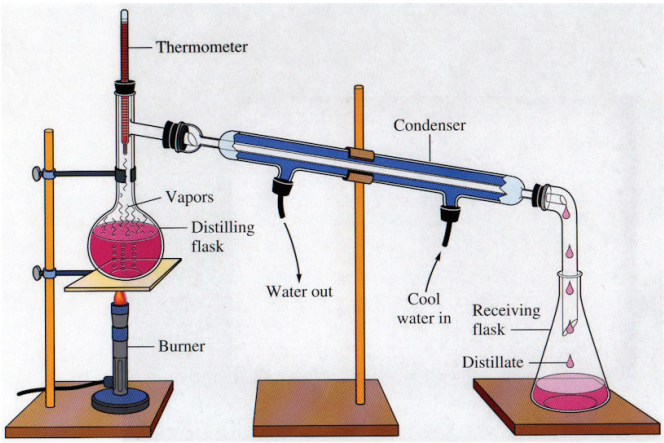


Figure 2.23: simple distillation set up

A mixture of two miscible liquids can also be separated by simple distillation. Liquids which mix with each other to form a solution are called miscible liquids. The mixture of alcohol and water, benzene and oil are some examples of miscible liquids. Consider a mixture of alcohol and water. Ethanol, which is an alcohol boils at 78°C, and water boils at 100°C. When the mixture is heated, the alcohol, which has the lower boiling point vaporizes more rapidly than the water. The vapor of alcohol passes through the condenser and then collected as a distillate in the receiver.

Activity 2.18

Perform the following tasks in groups and present your conclusion to the class. The following mixtures can be separated using a combination of separation techniques. Mention all the possible separation techniques.

i Mixture of salt, sand and water
ii Mixture of common salt, iron filling and salt
iii Mixture of oil, water and sand
iv sugar and clay

Is one separation method enough when salt and sand is mixed with water? Sometimes to separate such mixture may require combination of two or more techniques. For example, a mixture of common salt and sand can be separated by using the process of dissolving, filtration and evaporation. The first stage of separation is adding water to the mixture. The salt dissolves in water and forms a solution, but not the sand. Then by using filtration, the sand can be separated from the salt solution. Finally evaporation of the filtrate will cause the water to escape leaving the salt behind.

Activity 2.19

Perform the following tasks by asking your parents/guardians / grandparents/elders in the neighborhood on the separation techniques used in daily lives. Prepare a table of such activities of daily life in which sedimentation,
decantation, filtration and evaporation are used/occurs. Then present your findings to the whole class.

2.5.2. Application of Separation Techniques.

Table 2.5 application of various separation techniques

No.	Physical process	Application
	Filtration	• River water is potable • Separation of honey from its comb
	Evaporation	• Common salt is obtained from sea water on large scale.
	Distillation	• Alcoholic beverage such as whisky, gin, brandy, areki are manufactured.
	Magnetic separation	• To separate iron and steel from non –magnetic objects such as, glass, plastic, aluminum, etc.
	Sedimentation followed by decantation	• Drink homemade coffee, tella

Project Work

Separation of mixtures using local materials

There is a mixture of table salt, sulfur powder and iron filing inside a beaker.

By using any local materials that are found around you, you should try to separate this mixture into their components.

Hint: Both sulfur powder and iron filing are insoluble in water whereas sodium chloride (table salt) is soluble in water.

Write a group report: In your report indicate the separation techniques and the materials used during the separation processes.

Exercise 2.8

I. Give short answers for the following questions.

1. How would you separate the following?

a. salt solution b. Common salt and sand

c. iron and charcoal d. Oil, water and sand

e. nitrogen and oxygen

2. Name the separation technique shown in the following diagram

A

B

C

D

E

F

Key Terms

Matter	Magnetic Separation	Physical Changes
Particle Theory	Decantation	Compounds
Diffusion	Pure Substance	Elements
Physical Property	Mixture	Distillation
Chemical Property	Non Metal	Evaporation
Homogeneous Mixture	Heterogeneous Mixture	Filtration
Sublimation	Density	Freezing
Boiling Point	Freezing Point	

Unit Summary

- Matter is anything that has mass and occupies space. It can exist in three states: solid, liquid, and gas.
- The idea that matter is made up of tiny particles is called the Particulate nature of matter.
- Particle theory of matter tell us the particles of matter are always moving.
- Diffusion is the mixing and spreading out of a substance with another substance due to the movement or motion of its particles.
- The three states of matter (solids, liquids and gases) can be interconverted without changing the composition of the substance. to change a substance from one state to another, energy must be added or removed. Melting, freezing, evaporation, and condensation are all changes of state.
- A physical property can be measured and observed without changing the composition or identity of a substance.
- A chemical property is a characteristic of a substance that describes the way the substance undergoes or resists change to form a new substance.
- Matter can be classified in terms of its chemical composition into two broad categories: pure substances and mixtures pure substances.

- A pure substance is a single kind of matter that cannot be separated into other kinds of matter by any physical means.
- Pure substances are classified as elements and compounds.
- An element is a pure substance that cannot be broken down into simpler substances by ordinary chemical means.
- A compound is a pure substance composed of two or more elements that are combined chemically in a definite proportion by mass.
- A mixture is a physical combination of two or more pure substances in which each substance retains its own properties.
- A homogeneous mixture (also known as solution) has a uniform composition and properties throughout.
- Heterogeneous mixture is a mixture that does not have a uniform composition throughout.
- Changes are classified as either physical or chemical changes.
- A physical change is a process in which a substance changes its physical appearance but not its chemical composition.
- A chemical change is a process in which a substance undergoes a change in chemical composition.
- Mixtures can be separated using a variety of techniques. The process of separating the constituent substances of a mixture by physical methods, taking advantage of the differences in their physical properties is called separation process.
- Some of the methods used to separate mixtures are Separation by hand, Sieving Filtration, evaporation, magnetic separation, decantation and distillation

Review Exercise

Part I. Write 'True' for the correct statements, and 'False' for the wrong statements.

1. Depending upon the temperature, water can exist in solid, liquid or gas states.
2. A gas has neither a definite volume nor a definite shape.
3. Dust, smoke, bacteria, air born viral particles are components of particulate matter.
4. Elements can be further decomposed by ordinary chemical means.
5. Heterogeneous mixture contains one phase.

Part II: Choose the corect answers for the following questions

1. All of the followings are matter except
   A. plant B. stone C. air D. sound
2. Which of the following decrease during the phase (state) changes of
   Solids ⟶ liquids ⟶ gases
   A. Degree of order among particles
   B. Energies of particles
   C. Speed of particles
   D. Distance among particles
3. Which of the following is not the property of solids
   A. Solids have little tendency to diffuse
   B. Solids are extremely difficult to compress
   C. Solids are fluids
   D. Solids have definite volume and definite shapes
4. The particular physical state of substance depends on _
   A. temperature B. pressure
   C. strength of intermolecular force D. all
5. Which of the following has no definite shape and volume?
   A. Water B. carbon dioxide C. iron D. gold

6. The interaction of substance with light results
   A. taste B. odor C. color D. texture
7. Which of the following is not a physical change?
   A. sublimation of iodine C. tearing a piece of cloth
   B. burning of wax in a candle D. dissolving sugar in a tea
8. All of the following are heterogeneous mixtures except
   A soil C. salt solution
   B. mixture of water and oil D. blood
9. Which of the following is not a metal
   A. iron B. sulfur C. copper D. sodium
10. Which of the following has variable composition?
   A. Water B. salt C. milk D. silver
11. A mixture of sand and sugar can be separated by
   A. evaporation followed by distillation
   B. filtration followed by evaporation
   C. dissolution followed by filtration and evaporation
   D. dissolution followed by evaporation and filtration.
12. the conversion of a vapor directly to solid without passing through a liquid state is called_____
   A. fusion B. evaporation C. sublimation D. deposition
13. Distillation is used in the process of preparation of
   A. coffee B. areki C.tella D. honey
14. Identify the heterogeneous mixture among the following
   A. sea water B. blood C. bronze D. air
15. Grade 7 students in a certain school were given the task of separating, iron fillings, sand and salt. Which of the following process is the most appropriate order?
   A.Evaporation-dissolution - filtration- magnetic separation
   B. Dissolution –magnetic separation - filtration_ evaporation
   C. Magnetic separation -dissolution - filtration –evaporation
   D. Magnetic separation - dissolution – evaporation- filtration
16. I. Physical changes are easily reversible.
   II. Physical change do not produce new substance
   III. Physical change do not involve change in mass.

Which of the above statement are correct?

A.I B I and III C.II and III D. I, II and III

17. Which two state of matter are fluids?

A. Solid and liquid C. Liquid and gas

B. Solid and gas D. Plasma and solid

18. All of the following are same process. EXCEPT

A. Condensation C. Crystallization

B. Freezing D. Solidification

Part III: Match the items in column 'A' with items in column 'B'

A	B
1. Melting (fusion)	A. process of changing liquid to gas
2. Evaporation	B. process of changing liquid to solid
3. Sublimation	C. Process of changing solid to gas
4. Freezing	D. Process of changing gas to solid
5. Deposition	E. process of changing gas to liquid
6. Condensation	F. Process of changing solid to liquid

Part IV; Fill in the blanks with appropriate terms.

1. Immiscible liquids can be separated by using_____
2. Separating a solid from a solution by cooling is_____
3. The separation technique that involves heating a solution until the liquid changes into a gaseous state, leaving behind a solid is known as____
4. When rain falls, this is because the water vapor in the clouds is condensing into liquid _____?
5. I put dirty water into a funnel with a paper lining. Clean water comes out of the funnel into my beaker, and solid dirt gets left behind on the paper. This is an example of_____.

Part V: Give short answer to the following questions

1. What is diffusion?
2. Why do solids have fixed shape and volume?
3. Write the difference between extensive and intensive physical properties.

Unit 3 ELEMENTS,COMPOUND AND CHEMICAL REACTION

Learning Outcome

At the end of this unit, you will be able to:

- compare elements to compounds and how they are represented by symbols and formulae;
- identify and write symbols of common elements or compounds;
- name compounds given their formula and write formula given the name of the compound;
- use symbols and chemical formulae as a way of communicating information about elements and compounds;
- state and apply the law of conservation of mass to writing balanced equations;
- interpret chemical formulae of,compounds in terms of the elements present and the ratios of their atoms.

Main Contents

Introduction

Pure substance, whether an element or compound, has its own unique name, symbol or formula. Scientists use chemical symbols in place of the names of the elements because it helps for scientists in writing chemical formulas and equations. The symbols and formulas are designed in such a way that they are internationally accepted. Therefore, they enable all scientists in the world to communicate easily. Symbols and formulas of elements or compounds are used in certain combination-ratios as a short hand representation of chemical reaction and these short hand languages is known as chemical equation.

3.1. Elements and Their Representation

After completing this section, you will be able to:

- define element;
- identify symbols of some common elements;
- write chemical symbols for common elements.

Activity 3.1

Form a group and discuss the following questions and share your ideas with the rest of the class.

1. From your previous knowledge, what is an element?
2. List some common elements you are familiar with. Try to classify them as

   a. metal b. non -metal

3.1.1. Common Elements

An element is a pure substance that cannot be broken down into simpler substances by ordinary chemical means. An element compose of only one kind of matter (atoms).There are 118 known elements.92 out of 118 elements are naturally occurring elements.

As you have learnt in unit two, elements are classified as metal and nonmetals. Oxygen, aluminum, iron, calcium, sodium, potassium, magnesium, hydrogen, nitrogen, gold, silver, copper, sulfur, and chlorine are some common elements.

3.1.2 Chemical Symbols

Activity 3.2

Perform the following activities in group. Then present your opinion to the whole class.

1. What is an atomic symbol?
2. Explain why do some symbols for examples He, Cl and Si have two letters?

Scientists use symbols as abbreviation of names of an elements. An atomic symbol is defined as shorthand way of representing elements or atoms of an element. Every element has its own symbol. No two elements can have the same symbol.

How to write symbols of elements?

Chemists use chemical symbols in place of the names of the elements because they are much easier and quicker to write a symbol. A symbol for an element is taken from the first letter or the first letter plus another letters of the common name or Latin/Greek name of the element. If a symbol has one letter, it is written in capital, letter besides if it has two letters, the first is in capital and the second is in small letter.

For example, S stands for sulfur, O stands for oxygen, and K represents potassium. In the case of potassium, the symbol is derived from the Latin name, Kalium.

Why are not all elements symbolized by the first letter of their names?

The names of some elements such as carbon and calcium begin with the same letter "C". Therefore we cannot use the letter "C" as a symbol for both elements. Hence two letters are used for other elements except one. The first letter "C" is assigned as a symbol for carbon. The other element calcium is represented by two letter symbols Ca The same thing is true for hydrogen and helium. The first letter "H" is assigned as a symbol for hydrogen while "He" symbol stands for element helium.

Table 3.1: Name and symbols of some elements

Name of elements	Symbol	Name of elements	Symbol
Hydrogen	H	Magnesium	Mg
Helium.	He	Aluminum	Al
Lithium	Li	Silicon	Si
Beryllium	Be	Phosphorus	P
Boron	B	Sulfur	S
Carbon	C	Chlorine	Cl
Nitrogen	N	Argon	Ar
Oxygen	O	Calcium	Ca
Fluorine	F	Zinc	Zn
Neon	Ne	Bromine	Br
Magnesium	Mg	Iodine	I

Table 3.2: Symbols of element derived from Latin names

English name	Latin name	Symbol
Sodium	Natrium	Na
Potassium	Kalium	K
Iron	Ferrum	Fe
Copper	Cuprum	Cu
Silver	Argentum	Ag
Gold	Aurum	Au
Lead	Plumbum	Pb
Tin	Stannum	Sn
Mercury	Hydrargyrum	Hg

Exercise 3.1

I. Write true for correct statement and false for wrong statement

1.Ca is the symbol of sodium.

2.Water is not an element.

3.Elements are pure substances.

Choose the correct answer from the given alternatives.

4.Which of the following is the correct chemical symbol for silicon?

A. S B. Si C. SI D. Sl

5."C" stands for_____

A. calcium B. Chlorine C. Carbon D. Copper

III. Fill the missing symbols and names of the elements in the following table

Name of element	Symbol	Name of element	Symbol
Potassium		Iodine	
	He		B
Chlorine		calcium	
	Cu	Nikel	
Gold			H
	Li	silver	

3.2. Compounds and their Representation

After completing this section, you will be able to

- define compound as a substance formed when two or more elements chemically combined together.
- define valence numbers as the combining power of an atom.
- write the formulae of simple binary compounds using symbols and valences.
- name binary compounds.
- describe polyatomic ion.
- write the chemical formulas of common compounds that contain polyatomic ions.

- Name compounds containing polyatomic ions.
- Identify the elements and number of atoms, and give a chemical formula.

3.2.1. Compounds

Activity 3.3

Form a group and discuss the following questions and share your ideas with the rest of the class.

1. From your previous knowledge, what is a compound?
2. Copy the table on your exercise book and classify the substance as an element or, a compound.

Substance	Element	Compound
Sodium chloride (table salt)		
Water		
Gold		

As you have learnt in unit two, a compound is a pure substance consists of two or more elements which have been chemically combined. For example, water is a compound of hydrogen and oxygen. Each of its molecules contains two hydrogen atoms and one oxygen atom. There are many different compounds. Some examples of compounds are sodium chloride, iron sulfide, carbon dioxide, sugar, calcium carbonate, calcium oxide, etc.

3.2.2. Chemical Formulas

It is the symbolic representationof an element or a compound.
Chemical formulas can be classified as formulas of elements and formulas of compounds.

Formulas of Elements

The formula of an element consists of one kind of symbol.
A **molecule** is the smallest particle of an element or a compound that has a stable,independent existence.

The elements helium,neon,argon,krypton,xenon and radon are collectively known as noble gas, or monoatomic gases Because they exist uncombined as single atoms,they are also known as monoatomic gases. Their formula is the same as their symbol.Example He for Helium, Ne for Neon, Ar for Argon.

Some nonmetallic elements exist as molecules containing two, four, or eight atoms. Hydrogen, nitrogen, oxygen, fluorine, chlorine, bromine and iodine are found as **diatomic molecules.**

Table 3.3 symbols and formulas of diatomic elements.

Name	Symbol	Formula
Hydrogen	H	H_2
Nitrogen	N	N_2
Oxygen	O	O_2
Fluorine	F	F_2
Chlorine	Cl	Cl_2
Bromine	Br	Br_2
Iodine	I	I_2

Elemental formula also found in homo polyatomic molecules that contain more than two atoms. Examples Ozone-O_3, Phosphorus-P_4 and Sulfur-S_8

Formulas of Compounds

Elements combine to form compounds. Just as symbol is a shorthand way of representing an element, a chemical formula comprising two or more different symbols, is a short hand representation of a compound. In formulas of a compound, the following points are noticed.

- In each formula, the symbol of elements which form the compound are given. Each symbol is immediately followed by a subscript showing the number of atoms of that element.
- Chemical formulas indicate the relative number of atoms of each element present in the compound.
- For example, water (H_2O) is a compound of hydrogen and oxygen. Each of its molecules contains two hydrogen atoms (2H) and one oxygen atom (O).

Exercise 3.2

I.Choose the correct answer from the given alternatives.

1. Elements exists as a diatomic and polyatomic molecular form except_____.

A. phosphorus B Nitrogen C Oxygen D Neon

2. For which of the following do the atom and molecule have different formula?

A. Helium B. Argon C. Nitrogen D. Neon

3.2.3. Valence Number

Activity 3.4

Discuss in groups and share your ideas to the class

What is a valence number?

Elements combine in accordance with the laws of nature at atomic levels. Each element in a formula of a compound has a combining power. The combining power of an element is called **valence.** If we know the combining power (valence number) of the elements, it is easy to write the formula of a compound. Most common elements have valence 1, 2, or 3. Some elements have more than one valence number, which is different combining powers under different conditions. Common examples of these elements that have variable valence are iron, copper, lead and tin.Ions are atoms that have positive or negative charge. The number of negative or positive charge an ions carries is equal to the valence number of the ion.Thus, the valences of Cl^-, O^{2-} and Al^{3+} are 1, 2 and 3 respectively. The following table shows the combining power of some common elements.

Table 3.4: valences of some common elements.

Elements	Valence 1		Valence 2		Valence 3	
	Name	symbol	Name	symbol	Name	symbol
Metals	Lithium	Li	Magnesium	Mg	Aluminum	Al
	Sodium	Na	Calcium	Ca	Iron(III)	Fe
	Potassium	K	Iron(II)	Fe		
	Copper(I)	Cu	Zinc	Zn		
	Silver	Ag	Lead(II)	Pb		
Non-metal	Chlorine	Cl	Oxygen	O	Nitrogen	N
	Bromine	Br	Sulfur	S		
	Iodine	I				
	Fluorine	F				

3.2.4. Formulas of Binary Compounds

Binary compounds are compounds formed from two different types of elements. To write formulas of binary compounds, follow the following simple rule

i. Write the symbol of the elements.

ii. Write the valence number above the symbol.

iii. Criss-cross the valence numbers to conserve charge or to become the compound electrically neutral and write below the symbols. If the valence number is one, omit the subscript.

Examples

a. Write the chemical formula for Potassium iodide

Solution Step 1: K I

Step 2: $\overset{1}{K}\ \overset{1}{I}$

Step 3: $\overset{1}{K}\ \overset{1}{I} \rightarrow K_1I_1$

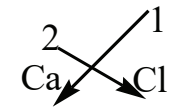
Since the subscript is 1 we omit and the chemical formula for Potassium iodide is KI.

b. Write the chemical formula for calcium chloride

Solution

Step 1 Ca Cl

Step 2: Ca Cl (2, 1)

Step 3  = $CaCl_2$

Therefore, the chemical formula of calcium chloride is $CaCl_2$

Exercise 3.3

I. Give short answers

1. Write the chemical formula for
   - A .Copper (II) oxide
   - B. Magnesium nitride
   - C. sodium chloride
   - D. Aluminum Oxide
   - E. Iron (III) Oxide
   - F. Iron (II) bromide
   - G. Silver Oxide
   - H. Calcium fluoride
2. Write the chemical name for
   - A. MgO B. FeS C. AgCl

II. Choose the correct answer from the given alternatives.

3. Which of the following is the chemical formula of aluminum nitride?
   - A. Al_5N_3 B. Al_3N_2 C. AlN_3 D. AlN
4. How many valence numbers does an Aluminium have?
   - A. 1 B. 2 C. 3 D.5

3.2.5. Naming Binary Compounds

Activity:3.5

Perform the following activities.

A student wrote this name for a compound made of calcium and sulfur: Sulfur calcium. What is wrong with this name? Write the correct name for the compound.

In naming a compound, the positive ion (metal) mentioned first followed by the negative ion (nonmetal). Binary compound is a compound that is made of only two different elements in a certain whole number ratio.

Rules for naming simple binary compounds.

1.If the binary compounds consists of metal and non -metal, the name of the metal named by its elemental name while the last letters of the non-metal is replaced by the suffix-ide.

Table 3.5 Names of nonmetallic elements in binary compounds

Nonmetallic Element	Name in Binary compounds	Non-metallic Element	Name in Binary compounds
Nitrogen	Nitride	Bromine	Bromide
Oxygen	Oxide	Iodine	Iodide
Fluorine	Fluoride	Phosphorous	Phosphide
Chlorine	Chloride	Sulfur	Sulfide

2.There are metals that form more than one positive ions. In naming compounds of metals with more than one valence number, state valence with Roman number in bracket to indicate positive charge. Thus Fe^{2+} is Iron (II) (read as '' iron two '') and Fe^{3+} is iron (III) (read as "iron three")

3.2.6. Polyatomic Ions

Ions are atoms or a group of atoms that have positive or negative charges. They can be simple ions as Cl^-, O^{2-} and Al^{3+} or polyatomic ions as NH_4^+,OH^-,NO_3^-, SO_4^{2-} and PO_4^{3-}

A polyatomic ion, also called compound ion is positively or negatively charged group of atoms. The following tables give the valence number of some polyatomic ions.

Table 3.6 some common valence of polyatomic ions

Valence 1	**Valence 2**	**Valence 3**
Ammonium ion (NH_4^+)	Sulfate ion(SO_4^{2-})	Phosphate ion (PO_4^{3-})
Hydroxide ion (OH^-)	Carbonate ion(CO_3^{2-})	Phosphite ion(PO_3^{3-})
Nitrate ion (NO_3^-)	Sulfite ion (SO_3^{2-})	
Nitrite (NO_2^-)		
Hydrogen carbonate (HCO_3^-)		
Hydrogen sulfate ion (HSO_4^-)		

In writing chemical formulas of compounds that contain polyatomic ions, follow the same steps you used for writing formulas of binary compounds and use bracket if the valence number is different from 1 and not simplified.

Examples

1. Write the formula for ammonium chloride

Step 1 NH_4^+ Cl^-

Step 2 NH_4^+ Cl^-

Step 3 NH_4^+ Cl^- 

So the molecular formula for ammonium chloride is $=NH_4Cl$

2. write the formula ammonium sulfate

Step 1 NH_4^+ SO_4^{2-}

Step 2 NH_4^+ SO_4^{2-}

Step 3 NH_4^+ SO_4^{2-} 

So the molecular formula for ammonium sulfate is $(NH_4)_2SO_4$

In naming compounds containing polyatomic ions, the name of metals and ammonium ion are written first followed by the name of the polyatomic ions.

Examples: NH_4Cl (ammonium chloride), $Al_2(SO_4)_3$ (aluminum sulfate) and $FeCl_3$ Iron (III)chloride,

Exercise 3.4

1. Fill in the blank by writing the formula of a compound

Ions	Nitrate	Sulfate	Carbonate	Phosphate
Na^+	_____	_____	_____	_____
Ca^{2+}	_____	_____	_____	_____
Al^{3+}	_____	_____	_____	_____
NH_4^+	_____	_____	_____	_____
Fe^{3+}	_____	_____	_____	_____

2. Name the following compounds.

A. NH_4Cl C. $NaHCO_3$

B. $Cu(NO_3)_2$ D. $FePO_4$

3. Which three elements are combine in magnesium carbonate?

4. Which four elements are combine in ammonium sulfate?

3.2.7 Interpreting Formula

Activity 3.6 Perform the following activities.
1. What information is obtained from the coefficient and subscript in a formula?

When a formula is interpreted, it will give qualitative and quantitative meanings. Chemical symbols and formulas with numbers around them at particular positions give specific information. Thus, symbols and formulas of elements have qualitative and quantitative meaning.

Qualitatively : A symbol represents the identity (kind) of the element. A formula represents the kinds or types of element involved in forming a compound. For example:

- Cl_2 qualitatively the subscript 2 shows a chlorine molecule.
- The symbol O represents an atom of oxygen. No other element can be represented by the symbol O.
- Fe stands for iron metal
- CaO is qualitatively stands for calcium oxide made from one atom of calcium and one atom of oxygen.
- $Ca(NO_3)_2$ qualitatively stands for calcium nitrate made from one atom of calcium and two nitrate groups.

Quantitatively: a symbol represents the number of atoms of the elements. A formula stands for one molecule or for formula unit of an element or a compound. In Cl_2 quantitatively 2 shows there are two atoms in chlorine molecule

- Number preceding symbols, called coefficient, indicates the number of atoms of the element in a formula.
  2Fe stands for two atoms of iron (The number 2gives a quantitative meaning while Fe itself gives a qualitative meaning.
- A subscript written after a symbol (to the right) indicates that the element is in molecular form. For example, Cl_2 a chlorine molecule and O_2 is oxygen molecule
- The coefficient of a molecule or formula unit indicates the number of molecules or formula unit of that substance.

- $3H_2O$;the coefficient three shows that there are 3 molecules of water $2CO_2$ the coefficient 2 shows that there are 2 molecules of carbon dioxide 4NaCl the coefficient 4 shows that there are 4 formula unit of sodium chloride.

CaO quantitatively it shows one formula unit of CaO

$Ca(NO_3)_2$ quantitatively it shows one formula unit of $Ca(NO_3)_2$

Exercise 3.5

1. What does $3H_2O$ represents?

A.$3H_2O$ atoms B. 6 H molecule

C.3 H_2O molecule D. 3O molecule

2 write the qualitative meaning for

A. 2Fe B. CO_2 C. O_2

3. Write the quantitative meaning for the following.

A. $3H_2$ B. $4H_2O$ C. 2NaCl

Project Work

Write and interpret formulae of common compounds

By using reference materials, such as a Science books and/or the Internet, try to discover the formulae of common compounds such as baking soda, Vinegar (acetic acid), lime ,sugar(sucrose),chalk, milk of magnesia etc. and interpret them in terms of the elements present and the ratios of their atoms

3.3. Simple chemical Reactions and Equations

After completing this section, you will be able to

- define chemical reaction and give examples;
- describe evidences that show chemical reaction has occurred;
- state the law of conservation of mass;
- conduct an experiment in a group to show simple chemical reaction;
- write a chemical equation;
- balance simple chemical equation by inspection;
- create and use models of particles to demonstrate balanced equations.

3.3.1. Simple chemical Reaction

Activity 3.7

Discuss in groups and share your ideas with the rest of the class

1. Give some examples of chemical changes that takes place in your home or school.
2. What kind of chemical changes occurred when you coo food?
3. Imagine that you drop a glass beaker and it breaks down.
   a. Is a new substance formed?
   b. Is this a physical change or chemical change?

The starting materials in chemical reaction called reactants, react alone or with each other to produce one or more new substances, called products.

A chemical reaction involves the transformation of reactants into products.

Reactants ⟶ Products

An arrow (⟶) separates the two side and can be read as 'produce', 'give','form','yield'.

Reactants are always written on the left hand side of the arrow while product is/are written on the right hand side by putting "+" sign read as 'and' if there are two or more products. The "+" sign means "combines with" or "reacts with". For example when magnesium is in its metal form it will burn very easily in air. In burning of magnesium, the reactants are magnesium and oxygen while the product is the white ash known as magnesium oxide.

$$2Mg + O_2 \longrightarrow 2MgO$$

Similarly in the reaction between iron and sulfur, the iron and sulfur atoms are reactants whereas the formed new substance Iron sulfide is the product.

By chemical reaction, some of the common examples of changes brought about Rusting of iron, Fermentation and Digestion of food.

3.3.2 Evidences that Show Chemical Reaction has Occurred

Activity 3.8

Form a group and perform the following activity.Then present your finding to the rest of the class.

Record and describe the various chemical changes that occur in your daily lives (e.g cooking food, etc.) and describe the evidence you use to determine that chemical reaction occurred.

In a chemical reaction, new products are formed from the reactants. How can you tell this happened? There are few signs that indicate a chemical reaction has occurred. These are:

1. Color changes

Gently heating black copper oxide with sulfuric acid produce a blue solution of copper sulfate.

Copper oxide + sulfuric acid ⟶ copper sulfate + water

(Black) (Blue)

Figure 3.1: blue copper sulfate solution

2. Evolution of a gas (formation of bubbles)

When magnesium is placed in hydrochloric acid, bubble of hydrogen gas are given off.

Magnesium + hydrochloric acid ⟶ magnesium chloride + hydrogen

3.**Change of temperature (heat change):**-either endothermic or exothermic

When potassium is placed in water, hydrogen gas is given off. The reaction produces so much heat that the gas burns.

Potassium + water ⟶ potassium hydroxide + hydrogen

4. Precipitate (formation of a solid)

If you mix solutions of silver nitrate and sodium chloride, a chemical reaction takes place. In the reaction insoluble solids are formed. This is called a precipitate. The solid is silver chloride.

Silver nitrate + sodium chloride ⟶ silver chloride + sodium nitrat

Figure 3.2 white precipitate of AgCl

3.3.3. Law of Conservation of Mass

Activity 3.9

Form a group and perform the following activity. Then present your opinion to the class.

When we burn something it gets lighter or, in other words, it loses mass. For example when paper burn, the solid ash left over is lighter than the original paper. Does it mean that mass is not conserved? Discuss in groups and present your ideas to the whole class?

The law of conservation of mass states that matter is neither created nor destroyed during a chemical reaction. It means that the mass of reactants is exactly equal to the mass of the products.

3.3.4. Investigating Chemical Reaction

Expriment 3.1

Title: burning of Magnesium ribbon

Objective: to investigate the chemical reaction

Apparatus: Burner, crucible, a pair of tongs

Chemicals: Magnesium ribbon

Procedure

1. Take about 5 cm of magnesium ribbon. Rub its surface gently with an abrasive. Notice its color and hardness.
2. Hold it by a pair of tongs and burn it.

Hazards!!!

In addition to being extremely bright, burning magnesium produces some ultraviolet light; avoid looking directly at it. The burning magnesium is very hot; do not touch it or let it come in contact with other flammable materials.

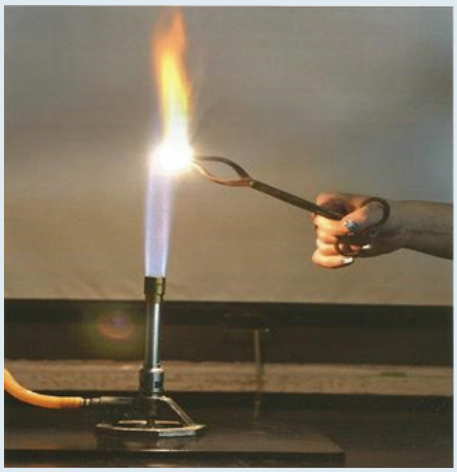


Figure 3.3: burning of magnesium

3.Collect the substance formed .Then add in a crucible and examine it carefully. Feel it. Notice its color.

Observation and analysis

1. What is the reactant materials?
2. Does it bend? It is shiny? Will it burn if heated again? Does it have any resemblance to the magnesium ribbon you started with?
3. Is chemical (change) reaction occur?

3.3.5. Writing and Balancing Simple Chemical Equation

Activity 3.10 Perform the following activity
1. What is chemical equation?

A chemical equation is a short hand expression of a chemical chanes (chemical reaction) through symbols and formulas.
In general, to write a chemical equation for a given reaction one can follow the following three steps.
Step 1: Write a word equation for the reaction.
Step 2: Change the word equation to a chemical equation i.e., write the correct symbol or formula for each reactant and product.
Step 3: Balance the equation so that it obeys the law of conservation of mass.
Example: the reaction between hydrogen and oxygen to give water.
Step1: Hydrogen + Oxygen → Water
Step 2: $H_2 + O_2 \rightarrow H_2O$
Step 3: $2H_2 + O_2 \rightarrow 2H_2O$

3.3.6. Balancing Chemical equation

Activity 3.11
Perform the following activity
Why should the chemical equation be balanced?

Chemical equation is balanced in order to obey the law of conservation of mass.So a balanced chemical equation is an equation in which the total number of atoms on the left hand side are equal to the total number of atoms on the right hand side. When we balance a Chemical equation, we have to change the coefficients not subscripts. This is because, changing subscripts changes the identity of substances. In balanced chemical equation:

Mass reactants = mass of products

Number of atoms in reactant side = Number of atoms in product side

There are many methods of balancing chemical equations. Only 2 methods of balancing chemical equations are discussed in this grade level, namely

1. The inspection method
2. Least common multiple (LCM) method

1. The inspection method- is trial and error method. It involves examining the equation and adjusting the coefficients until each kind of atoms are equal on the reactant and product sides.

For example, to balance the equation when nitrogen react with hydrogen

to give ammonia

$N_2 + H_2 \longrightarrow NH_3$

Balance nitrogen by placing 2 before ammonia

$N_2 + H_2 \longrightarrow 2NH_3$

Now you have 6 hydrogen atoms on the product side. To balance hydrogen write a coefficient 3 before H_2

$N_2 + 3\ H_2 \longrightarrow 2NH_3$

Finally check whether the equation balanced or not

Reactants	Products
N (2)	N (2)
H (6)	H (6)

Therefore, the equation is balanced

2. Least common multiple (LCM) method

The steps used in this method are shown by the following examples. Consider the reaction between aluminum and oxygen to form aluminum oxide.

Step 1: Represent the reaction by word equation

Aluminum + oxygen ⟶ aluminum oxide

Step 2: Write the correct formula for each of the reactants and products

$Al + O_2 \longrightarrow Al_2O_3$

Step 3: Find the total valence number (subscript time's valence number) and place above each symbol and formula

$$\overset{3}{Al} + \overset{4}{O_2} \longrightarrow \overset{6}{Al_2}\overset{6}{O_3}$$

Step: 4 Find the LCM of total valence number and place it above the arrow. The LCM of 3, 4, and 6 is 12

$$\overset{3}{Al} + \overset{4}{O_2} \xrightarrow{12} \overset{6}{Al_2}\overset{6}{O_3}$$

Step 5: Divide the LCM by each total valence number to obtain the coefficients for each of the reactants and products. Place the coefficients thus obtained in front of the respective formula.

$4Al + 3O_2 \longrightarrow 2\,Al_2O_3$

Check

Reactants	Products
Al (4)	Al (4
O (6)	O (6)

So the equation is balanced

Exercise 3.6

1. Balance the following by inspection

   a. $CaCO_3\,(s) \longrightarrow CaO(s) + CO_2(g)$
   b. $H_2 + I_2 \longrightarrow HI$
   c. $C_2H_2(g)+O_2(g) \longrightarrow CO_2(g)+H_2O(l)$
   d. $Ca + H_2O \longrightarrow Ca(OH)_2 + H_2$
   e. $Fe_2O_3 + CO \longrightarrow Fe + CO_2$

2. Balance the following chemical equations

   a. $Fe + O_2 \longrightarrow Fe_2O_3$
   b. $Cu + H_2SO_4 \longrightarrow CuSO_4 + SO_2 + H_2O$

3.4 Use of Chemical Reactions in Every Day Situation

After completing this section, you will be able to:

- describe the uses of chemical reactions in everyday situations.

Activity 3.12

Discuss in groups and share your ideas with the rest of the class.

1. How do the local people in Ethiopia prepare alcoholic beverages like "Tella"? What raw materials are used? Is the process a chemical change?
2. Give some examples of useful chemical reactions such as fermentation in brewing which produces carbon dioxide and ethanol/

3.4.1. Uses of Chemical Reaction

Chemical reaction happen everywhere. It happen inside your body to keep you alive. For example, reactions to digest food i.e. the breakdown of large molecules (protein, starch and fats) into smaller ones, so that they can be absorbed. Chemical reactions are an integral part of technology of culture, and indeed of life itself. Burning fuels, smelting iron, making glass and pottery, brewing beer, and making wine and cheese are among many examples of activities incorporating chemical reactions that have been known.

Some Important Chemical Reactions

1.Synthesis of ammonia:

Reaction: $3H_2(g) + N_2(g) \longrightarrow 2NH_3(g)$

Hydrogen gas and nitrogen gas are combined in the presence of a catalyst at high temperature and pressure to produce ammonia gas.

Significance: Synthesis of ammonia leads to the production of fertilizer (ammonium nitrate) and to the production of ammunitions.

2. Combustion of hydrogen

$2H_2(g) + O_2(g) \longrightarrow 2H_2O(l)$

Hydrogen gas and oxygen combine to produce liquid water.

Significance: In the forward direction this is a spontaneous reaction that explosively oxidized hydrogen to water.

3. Combustion of methane (hydrocarbons)

Reaction:$CH_4(g) + 2O_2(g) \longrightarrow CO_2(g) + 2H_2O(g)$

Methane gas and oxygen gas combine exothermically to produce carbon dioxide gas and water vapor.

Significance: Methane is the simplest of the hydrocarbons, all of which combine with oxygen and undergo oxidation. If the oxidation is complete the products are carbon dioxide (a greenhouse gas) and water.

4. Photosynthesis

Reaction: $6CO_2(g) + 6H_2O(l) \longrightarrow 6O_2(g) + C_6H_{12}O_6(aq)$

Carbon dioxide and water combine in the presence of sunlight (and many biologically catalyzed reactions) to produce oxygen and glucose (sugar). Significance: Our atmosphere is 21% oxygen - in spite of the tendency of oxygen to react with so many substances. The constant level of oxygen is maintained by the many plants that inhabit our planet through the reaction of photosynthesis. This is truly solar energy at its most efficient and productive.

5. Synthesis of sulfuric acid

Reaction: $2S(s) + 3O_2(g) + H_2O(l) \longrightarrow 2H_2SO_4(aq)$

Sulfur is first oxidized to sulfur dioxide and then to sulfur trioxide. This gas is bubbled through water to produce sulfuric acid.

Significance: Sulfuric acid is a very important chemical and an indicator of a nation's industrial strength.

Key terms

Element	Balanced Chemical Equation
Reactants	Chemical Symbol
Valence	Products
Chemical Formula	Binary Compounds
Chemical Equation	Molecule
Poly Atomic Ion	Law Of Conservation Of Mass
Chemical Reaction	Inspection Method
Lcm Method	

Unit summary

An atomic symbol is defined as shorthand way of representing elements or atoms of an element.

A compound is a pure substance consists of two or more elements which have been chemically combined.

Chemical formula is the symbolic representation of an element or a compound.

A molecule is the smallest particle of an element or a compound that has a stable, independent existence.

The combining power of an element is **called valence.** Most common elements have valence 1, 2, or 3.

Binary compounds are compounds formed from two different types of elements. In naming a compound, the positive ion (metal) mention first followed by the negative ion (nonmetal).

A polyatomic ion, also called compound ion is positively or negatively charged group of atoms.

A chemical reaction is a process in which some substances is changed into one or more different new substances. The starting materials in chemical reaction called **reactants,** react alone or with each other to produce one or more new substances, called **products.** There are few signs that indicate a chemical reaction has occurred. These are: color change, evolution of gas, heat change and formation of precipitate.

The law of conservation of mass states that matter is neither created nor destroyed during a chemical reaction.

A chemical equation is shorthand expression of a chemical changes (chemical reaction) through symbols and formulas.

There are many methods of balancing chemical equations. Some of them are inspection method, Least common multiple (LCM) method.

Review Exercise

I Write 'True' for the Correct Statements and 'False' for the Wrong Statements.

1. A compound is pure substance.
2. In a chemical reaction atoms are neither created nor destroyed.
3. O_2 and 2O have the same meaning.
4. The symbol copper denoted by Co.
5. When we balance chemical equation, we change the subscript but not the coefficient.
6. Respiration is a chemical change (chemical reaction).
7. The combining power of an element is called valence.
8. The formula and symbol of nitrogen is the same.
9. A number in front of a symbol or formula is coefficient
10. A chemical symbol is a short hand notation for the chemical name of an element.

II.Choose the correct answer from the given alternatives.

11. The Latin name of silver is
    A. Argentum B. Kalium C. Natrium D. Cuprum
12. In $4O_3$ the coefficient and subscript respectively
    A.3, 4 B.4,3 C.7,3 D.3,7
13. The formula of nitrate ion is
    A. NO B. NO_3^- C. NO_2^- D. N_2^-
14. All of the following elements can exist as diatomic molecules EXCEPT
    A Hydrogen B Oxygen C Sodium D Chlorine

15. What is the chemical formula for Iron (III) chloride?

A. $FeCl_2$ B.FeCl C.$FeCl_3$ D. Fe_2Cl_3

16. The chemical symbol for Helium is

A. H B. Hl C. He D.HE

17. Which of the following is the correct name of MgO?

A. Magnesium oxygen B. oxygen magnesium

C. Magnesium oxide D. molybdenum oxide

III. Give short answers

18. Write the symbol of

a. Zinc b. phosphorous c. Mercury d. calcium e. Tin

19. Balance the following chemical equations

a. $Al(OH)_3 \longrightarrow Al_2O_3 + H_2O$

b. $NH_3 + O_2 \longrightarrow NO + H_2O$

c. $Ca(OH)_2 + H_3PO_4 \longrightarrow Ca_3(PO_4)_2 + H_2O$

20. Write chemical equations for the following reactions and balance them.

a. Aluminum + bromine $\longrightarrow$ Aluminum bromide

b. Sodium + Nitrogen $\longrightarrow$ Sodium nitride

c. Potassium + Water $\longrightarrow$ Potassium hydroxide + Hydrogen

Unit 4 CELLS AS THE BASIS OF LIFE

Learning Outcomes

At the end of this unit, you will be able to:

- define a microscope;
- explain the use of a microscope;
- distinguish the different types of microscopes;
- describe the basic parts and functions of a microscope;
- use a microscope to view objects;
- define a cell;
- explain how cell was discovered and who discovered it;
- draw a cell and label its major parts;
- describe the functions of the major structural parts of a cell;
- distinguish between unicellular and multicellular organisms;
- give examples of cell shape;
- explain why cell shape and structure vary;
- discuss the differences of cell, tissue, organ and organ system;
- define respiration and write its chemical equation;
- define photosynthesis and write its chemical equation.

Main contents

Introduction

This unit deals with cell as the basis of life and organized in to two sub units. The first one deals with the purpose and types of microscope. The second part deals about the cell, this section focuses on discover and definition of cell, structures of cell and their function, types of organism, and level of organization of organism. A cell is the smallest unit of life. Most cells are so small that they cannot be viewed with the naked eye. Therefore, scientists must use microscopes (magnifying) instrument to study cells.

4.1. Microscope

At the end of this section, you will be able to:

- identify the major parts and functions of a basic microscope;
- use a microscope to view objects;
- discuss the role of a microscope;
- differentiate between simple and light microscope;
- draw diagram of a microscope and label the major parts;
- build microscope from locally available materials.

Microscope have opened up a whole new dimension in science, by using a microscope scientists were able to discover the existence of the microorganisms, study the structure of cell, and see the smallest parts of plants, animals, and fungi. Cells are the smallest units from which all life forms are made.

Activity 4.1: discuss in group and share your ideas

What kind of organisms found in your environment?
How can you observe those can not see by your naked eye?

4.1.1 Purpose and Invention of a Microscope

What is a microscope?

A microscope is an instrument that is used to observe objects which Can not see clearly with naked eye. A microscope uses lenses or system of lenses to produce a magnified image of an object under study. Microscopic means invisible to the eye unless aided by a microscope. The science of investigating small objects using such an instrument is called microscopy.

Word Roots and Origins

The word microscope is derived from two Greek words
"Micro" meaning tiny and "scope" meaning to view or look at

What is the use of microscope?

Microscope magnifies the size of the object observed so that it looks bigger than its actual size. This offers a chance to closely study and learn more about smaller organisms like cell and microorganism.

Who invented a compound microscope and when?

Dutch spectacle makers Zaccharias Janssen and Hans Lipperhey (1595) are noted as the first men to develop the concept of the compound microscope by placing different types and sizes of lenses in opposite ends of tubes.

In 1665 Robert Hooke an English scientists built compound microscopes, which have multiple lenses. However, his microscope is a compound microscope, the lenses are not very good and magnifications of more than 30x are very blurred and do not show much detail. Later in the 1674 century, Anton van Leeuwenhoek Dutch merchant began polishing and grinding lenses when he discovered that certain shaped lenses increased an image's size. The glass lenses that he created could enlarge an object many times. The quality of his lenses allowed him, for the first in history, to see the many microscopic animals, bacteria and intricate detail of common objects. Leeuwenhoek is considered the founder of the study of microscopy and played a vital role in the development of cell theory.

4.1.2. Types of a Microscope

There are many types of microscopes, and they may be grouped in different ways. According to the type of radiation they use for observation, microscopes grouped into two main types of microscopes these are the **light microscope** and **electron microscope**. The light microscope uses a beam light to form the image of an object, while the electron microscope uses the beam of electron to form the image. Based on the number of lenses it has and uses the light microscopes are categorized into two ***simple microscope*** and c***ompound microscope.***

Simple light microscope

A simple microscope consists of a single convex lens that is capable of magnifying an object. A hand lens (magnifying glass) and reading lens an example of simple microscope. They can magnify about ten times (10X) to twenty (20X).

Figure: 4.1 types of simple light microscope

Compound light microscope

Compound microscope is a microscope that uses multiple lens systems at the same time to improve magnification and resolution. The two lens systems are the eyepiece (ocular) lens and the objective lenses. The objective lenses include:

- Lower power objective (4x)
- Middle power objective (10x)
- High power objectives (40)
- Oil immersion lenses (100x)

Microscope has two major abilities: magnification and resolution

1. **Magnification** is increasing the size of an object to be viewed.
2. **Resolution** is ability of the microscope to show the detailed or the scattered part of an object. It helps us to distinguish between two separate points

Key Terms

Microscope: an optical instrument used to observe very small objects.
Microscopic: very small objects which are only viewed with microscope.
Magnifications: increasing the image of an object
Lens: a piece of glass used to converge or diverge light and form optical images.
Resolution is ability of the microscope to show the detailed or the scattered part of an object.

Key Terms

Monocular compound microscope: A compound microscope with single eyepiece lens.
Binocular compound microscope: compound microscope with two eye piece lens

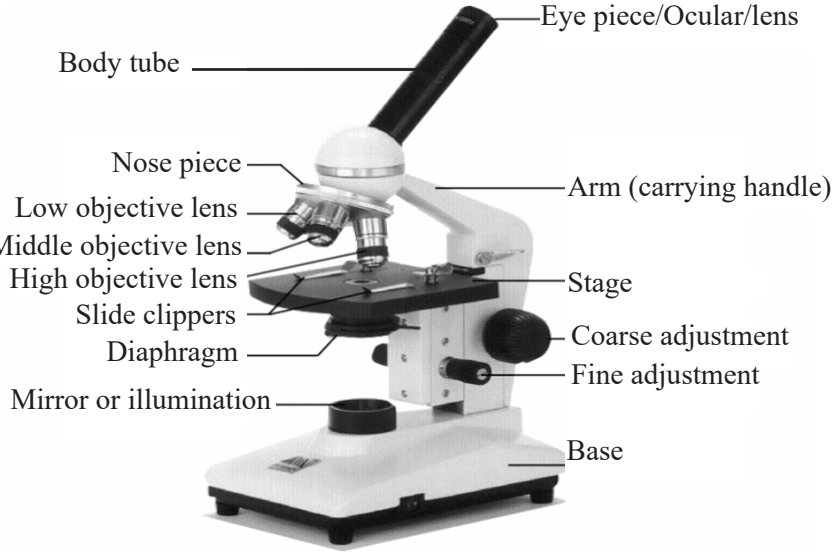


Figure 4.2 Parts of compound microscope

N.B: Your teacher will provide you with a microscope so that you can identify the parts and their functions

4.1.3. Basic parts of compound light microscope

Table 4.1 parts of compound microscope and their function.

Parts	Function
Base	Support the microscope
Arm	Used to carry the microscope
Stages	Supports the glass slide and contains the specimen being Observed.
Stage clips	Holds the slide in place on the stages
Eyepiece	Magnifies image for the viewer
Objective lens	Low, medium and high power lenses that magnifies the specimen
Course adjustment	Large knob used for focusing the images under low power. produce rough focus of image

Fine adjustment	Smaller knob used for focusing the image with high power objectives. Sharp focus image
Diaphragm	Controls the amount of light that pass through the specimen
Light course	Provide light for viewing the specimen
Body tube	Separates the objective and the eyepiece and assures continuous alignment of the optics.

Activity 4.2:
Based on figure
4.3 label each parts of a microscope and relates your answare to the actual parts of microscope.

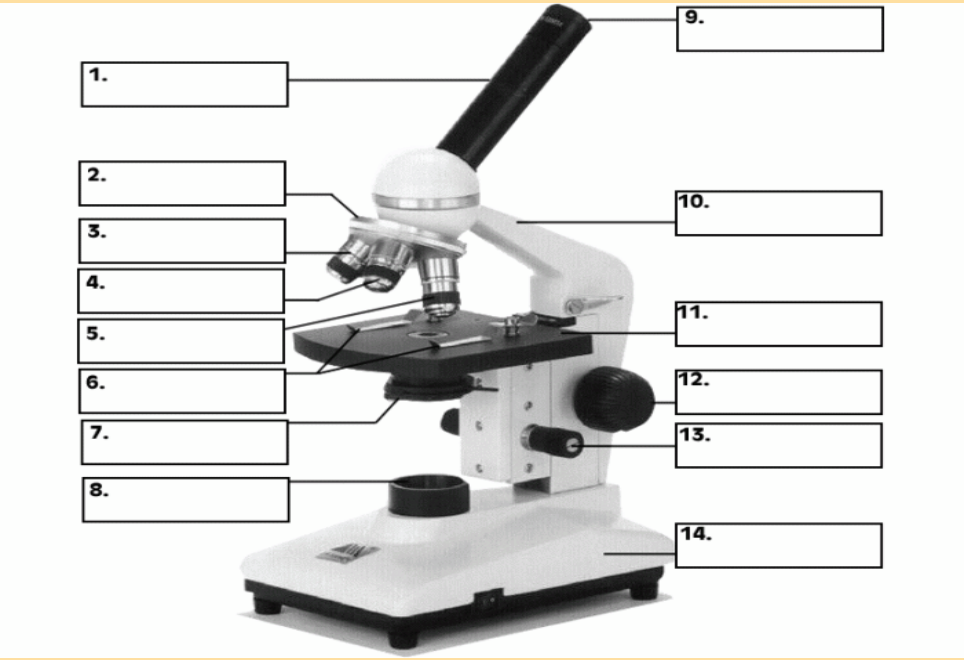


Figure 4.3 Parts of microscope

Magnification Power of Compound Microscope

Activity 4.3 work in a group

- What component of a microscope is used to magnify an object and in which parts of a microscope it is located?
- Discuss the total magnification of a compound microscope.

The eyepiece lens usually magnifies ten times and is labeled 10X. The objective lenses magnify four to hundred times. The total magnification of an object is calculated by multiplying the magnification of the objective lens by the magnification of the ocular lens. For example, if the magnification of the eye lens is 10X and the magnification of the objective lens is 4X, then the total magnification is 40X. Because two lenses. are used, compound microscopes are capable of higher magnifications than simple microscopes, which use only one lens.

key terms

Mounting: is preparing a specimen for observation under a microscope.

Focusing: is adjustment of focus to observe specimen clearly.

Specimen: a sample of a substance or material for examination or study

Experiment 4.1

Practicing mounting and focusing

Materials you require:

- Clean slide and cover slip
- Very fine fiber
- Dropper with nipple
- Forceps
- Water in a beaker
- Compound light microscope

Procedure:

1.Lay down the very fine fiber on a clean microscope slide as shown in the figure 4.4

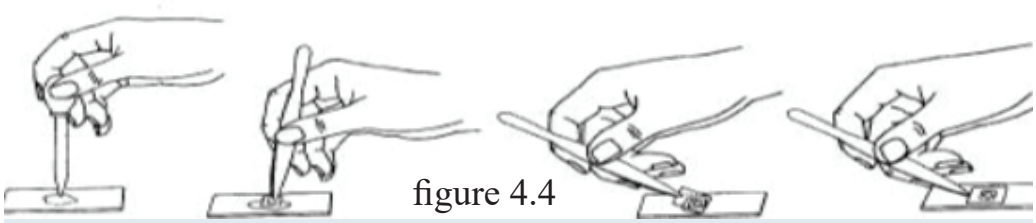


figure 4.4

2. Add a drop of water to a slide | Place the specimen in the water | Place the edge of a coverslip on the slide so that it touches the edge of the water | slowly lower the coverslip to prevent forming and trapping air bubble

- Place one drop of water directly over the specimen and cover it with a cover slip.
- If you put too much water over the specimen, cover slip will float on top of the water.
- Air bubbles confuse the observer.
- Cover slips protect objective lens and keep the specimen in position

3. Place the slide on the microscope stage, with the specimen directly over the center of the glass circle on the stage (directly over the light).
4. Always start and end with Low Power objective. Change objective lens to the lowest point, then focus using first the coarse knob, then with the fine focus knob.
5. Adjust the Diaphragm as you look through the Eyepiece, and you will see that more detail is visible when you allow in less light! Too much light will give the specimen a washed-out appearance.Try it out!!
6. Use low power, then, without changing the focus knobs, switch it to medium power. Move the slide until you are able to see clearly through the lens.
7. Use the Medium and High Power remember that you only use the fine focus knob!
8. Click the high power objective lens in position and only use the fine adjustment knob to focus on specimen. At this point, if the specimen is too light or too dark, try adjusting the diaphragm.
9 . Then, focus using the fine adjustment for sharp focusing. Do not use the coarse adjustment

11. Explain what change you have observed. Is the fine fiber compact or relaxed? Draw it.

Activity 4.4 Project work

Construct a model of light microscope from locally available materials in group of five students and submit it to your teacher.

Exercise 4.1

Choose the best answer from the given alternatives

1. Of the followings lists choose the correct order in which light passes through it.
   A. mirror----objective------ lens----eyepiece --- lens
   B. mirror-----slide----------objective lens-----eye piece
   C. Lens-------slide----------eyepiece lens-------objective lens
   D. Eye piece-----objective lens----slide------mirror
2. Which parts of a microscope are combined to give magnified view of specimen?
   A. Light source and objective lens
   B. eye piece and objective lens
   C. Stages and eyepiece
   D. eye piece and focus knob
3. To focus on specimen is the best to start with which objective lens?
   A. lower magnification C. intermediate magnification
   B. High magnification D. Oil emersion
4. To which part of a microscope do you look through to see an object magnified?
   A. Eyepiece B. Stage C. Focus knob D. objective lens
5. Which One of the following is not the function of a microscope.
   A. Magnifying the image of the sample.
   B. Showing the details of the sample.
   C. Enabling one to observe something seen with naked eye.
   D. Enlarging the size of the sample.

6. Magnification power of a microscope is related to its ability to:
   A. increase the size of the image.
   B. show the fine details of the sample.
   C. resolve the image.
   D. all of the above

4.2. Cell

At the end of this section, you will be able to:

- explain how cell was discovered;
- draw and label the basic structures and functions of a cell;
- explain why cell shape and structure vary;
- distinguish between unicellular and multicellular organisms;
- differentiate among cell, tissue, organ and organ system with examples;
- examine the importance of cellular respiration and photosynthesis.

Introduction

In the previous section you have already learnt about a microscope; the instrument that magnifies the images of an object. This enables scientists to look at and study smaller things like cells which are not seen by naked eye. In this sub unit, you will learn about the basic structural unit of life, which is the cell; cell may be compared to bricks. Bricks are assembled to make a house, similarly cells assembled to make the body of organism.

Word Roots and Origins

Word cell comes from the Latin **"cellula"** meaning a small room.

4.2.1 The Discovery and the Definition of Cell

Many biologists and other scientists contributed to the discovery of cells. Among this the English scientists Robert Hooke (1665) was he first to use the cell for he observed at very tiny slice of cork through his microscope. He noticed that the cork was porous and comprised of many tiny square boxes that remained him of the small rooms in monastery. Hence, the word cell comes from the Latin cellula meaning "a small room".

Soon after Robert Hooke discovered cells in cork, **Anton van Leeuwenhoek (1674)** made his own simple microscope with only one lens. However, van Leeuwenhoek is very skilled at grinding lenses and so his microscope can achieve magnifications of 300X. He was the first person sees living, moving unicellular organisms (Protista) in a drop of water. He calls the moving organisms 'animalcules'. He also sees bacteria (from his teeth), which he also calls 'tiny animalcules'.

By the late 1830s, botanist **Matthias Schleiden** and zoologist Theodor Schwann were studying tissues and proposed the unified cell theory, which states that:

- all living things are composed of one or more cells,
- the cell is the basic unit of life, and
- all new cells arise from pre existing cells.

What is a cell?

Cells are the basic units of living organisms which are responsible to carry out basic structural, functional, and biological unit of all known organisms.

Cells are the building blocks of all living beings provide structure to the body. Organisms may be made up of a single cell or many cells. Cells are complex and their components perform various functions in an organism. It comprises several cell organelles that perform specialized functions to carry out life processes.

4.2.2. Structure of a cell

Cells are the tiny structural units of life, are made up of different parts. The parts of cells are known as sub-cellular structures or organelles. Different sub-cellular structures carry out different functions in cells. When observed under compound or electron microscope, all cells share four common components:

1) **A plasma membrane,** an outer covering that separates the cell's interior from its surrounding environment;

2) Cytoplasm, consisting of a jelly-like region within the cell in which other cellular components (organelles) are found

3) DNA, the genetic material of the cell and

4) Ribosomes, particles that synthesize proteins. Animal and plant cells share some common features like the cell membrane, nucleus and cytoplasm.

There are several types of organelles within organism cell. The sub cellular structures found in cells are cell wall, cell membrane, mitochondria, plastids, ribosomes, nucleus, Golgi apparatus, lysozyme, endoplasmic reticulum, vacuoles and others. Organelles

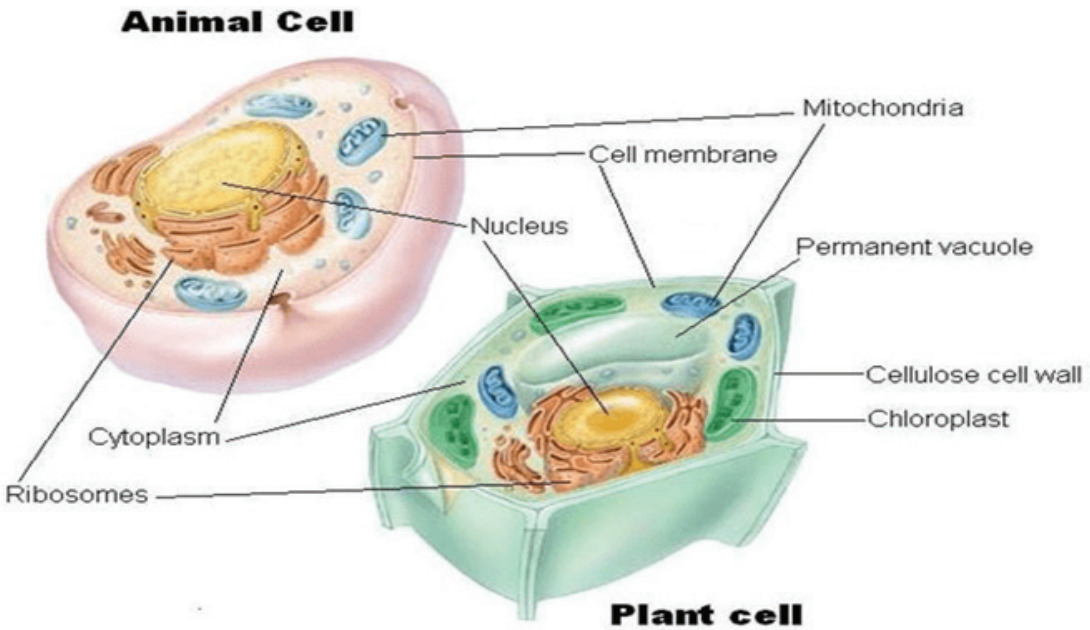


Figure 4.5: Basic structure of animal and plant cell

Table 4.2 summary of the common and difference parts of animal and plant cell

	Name of part	Description	Function
Animal and plant cell	Cytoplasm	jelly-like, with particles and organelles in	contains the cell organelles, e.g. mitochondria, nucleus, site of chemical reactions
	cell membrane	a partially permeable layer that forms a boundary around the cytoplasm	prevents cell contents from escaping controls what substances enter and leave the cell
	nucleus	a circular or oval structure containing DNA (genetic material)	controls cell activities controls cell division controls cell development
Plant cells only	cell wall	a tough, non-living layer made of cellulose surrounding the cell membrane	prevents plant cells from bursting allows water and salts to pass through (freely permeable)
	Vacuole	a fluid-filled space surrounded by a membrane	contains salts and sugars helps to keep plant cells firm
	Chloroplast	an organelle containing chlorophyll	• Site of photosynthesis • traps light energy for photosynthesis

Exercise 4.2

1. Compare and contrast animal cell with plant cell?
2. What are the comen components of all cell?

Structures and functions of organelles

Cell wall: The outer covering of cells that surrounds the cell membrane in plant cell. The cell wall is a rigid covering that protects the cell.

Cell membrane: Is the outermost covering of the cell that separates the content of the cell from its external environment. It Controls materials that get in and out of the cell.

Nucleus: controls all activities of the cell.

Cytoplasm: it contains different sub-cellular structures in which chemical processes take place. The part of the cytoplasm that does not contain any organelles is referred to as the cytosol.

Mitochondrion: Known as the powerhouse of the cell, the mitochondrion (plural: mitochondria) is the double-membrane organelle where the process of cellular respiration takes place.

Chloroplast: Specific/unique to plant cells, chloroplasts are double-membrane organelles and serve as site of photosynthesis.

Ribosome: Ribosomes are the sites where protein synthesis occurs. Because protein synthesis is essential for all cells, ribosomes are found in almost in every cell,

Endoplasmic Reticulum: is a series of interconnected membranous tubules that collectively modify proteins and synthesize lipids. Most cells contain two types of endoplasmic reticulum: the rough and the smooth. **The rough endoplasmic reticulum** (RER) is so named because the ribosomes attached to its cytoplasmic surface give it a studded appearance when viewed. Protein molecules undergo modifications such as folding or addition of sugars. **The smooth endoplasmic reticulum** (SER) is continuous with the RER but has few or no ribosomes on its cytoplasmic surface.

The SER's functions include synthesis of carbohydrates, lipids (including phospholipids), and steroid hormones; detoxification of medications and poisons; alcohol metabolism; and storage of calcium ion.

The Golgi apparatus is a series of flattened membranous sacs. The sorting, tagging, packaging, and distribution of lipids and proteins take place in the **Golgi apparatus** (also called the Golgi body).

Lysosomes

In animal cells, the lysosomes are the cell's "garbage disposal." Digestive enzymes within the lysosomes aid the breakdown of proteins,polysaccharides, lipids, nucleic acids, and even worn-out oranelles. In single-celled organism lysosomes are important for digestion of the food they ingest and the recycling of organelles.

Vesicles and Vacuoles

Vesicles and vacuoles are membrane-bound sacs that function as storage and transport materials. The central vacuole in plant cell plays a key role in regulating the cell's concentration of water in changing environmental conditions.

Peroxisomes

are small, round organelles enclosed by single membranes. They carry out oxidation reactions that break down fatty acids and amino acids. They also detoxify many poisons that may enter the body. Alcohol is detoxified by peroxisomes in liver cells.

4.2.3. Cell Shape and Size

Different cells have different shapes and their unique morphologies are directly related to their function:

- Plant cells, in general, have rectangular, rigid walls, and distinct edges. Such structure is contributed by the presence of cell wall that forces the cell to have a definite shape.
- Unlike plant cells, animal cells tend to have more irregular body shapes due to the absence of cell wall in their overall structure.
- Microorganisms like bacteria have three types of cell shape: oval (cocci), rod-shaped (bacilli), spiral, star-shaped, and rectangular.

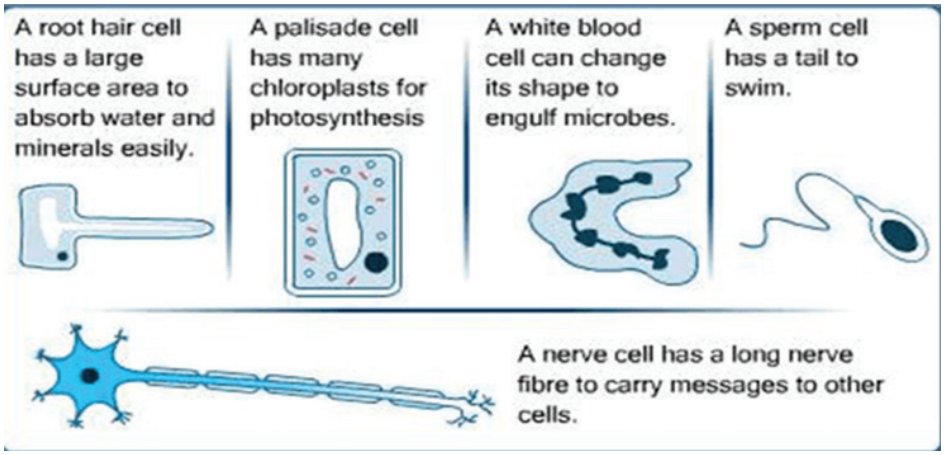


Figure: 4.6 the shapes and sizes of different cell

Like shapes, the size of cells is also linked to their functions. Depending on the type of organism, the size of the cell greatly varies.

- In particular, egg cells are the largest cells that an organism has. This is very much related to their function as the development of the zygote after fertilization requires huge amounts of energy. Approximately, the human egg cell measures 0.12 mm in diameter.
- On the other hand, the smallest cell is that of the parasitic bacterium Mycoplasma gallicepticum. This bacterium, which thrives in the bladder, respiratory and reproductive tracts of mammals. This cell has an average diameter of 0.0001 mm

4.2.4 Unicellular Organisms

What is a unicellular organism?

A cell is the basic unit of life. All living organisms are composed of one (unicellular) or more (multicellular) cells.

Unicellular organisms are those organisms composed of one cell. They are typically microscopic in nature and cannot be seen with naked eyes.Life processes such as excretion, digestion, feeding and reproduction occur in one cell. Examples of unicellular organisms include different bacteria, most algae, unicellular fungi (yeast) and protozoans such as, amoeba and paramecium

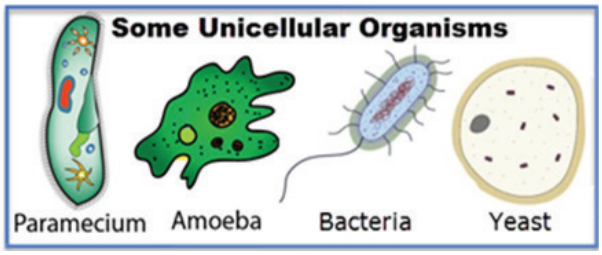


Figure: 4.7 some unicellular organisms

Experiment 4.2

Examining different water samples and prepared slides for the presence of organisms In the first part of this activity, you will examine prepared slides of unicellular organisms under a compound microscope. in the second part of this activity; you will prepare a wet mount of amoeba culture.

Materials you will require:Hand lens,Water samples from pond, river, lake, well or standing water, Compound light microscope, Prepared

Procedures:

I Observation using a hand lens

1. Collect water samples in open mouth container (beakers).
2. Observe the surface of each sample with a hand lens for the presence of living organisms.
3. Can you observe anything moving? Please, draw it.

II Observation of prepared slides

1.Place a prepared slide of amoeba on the stage of the microscope.

1. With your microscope on low power, observe the slide.
2. Move the slide around on the stage until you find some cells.
3. Now, using the medium- or high-power objective lens, focus on one cell and observe and draw what you see.
4. Label all visible structures.
5. Repeat steps 1- 5 for prepared slides of Paramecium, euglena, yeast and bacteria.
6. For each organism you view, be sure to include the name of the organism and the total magnification used.

After completing these practical activities, compare your drawings with the figures given on; figure 4.7

4.2.5 Multicellular organisms:

are organisms that are made of up many cells. Plants and animals are examples of multi cellular organism. In multi cellular organism cell are specialize to perform different functions.Human being are multicellular organism consist different types of cell like blood cells, skin cells, brain cells, heart cells and many types of other cells. Similarly plants have different cell like stem Cells, root cells, and many other types.

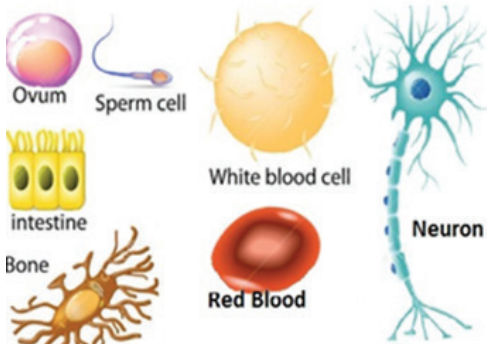


Figure: 4.8 some different types of cells that build up human body

Cells are the basic building blocks of living things. The human body is composed of trillions of cells, all with their own specialized function.

4.2.6 Levels of Organization of Living Things

Cell, tissue, organ, organ system, organisms

1. Cell: All living things are made of cells; the cell itself is the smallest important unit of structure and function in living organisms it performs various metabolic functions like providing structure and rigidity to the body, converting food into nutrients and energy, and others.

Regardless of their small size, cells are organized in a precise manner.

At cellular level, organisms can be classified into two: single-celled organisms (unicellular) and multiple-celled organisms (multi-cellular)

2. Tissue: In most multicellular organisms, cells combine to make tissues, which are groups of similar cells carrying out the same function. For example, muscle tissue, connective tissue, and nervous tissue. Like cells, tissues perform metabolic processes that keep the organism alive.

3. Organs: are collections of tissues grouped together based on a common function. Organs are present not only in animals but also in plants.In plants, their organs include the flowers, roots, stems, and the leaves. On the other hand, organs of animals include the brain, heart, stomach, eyes, and many more

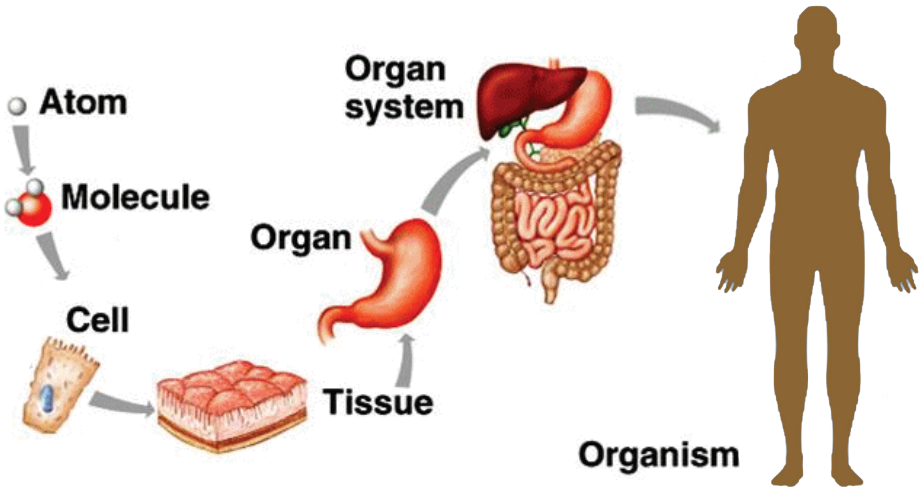


Figure: 4.9 Level of organization in human being

4. Organ system: is a higher level of organization that consists of functionally related (associated) organs. While each organ system in an organism works as a distinct entity, they all function in cooperation with each other in order to help keep the organism alive. In plants, organ systems include the root and shoot system, while animal organ systems include the *digestive system, nervous system circulatory system, and others.*

5. An organism can be simply defined as any living thing that is composed of various organ systems that function altogether. **Organisms** are individual living bodies.

Exercise 4.3

Choose the best answer for each of the following questions.

1. Which one of the following is true about cells?
   A. Cells are generally too small to be seen with the naked eye.
   B. Cells are the structural and functional units of all life forms.
   C. Cells are the smallest units that carry out all process of life.
   D. All of the above
2. Which one of the following is part of a cell and only found in animal cell?
   A. Cell membrane
   B. Chloroplast
   C. Cell wall
   D. Lysosome
3. Which one of the followings is a unicellular organism?
   A. amoeba B. yeast C. bacteria D. Paramecium E. all
4. Which one the following is not an organ?
   A. heart B. lung C. nerve D. stomach
5. Which one is formed from groups of similar cells carrying out the same function?
   A. Organ B. tissue C. organ system D. Organism

4.2.7 Respiration and Mitochondria.

Activity 4.5: Discuss in a group

- How is the energy in food we eat released to our body?
- What is the role of oxygen that we inhale during breathing

Respiration is the process in which both plants and animals break down simple sugars into carbon dioxide and water and release energy in the form of adenosine triphosphate (ATP). The ATP is used for all the processes that occur within a cell that need energy.

Cellular respiration occurs in the mitochondria of all organisms. **Mitochondria** (singular, mitochondrion) are often called the "powerhouses" or "energy factories" of a cell because they are responsible for making adenosine triphosphate (ATP), the cell's main energy-carrying molecule. Mitochondria are oval-shaped, double-membrane organelles. Each membrane is a phospholipid bilayer embedded with proteins. The inner layer has folds called **cristae**, which increase the surface area of the inner membrane. The area surrounded by the folds is called the mitochondrial **matrix**. The cristae and the matrix have different roles in cellular respiration.

The formation of ATP from the breakdown of glucose is known as **cellular respiration** and occurs in two ways, **aerobic respiration** and **anaerobic respiration.**

Aerobic Respiration

Cellular respiration or aerobic respiration is a series of chemical reactions which begin with the reactants of sugar in the presence of oxygen to produce energy, carbon dioxide and water as waste products.

Cellular respiration has four stages

1. Glycolysis.
2. Oxidation of pyruvates or Link reaction
3. Tricarboxylic Acid (TCA) or Citric Acid cycle
4. Electron transport chain (ETC) and chemiosmosis

In total, the resulting product of aerobic cellular respiration from a single glucose molecule can be up to 38 ATP.

$C_6H_{12}O_6 + 6\ O_2 \rightarrow 6\ CO_2 + 6\ H_2O + 38ATP$

(Glucose + 6 Oxygen → 6 Carbon Dioxide + 6 Water + 38ATP)

What is the role of Oxygen in cellular respiration?

Oxygen is an essential molecule in cellular respiration. Basically, oxygen can be found at the end of the ETC (during aerobic respiration) where it accepts electrons while picking up protons in order to produce water molecules.

Because of this, oxygen is also called as the "final electron acceptor" / Oxidizing agent. When oxygen levels are depleted, electrons will be simply dispersed and the electron transport chain will discontinue.

Anaerobic Respiration: Respiration in animal cells can take place anaerobically (without oxygen), to transfer energy; it simply involves the incomplete breakdown of glucose into lactic acid. This occurs when the body can't supply enough oxygen for aerobic respiration, such as during vigorous exercise

• In animal cells, this process is called the **lactic acid fermentation**. It can be simplified in the following equation:

$C_6H_{12}O_6 \rightarrow 2\ CH_3CH\ (OH)\ COOH + 2\ CO_2 + 2\ ATP$

Glucose → lactic acid + carbon dioxides+ 2ATP

• On the other hand, microorganisms like yeast respire without oxygen produce ethanol and carbon dioxide. Such process is referred to as the **ethanol or alcohol fermentation.**

$C_6H_{12}O_6 \rightarrow 2\ C_2H_5OH + 2\ CO_2 + 2\ ATP$

Glucose→ Ethanol + carbon dioxides + 2ATP

In both types of fermentation process, only 2 ATP are produced from a glucose molecule

4.2.8 Photosynthesis and chloroplast

Plants make sugar by using energy from sunlight to change carbon dioxide (CO_2), a gas absorbed from the air, and water (H_20) taken from the ground by roots into glucose ($C_6H_{12}O_6$) and oxygen (O_2). This process is called **photosynthesis** and occurs in the **chloroplast** of the plant cell.

Chloroplasts have outer and inner membranes, within the space enclosed by a chloroplast's inner membrane; is a set of interconnected and stacked, fluid-filled membrane sacs called thylakoids. Each stack of **thylakoids** is called a **granum** (plural grana). The fluid enclosed by the inner membrane and surrounding the grana is called the **stroma**.

The chloroplasts contain a green pigment called **chlorophyll**, which captures the energy of sunlight for photosynthesis.
Photosynthesis is a series of chemical reactions that convert carbon dioxide and water into glucose (sugar) and oxygen in the presence of sunlight.

Carbon dioxide + Water $\xrightarrow[chlorophll]{light}$ Glucose (sugar) + Oxygen+ Water

$6CO_2 + 12H_2O \quad C_6H_{12}O_6 + 6O_2 + 6H_2O$

Importance of photosynthesis and respiration

The process of photosynthesis plays important roles. Photosynthesis converts radiant or solar energy into chemical energy. It provides oxygen in atmosphere for all living organism to perform cellular respiration. It maintains the balanced level of oxygen and carbon dioxide ecosystem.

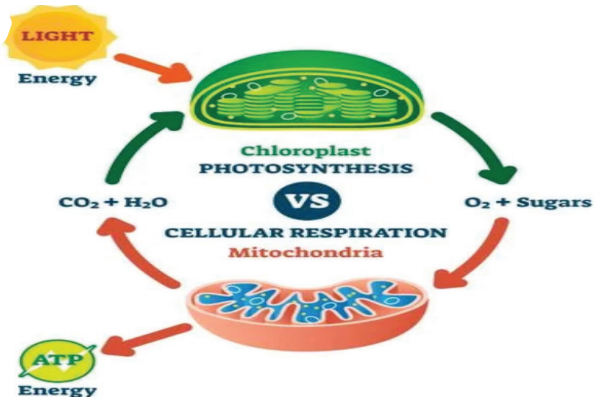


Figure 9.10 the relationship between respiration and respiration

On the other hand carbon dioxide produced by organism from cellular respiration would be used during photosynthesis. The following diagram illustrates the important relationship between cellular respiration and photosynthesis.

Exercise 4.4

1. Compare the functions of the mitochondrion and chloroplast. How are they similar?
2. Write the chemical equations of cellular respiration and photosynthesis
3. Cmpare anaerobic and aerobic respiration?

Unit Summary

- A microscope is an instrument that is used to observe objects can not clearly with the naked eye.
- Microscopes are grouped in different type According to the type of radiation they use for observation or image formation. The most common are light and electron microscope.
- The light microscope uses a beam light to form the image of an object, while the electron microscope uses the beam of electron to form the image.
- A simple microscope consists of a single convex lens that is capable of magnifying an object.
- Compound microscope is a microscope that uses too lens systems at the same time.
- Microscope has two major abilities these magnification and resolution
- Magnification is increasing the size of an object to be viewed.
- Resolution is ability of the microscope to show the detailed or the scattered part of an object.
- Mounting is the process preparing a specimen for observation under a microscope.
- A cell is the smallest unit of a living thing. Thus, cells are the basic b uilding blocks of all organisms.
- All cells commonly have cell membrane, cytoplasm, nucleus or DNA and ribosome
- Unicellular organisms are those organisms that are made up of single cell.
- Multicellular organisms are those organisms that are made up of many cells.
- Multicellular organisms have different levels of organization like cell, tissue, organ, organ system, and organisms.
- Mitochondria and chloroplast are double membrane organelles perform cellular respiration and Photosynthesis respectively.
- The formation of ATP (energy) from the breakdown of glucose using oxygen is known as cellular respiration.
- Photosynthesis is food making process in green plants using CO_2, water, chlorophyll pigments and light from the sun in chloroplast.

Review Exercise

Part I, Choose the correct answer from a given alternative options.

1. The _____ is the basic unit of life.
   A. Organism B. cell C. tissue D. organ
2. Which of these structures do all cell share?
   A. nuclear envelope C. organelles
   B. cell walls D. plasma membrane
3. which of the following scientists discover the cell
   A. Robert Hooke B. Anton van Leeuwenhoek
   C. Zaccharias Janssen D. Hans Lipperhey
4. Which of the following feature will help you in distinguishing a plant cell from an animal cell?
   A. cell wall C. mitochondria
   B. Cell membrane D. nucleus
5. Which part of the cell contains organelles?
   A. cytoplasm B. Cytosol C. Cell wall D. nucleus
6. The shape and size of the cell are directly related to:
   A. The size of organism B. their functions
   C. environment D. all
7. Which of the following levels of biological organization shows the correct order from simplest to complex level?
   A. Organism ------ organ system ------ organ ------ tissue ------ cell
   B. Tissue ------ cell ------ organ ------ organ system ------ organism
   C. Cell ------ tissue ------ organ ------ organism ------ organ syste
   D. Cell ------ tissue ------ organ ------ organ system ------ organism
8. Which organelles are responsible for digesting cell waste and foreign bacteria?
   A. Golgi apparatus
   B. Cytoskeleton
   C. Nucleus
   D. Lysosomes

9. Which of these organelles modifies cell products and then packages them for distribution?

A. The nucleus C. The mitochondrion

B. The cell membrane D. The Golgi apparatus

10. The 'powerhouse' of the cell that generates the cell's energy-rich ATP molecules is the:

A. Mitochondrion C. Chloroplast

B. Smooth ER D. Nucleus

Part II: Match items given in column 'B' with items given in column'A'

"A"	**"B"**
1. Cell wall	A. sorting, tagging and distribution of lipids
2. Nucleus	B. provides support, and gives shape to the cell
3. Ribosome	C. it direct and control cell activities
4. Mitochondria	D modify proteins and synthesize lipids
5. Chloroplast	E. the power house of cell
6. Golgi apparatus	F. the site were photosynthesis occur
7. Endoplasmic reticulum	G. Synthesize protein

Part III: Fill in the blank spaces with correct answer.

1. The instrument used to observe cells is _____
2. The ability of the microscope that makes the specimen appear large is known as----------
3. The objective lenses of a compound microscope are------ ,---- ,----- and-----
4.---------- is preparing a specimen for observation under a microscope

Part IV: Give short answers for the following questions.

1. Cells consist of many organelles, yet we do not call any of these organelles as structural and functional unit of living organisms why?

2. Explain how do you calculate the total magnification of your specimen when using a compound light microscope.
3. Reorder the following parts of living things from the largest to the smallest (Cell, organ, organism, tissue, organ, system, organelle).
4. Write the difference between magnification and resolution.
5. The figure 4.10 given below is the structural organization of the animal cell. Indicate the organelles that is indicated by each number.

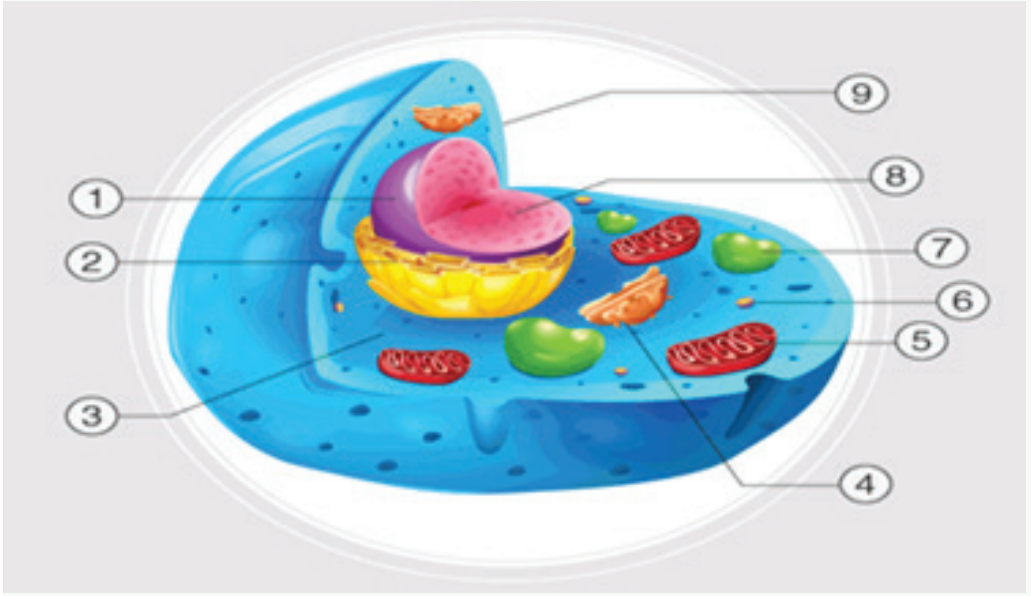


figure 4.10

Key words

Microscope	Resolution	Unicellular	Lens
Magnification	Mounting	Multicellular	Cell
Monocular	Binocular	Cellula	Tissue
Respiration	Photosynthesis	Anaerobic	Aerobic

Unit 5 LIVING THINGS AND THEIR DIVERSITY

Learning Outcomes

At the end of this unit, you will able to:

- distinguish between living and non-living things by describing the features that characterize living organisms;
- discuss if movement i.e. locomotion can characterize all living things or not;
- define classification and its purpose;
- explain the purpose of scientific name;
- list down the hierarchical levels in the classification of organisms;
- describe the distinguishing characteristics of kingdom animalia, Plantae, Protista, Monera and Fungi;
- list common examples of animals, Plantae, Protista, Monera and Fungi;
- describe the body plan of a common animals, Plantae, Protista, Monera and Fungi;
- describe habitats of animals, Plantae, Protista, Monera and Fungi.

Main contents

Introduction

There are at least five million different kinds of living things in the world. These organisms are classified according to their similarities and differences. The need for classifying living things is to identify them and to study their relationship, their origin and development and to understand how life originated.

If you observe your environment you realize that there are millions of different organisms living on the Earth, including animals, plants and microorganism. In this unit you are going to learn about the characteristics of life, naming and how to classify those organisms in to different group.

Activity: 5.1 Discuss in group and present your ideas to classmates.

- What makes living things different from non-living things? Consider as an example butterfly and stone.

5.1. Living Things

At the end of this section you will be able to:

- differentiate between living and non-living things;
- organize and describe characteristics of living things;
- justify why movement or locomotion from one place to another cannot be a defining characteristic of all organisms;
- relate diversity with classification of organisms;
- justify why scientific names of organisms should be used in science than the local names;
- analyze and describe the relationships of the hierarchical levels (Kingdom to Species) in the classification of organisms.

5.1.1. Characteristics of Living Things

Living things belong to any organism or a life form that possesses or shows the characteristics of life or being alive.

However, a living thing possesses certain properties that help define what life is. All groups of living organisms share several key characteristics or functions: movements, respiration, sensitivity, growth, reproduction and excretion. When viewed together, these seven characteristics serve to define life.

- **Movement:** An action by an organism or part of an organism causing a change of position or place. Most single-celled creatures and animals move about as a whole. Fungi and plants may make movements with parts of their bodies.
- **Respiration:** The chemical reactions that break down nutrient molecules in living cells to release energy for metabolism
- **Sensitivity:** The ability to detect or sense stimuli in the internal or external environment and to make appropriate responses. Organisms can respond to diverse stimuli. For example, plants can grow toward a source of light, climb on fences and walls, or respond to touch. Even tiny bacteria can move toward or away from chemicals or light.
- **Growth:** A permanent increase in size and dry mass by an increase in cell number or cell size or both. Non- living organisms grow by addition of new material to the out sides surface, however living organism grow from within using food
- **Reproduction:** All living organisms must have the ability to reproduce. Living things make more organisms like themselves. Whether the organism is a rabbit, or a tree, or a bacterium, life will create more life. Reproduction is the process of making the next generation and may be a sexual or an asexual process.
- **Excretion:** The removal from organisms of toxic materials, the waste products of metabolism (chemical reactions in cells including respiration) and substances in excess of requirements Example the process of respiration produces west product, carbon dioxides, which can be harmful in excess and must be removed

Activity 5.2 Discuss in group and share your ideas to the class

1. Are plants categorized in living things or nonliving things? Justify why?
2. Why movement or locomotion from one place to other cannot be taken as defining characteristics of all living things?

• **Nutrition:** the taking in of materials for energy, growth and development; plants require light, carbon dioxide, water and ions; animals need organic compounds, ions and usually need water

Exercise 5.1

1. List and explain the characteristics of life.
2. Explain why movement or locomotion from one place to another cannot be a defining characteristic of all organisms?

5.1.2. Classification ond Scientific Names of Organisms

Scientists have found and describe approximately1.75million species, moreover new species are being discovered every day. With such diversity of life on the earth how does one go about making sense of it all? One ways to make sense of it is by classification. Classification is an important step in understanding the present diversity and past evolutionary history of life on Earth. It helps make sense of the over whelming diversity of living things.

What is classification?

Classification is sorting or grouping things together on the basis common features /defined characteristic or criteria.

The science of classification is called taxonomy. Classification is usually a hierarchical process. One begins with general and broad differences, and then one systematically introduces more and more detailed and specific criteria.

Why do we classify organisms?

Biologist/ taxonomist classify living things because of:

- To identify those most at risk of extinction and to understand common ancestors.
- It helps scientists to sort organisms in order and to make easy for study.
- It helps them to identify new organisms by finding out which group they fit.

Taxonomy has two branches: the naming of organism or **nomenclature** and Placing of organism in to group, or **systematic** which is done on the basis of their similarities and differences.

There are two ways of classifying organisms. These are artificial and natural.

The artificial classification based on one or a few easily observed characteristics and usually designed for practical purpose with an emphasis on convenience and simplicity.

Example: You could put all the animals that fly in the same group. This group would then include birds, bats and many insects. You could put all animals that live in water and have streamlined, fish-like bodies in the same group. This group would then include fish and whales. They based on arbitrary groupings and have little meaning.

Natural or biological classification system tries to use natural relationships between organisms it consider more evidence than artificial classifications including **internal** as well as **external**\ features. It is a scientific method of classification that groups organisms that share common features.This classification is not random, but rather it describes evolutionary relationships. As a consequence, it is always necessarily hierarchical, where the important features inherited from a common ancestor determine the group in which the organisms are placed. For example, humans and whales both feed their young on milk, which is a characteristic inherited from a common ancestor. This similarity places them under the same class, mammals, even though their habitats are completely different.Each organism is grouped into one of five large groups or

kingdoms, which are subdivided into smaller groups called phyla (singular: phylum) and then smaller and smaller groups with other names.

KEY WORDS

Taxonomy: the science of classifying and naming organisms.

Classification: sorting things based on defined characteristics or criteria

Nomenclature: naming of organism

Genera: a group of closely related species.

Species: is the group of organism that can reproduce to produce fertile offspring

Binomial system is an internationally agreed system in which the scientific name of an organism is made up of two parts showing the genus and the species.

The smallest natural group of organisms is the species. A species can be defined as a group of organisms that can reproduce to produce fertile offspring.

Scientific Naming of Organism

The Binomial System

Organisms were first classified by a Swedish naturalist called Carl Linnaeus (1707 to 1778) in a way that allows the subdivision of living organisms into smaller and more specialized groups. He designed a scientific system of naming organisms called binomial nomenclature. The binomial system of naming species is an internationally agreed system in which the scientific name of an organism is made up of two parts showing the genus and the species. Binomial means 'two names'; the first name gives the genus and the second gives the species. Carl Linnaeus named organisms in Latin using the binomial system.

The scientific naming or binomials naming system should follow the following rules

- Should contain two names (first and the second)
- The first name is the name of the genus name to which group the organisms belongs and it should begin in capital letter
- The second name is the name of species to which the organism belongs.it is written in small letter.

- The scientific name must always be either written underlined or printed in italics.

The scientific name of our human race is Homo sapiens / **Homo sapiens,** similarly the scientific name of some organism listed below in the table Below Table 5.1

No	Common name	Scientific name
1	African elephant	*Loxodonta africana*
2	Lion	*Panthera leo*
3	Teff	*Eragrostis tef*
4	Domestic cat	*Felis catus*

Table 5.1

Activity 5.3

The activity below is not Life Sciences related, but expresses the process of classification. Discuss in group and present the result of your discussion the relationships among the administrative structure: Country, Region, Zone, Wereda and correlates with biological levels of classification

An organism will always have only one scientific name even though they might have more than one common name. Before Linnaeus, the use of common names to refer to organisms caused confusion because there were regional differences in these common names. Because of the scientific names is international agreed and universal it avoids the confusion of local variation in common names. For instance, every biologist will understand that Felis catus means 'house cat' without resorting to the dictionary, no matter what language they speak.

Exercise 5.2

1. Explain the importance of scientific naming organisms.
2. Define what is species?

5.1.3. Hierarchy in the Classification of Organisms (Kingdom to Species)

Linnaeus eventually extends the binomial system to include more groups than just genus and species. He arranged in hierarchy with largest group, the kingdom at the top of hierarchy, the groups he proposed are still used today and, in descending order of size: The sequence of classification is: **Kingdom, Phylum, Class, Order, Family, Genus, and Species**

Figure 5.1: Diagram showing hierarchy of classification

When trying to identify animals, it is this hierarchy or ranking scheme that we follow. We start by identifying the kingdom, to which an organism belongs, then its phylum, class, family, order, and so on.

As you go through the classification hierarchy, you will see that scientists have used broader features to put organisms into kingdoms, which are the largest groups of organisms. When you move down towards the species, which are the smallest groups of organisms, features are becoming specific. In other words, two organisms that belong to the same species share more features than those in the same kingdom but in different species.

Activity: 5.4 Make you own mnemonic to remember the sequence of the classification system This activity allows the learners to be creative. Give the learners the freedom to choose which platform suits them best. The learners have fun and learn the classification system at the same time.

Instructions: Make an easy to remember memory aid to remember the sequence of levels of the classification system.
Materials: pen, paper, imagination!
Procedure: coin mnemonics using first letters of the levels such as-K for Kingdom, -P for Phylum, etc. for frequent memorization easy and of the levels. Example of such mnemonic as: King Philip came over for Good Spaghetti (KPCOFGS) rehearse this mnemonic at the beginning of your class until thisunit is completed.

KEY WORD
Kingdom: are grouped of related phyla or divisions
Phylum or Division: a group of related classes
Class: a group of related orders
Order: a group of related families
Genus: a group of closely related species.
Family: a group of related genera
Species: basic unit of classification or taxonomy

Taxon	**Dog**	**Tiger**	**Maize**
Kingdom	Animalia	Animalia	Plantae
Phylum	Chordata	Chorodata	Magnoliophyta
Class	Mamalia	Mammalia	Liliopsida
Order	Carnivora	Carnivora	Poales
Family	Canidae	Felidea	Poaceae
Genus	*Canis*	*Felis*	*Zea*
Species	*Familiaris*	*Tigris*	*Mays*

Table5.2 The taxonomic breakdown of a few familiar animals and plant

Exercise 5.3

1. Write the classification hierarchy for human being from kingdom to species.

A species can be defined as a group of organisms withsimilar features and these organisms are capable of breeding and produce fertile offspring.You are probably aware of the fact that horses and donkeys belong to the same kingdom, phylum, class, order, family as well as genus but they are from different species. Therefore, if a donkey and the horse happen to breed, they produce an offspring called a mule. The mule is infertile, meaning that it cannot reproduce offspring because it is a product of organisms of different species.Classification hierarchy has many uses. First, it helps scientists to sort organisms in order. Second, it helps them to identify new organisms by finding out which group they fit. Third, it makes easier to study organisms when they are sorted in groups.

Exercises 5.4

I. Choose the best answer from a given alternative options

1.The Swedish scientist who create the binomial naming system was_____

A. Robert Hook B. Carlous Linnaeus

C. Robert H. Whittaker D, Charles Darwin

2.Which of the following in the classification system is the smallest?

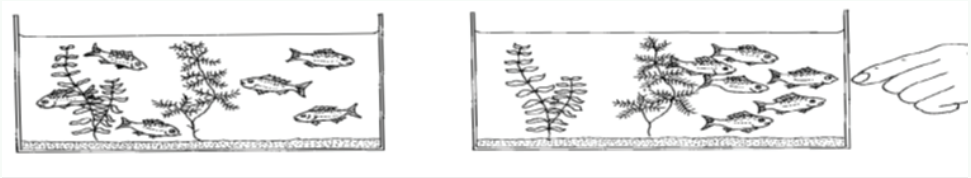
A. Kingdom B. species C. Genus D. class

3. Excretion, irritability and reproduction are characteristics of:

A. all animals and plants B. animals only

C. plants only D. some animals and some plants only

Figure: 5 .2 below shows how fish react when the glass on one side of an aquarium tank is tapped with a finger.



4. What characteristics of living organisms does this demonstrate?

A. Excretion and movement B. excretion and nutrition

C. growth and irritability D. irritability and movement

5. The scientific name of human being is Homo sapiens, the second parts of the scientific name represent-------

A. genera name B. kingdom name

C. species named D. family name

II. Complete the blank space below by choosing the words from list

List: excretion, growth, Sensitivity, movement, nutrition, organisms, reproduction, respiration

A. Living things are often called _____.

B. All living things release energy from their food in a process called _____, which happens inside their cells

C. Some of the energy is used for _____, which usually happens more quickly in animals than in plants.

D. The food from which the energy is released is taken into the body in a process called _____.

E. All living things get bigger as they get older. This process is called _____

F. The production of young is called _____.

G. Waste substances are removed from organisms by the process of _____

H. The seventh characteristic shown by all living organisms is _____, which means that they are sensitive to things around them.

5.2. The Kingdom of Life

At the end of this section you will be able to:

- compare the five Kingdoms of living things by describing their distinguishing characteristics;
- summarize the commonest examples of organisms belonging to each Kingdom;
- describe the body plans of insects such as butterfly, amphibians such as frogs, mosses, liverworts, ferns, conifers such as junipers;
- Junipers, flowering plants;
- relate each Kingdom of organisms to their major habitat types as aquatic, terrestrial or moist.

Activity: 5. 6 Brain storming

i. Have you ever thought of multitudes of life forms that surround us and can you list some of them?

ii. Make groups of five, and then categorize the following organisms given in the following charts into some named group and present the results of your discussion for the class.

Figure: 5.3 Some diversity of life on planet earth

The five kingdom system is the most common way of grouping living things based on simple distinctive characteristics. The five-kingdom system was developed by Robert H. Whittaker in 1969 and was built on the work of previous biologists such as Carolus Linnaeus. Living things can be classified into five major kingdoms:

- Kingdom Animalia
- Kingdom Plantae
- Kingdom Fungi
- Kingdom Protista
- Kingdom Monera (Bacteria)

Activity5. 7

Categorize the organism listed in figure 5.3 above based on the five kingdom of life and explain your reason to classify under any of the kingdom?

5.2.1. Kingdom Animalia

Major characteristics of animals

Members of the animal kingdom are eukaryotic and multicellular but have no cell wall or photosynthetic pigments. They are mostly motile and they are heterotrophic, which means they must feed on other organisms and cannot make their own food. They reproduce sexually or asexually. Animals store carbon as glycogen and fat.

Major groups of animal and their habitats

1. Invertebrates - lack back bones

Insects (Arthropods)

The insects form a very large class of arthropods. Insects live in every possible environment on Earth and are among the most highly adapted of all animal species

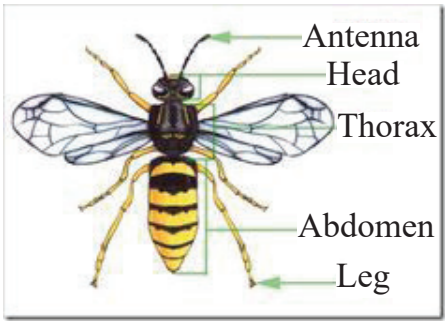


Figure: 5.4 the body structure of insect

Bees, butterflies, mosquitoes, houseflies, earwigs, greenfly and beetles are just a few of the subgroups in this class. Insects have segmented bodies with a firm exoskeleton, three pairs of jointed legs, compound eyes and, typically, two pairs of wings. The body is divided into three parts: head, thorax and abdomen regions. Insects have only one pair of antennae and only three pairs of legs and have no limbs on the abdominal segment. They reproduce sexualy having complet or incomplet metamorphosis.

KEY TERMS

Heterotrophs: heterotrophs are organisms that cannot produce their own food.

hydroskeleton: is flexable skeleton supported by fluid pressure.

Metamorphosis: is a process by which animals undergo extreme, rapid physical changes sometime after birth.

Exoskeleton: is the external skeleton that supports and protects an animal's body.

Worms: Worms are members of several invertebrate phyla, animals that typically have a long cylindrical tube-like, flattened, or leaf like shaped body, no limbs, no eyes and have hydro skeleton. It includes Platyhelminthes (flatworms), Annelida (segmented worms), Nemertea (ribbon worms), nematode (roundworms, pinworms) etc. They vary in size from less than 1 mm (0.04 inch) in certain nematodes to more than 30 m (100) They live in marine, freshwater, and terrestrial habitats. Some types of worms are parasitic, others are free-living.

Figure: 5.5 different worms of phylum Annelida

Activity: 5.8

Collecting and examining flatworms and insects

Materials you require:

A. containers, nets, hand lenses, alcohol solution

Procedure:

1.You may need to use nets to catch some of the organisms. Take care handling any organisms which may sting or bite. Keep different types of specimen you collected (butterfly, grasshopper, spider, Bees, mosquitoes) in appropriate container and examine as the following

2. What features the specimens have in common?

3. Examine their characteristic features, i.e. number of limbs, presence and number of antennae and number of body parts, presence and number of wings.

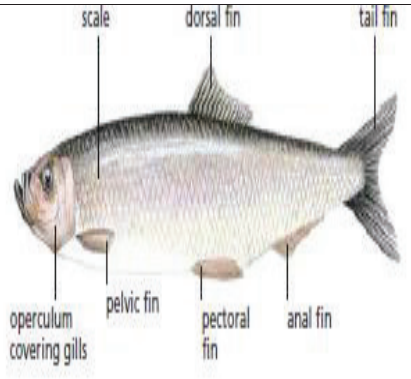
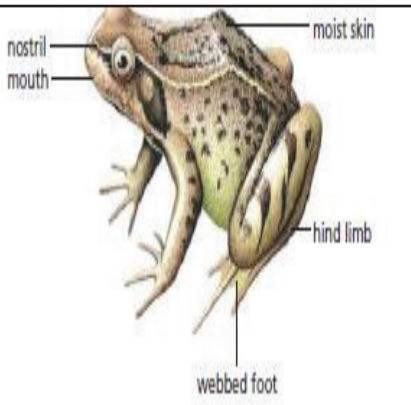
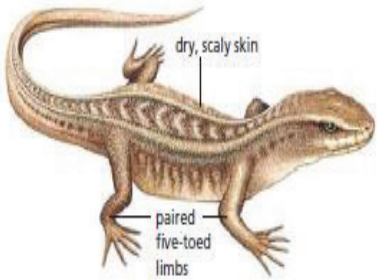
4. You should then make a table of characteristic features like in the following tables

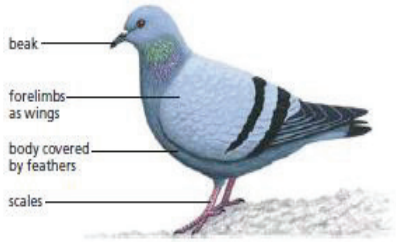
5. Make large well-labeled drawings of each of their specimens

Specimen	Number of body parts	Number of limbs	Antenna	wings
Butterfly				
Grasshopper				

B. Looking at Platyhelminthes (flatworms)

Material used: preserved or fresh specimens of Platyhelminthes (flat worm) hand lenses. Observe, draw and label specimens of these invertebrate phyla

Class	Main features	Examples
FISH	▪ Poikilothermic (cold blooded) vertebrates. ▪ Many of them have a smooth, wet scales on skin, and streamlined shape ▪ Breath by gills, reproduce sexually and fertilization is external ▪ Fins (also used for balance) ▪ Live only in water (aquatic environment) **Example** - star fish - whales - jelly fish	
AMPHIBIAN	▪ Poikilothermic (cold blooded) ▪ Have moist skins with a good supply of Capillaries ▪ Have lungs and skin for breathing ▪ Fertilization external, produce jelly-covered eggs in water ▪ Four limbs, back feet are often webbed to make swimming more efficient ▪ Live both in water and on land. **Example** - froge, toads	
REPTILES	▪ Poikilothermic (cold blooded) ▪ dry skin, with scales ▪ four legs (apart from snakes) ▪ produce eggs with a rubbery, waterproof shell; laid on land ▪ Have lungs for breathing ▪ Most reptiles live in warm habitats **Example** - corocodiles - snakes - lizards -turtles	

Class	Main features	Examples
BIRDS	▪ are homoeothermic (warm blooded ▪ feathers, with scales on legs ▪ two wings and two legs produce eggs with a hard shell, laid on land ▪ lungs for breathing; beak ▪ live in water and on land **Example** - pigeon and duck	
MAMMALS	▪ are homoeothermic (warm blooded) ▪ produce live young ▪ lungs for breathing ▪ females have mammary glands to produce milk to feed young; ▪ four types of teeth ▪ live on land **Example** - cow, human and sheep	

Activity: 5. 9 collecting and examining amphibian (frog)
Material you require: transparent container or cage, pairs of forceps, pairs of gloves.
Procedure: 1.Collect live or freshly killed toads or frogs keep in transparent container or cage.
2. Examine the head and trunk regions of the toad. You should note and identify the following characteristic features: Mouth, Nostrils, Eyes Ear, Trunk, limbs.
3. Does the toad/frog have a tail?
4. Make a large well-labelled drawing of the toad/frog as seen from the slide

Key word

Poikilothermic: an animal whose internal temperature varies considerably.

Homoeothermic: organisms able to maintain a constant internal body temperature.

Exercise 5.5

I. choose the corect answer from the given alternatives

1. Which of the following vertebrates are characterized by four limbs with back feet are often webbed, moist skin and live both in water and on land?

   A. Mammal B. fish C. reptiles D. amphibian

2. Which of the following is not true about insect? They posses

   A. three segmented body B. a pair of antennae

   C. two pairs of legs D. typically two pair's wings

3. Writes the distinguishing characteristic of mammals and birds.

5.2.2. Kingdom Plantae

Major characteristics of plants

Organisms belonging to the plant kingdom are eukaryotic and multicellular organisms. They have a distinct cell wall made of cellulose. Cells are organized into true plant tissues. Plants contain plastids and photosynthetic pigments such as chlorophyll. They are non-motile. Plants make their own food by photosynthesis and are therefore said to be autotrophic. Plants undergo both sexual and asexual reproduction. They store food as starch.

Major groups of plants and their habitats:

Important examples of plants are mosses, ferns, conifers and flowering plants.

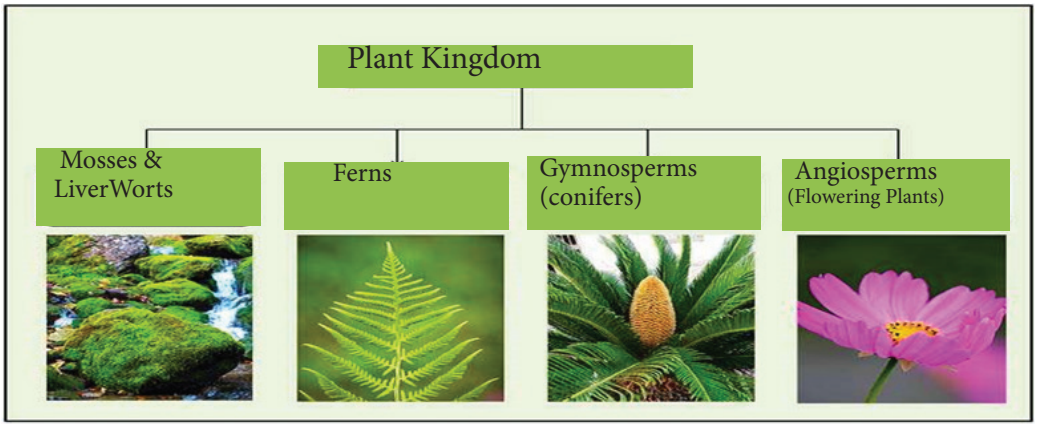


Figure: 5.6 major groups of plants

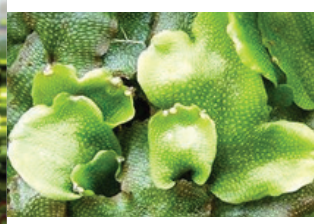
Non vascular plants (lack vascular tissue)

Vascular systems consist of xylem tissue, which transports water and minerals, and phloem tissue, which transports sugars and proteins.

Bryophytes: The Phylum Bryophyta, are the most diverse group with more than 10,000 plant species.

- This phylum includes the mosses, liverworts, and hornworts.
- They lack vascular tissue and wood that can render them structural support.
- They also lack true leaves, stem, and roots that can help them transport water and nutrients.
- Live in moist places and somehow have adapted several methods that can help them thrive in dry periods.
- Reproduce through spores.
- Play an important role in minimizing erosion along bodies of water, carrying out water and nutrient cycling

Mosses



Liver wort

Horn wort

Figure 5: 7 different types' bryophytes

Activity: 5. 10 collecting and examining mosses
Materials you require: microscopes, hand lenses, scalpel blades, forceps, microscope slides and cover slips.
Procedure: 1. in groups search around the school for moss plants around damp walls, rocks, tree barks or damp verandas. Then you should carry collected specimen into the laboratory for detailed study.
2. With the help of a hand lens examine the specimen carefully and identify the parts.
3. You draw and label your specimen.

Figure 5.8 the structure of pteridiophytes (fern)

Key terms:
Vascular Plants: possess vascular tissues (xylem and phloem) that aid them to transport water, minerals and food

Pteridophytes

- have well-developed xylem and phloem
- Pteridophytes are seedless plants but they pass their genetics to offspring through spores that are located on the underside of their leaves known as sporophylls.
- Unlike bryophytes, they are already vascular plants and capable of transporting fluids.
- The stem and leaves have sieve tubes and water conducting cells similar to those in the xylem and phloem of a flowering plant.
- The stem is usually entirely below ground and takes the form of a structure called a rhizome.
- The leaves of ferns vary from one species to another
- Pteridophytes have already adapted to a wide range of habitat: they can be aquatic, terrestrial, and even cold-resistant.

Activity: 5. 11 Collecting and examining a fern

Material you require: hand lenses, scalpels, clean slides, cover slips, and microscopes.

Procedure

1. In groups, you should search for a fern along rivers/stream banks, shady areas beneath trees and along fences.
2. Examine your specimens and identify as many structures as they can.
3. Draw and label their specimen.
4. Observe the lower surface of the leaves (fronds).
5. Draw the lower surface of the specimen showing the arrangement of the spore-forming bodies if there are any there.
6. Using forceps or a needle, remove a capsule if they can see one, mount it on a slide and view under low power. Draw what you see.

Exercise: 5. 6

1. What is the difference between bryophytes and pteridophytes
2. Writes the importance of bryophytes in the environments

Gymnosperms (confers plant)

- The name "gymnosperm" literally means "naked seed", which is exhibited by the members by having cones instead of seeds to reproduce. Their seed are not enclosed in fruit.
- They are widely distributed in the planet but dominate the temperate and arctic regions. The stem and leaves have sieve tubes and water conducting cells similar to those in the xylem and phloem of a flowering plant.
- They are characterized by having wood, and green needle-like or scale-like foliage.
- gymnosperms are good sources of wood and paper

Fig 5.9 structures of gymnosperm

Activity: 5.12

collecting and examining conifers

Material you require: saw, conifer leaves and cones.

Procedure: In groups, you should search and collect a conifer tree in you school compound.

1. Obtain some conifer leaves and cones.
2. Observe them carefully.
3. Make large well-libeled drawings of the leaves of conifer.
4. Examine some conifer cones. Note the seeds attached to the cone. Carefully remove one seed from the cone of conifer and draw it.

B. Collecting and examining angiosperm

Material you require: bean/pea plants with flowers and bean/pea seed, maize plants with flowers and maize grain, hand lenses.

Procedure: Make a collection of flowering plants around your school. Identify them and then classify them according to whether they are monocotyledons or di cotyledons

1. Collect a bean/pea plant and a maize plant.
2. Compare their roots, stems, leaves, flowers and seeds.
3. Make a table of differences between the bean plant and the maize plant.
4. Draw well-labeled diagrams of the bean plant and the maize plant.

Angiosperms (flowering plants)

- They have true root, stem, leaves and flowers as reproductive organ and the seed are enclosed in fruit.
- They reproduce by seeds which are formed in flowers.
- Flowering plants are divided into two subclasses: monocotyledons and cotyledons.
- Their leaves are usually broad and the leaf veins form a branching network

Most angiosperms are good sources of food, medicine, clothing fibers, and wood.

Figure 5:10 flowering plants

Activity 5.13

Develop a table that simplifies and summarizes the kingdoms from mosses to flowering plants as follows. You should copy the example shown here and fill it in.

Division	Characteristics	Examples
Bryophyta		
Pteridophyta		
Gymnosperm		
Angiosperm		

EXERCISE: 5.7

Choose the best answer from the give suggested option

1. True root, stem and leaves are found in_____.
   A. lichens B. algae C. fungi D. ferns
2. A group of plant characterized by the possession of a masses of spore bearing structures under side of their leaves are;
   A. ferns B. lichen C algae D. worts
3. A seed bearing but non flowering plants are
   A. algae B. angiosperm C. gymnosperm D. ferns
4. Which one of the following does not belong to Bryophytes?
   A. Liverwort B. mosses C. hornwort D. ferns
5. Which of the following group is vascular plants?
   A. liverworts B. mosses C. Ferns D. Hornwort

5.2.3 Kingdom Protista

Major characteristics of Protista

Protista are eukaryotic and can be unicellular. The kingdom protista is diverse group.one ways to classify protists according to the way they obtain nutrition. There are animallike, plantlike, funguses like protistis. They reproduce sexually or asexually. Important examples of protists include the organism known as Plasmodium (which causes malaria), Amoeba and Euglena, Trypanosomes.

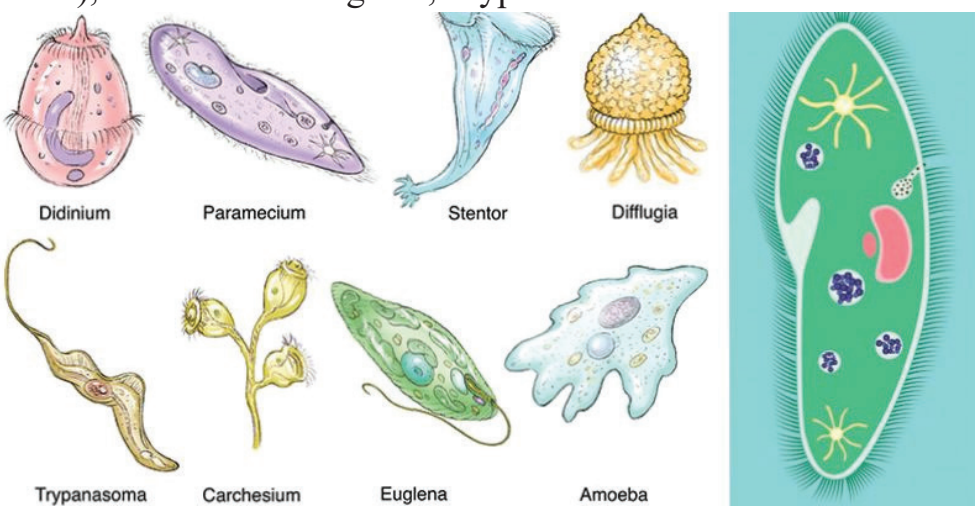


figure 5.11 Protista

There are two major groups of Protista include:

i. Protozoans: are similar to animal cells in that they do not have cell walls Organisms such as Amoeba and Paramecium take in and digest solid food and thus resemble animals in their feeding. They may be called unicellular 'animals'

ii. Protophyta: the plant-like cells which do have cell walls and are similar to algae.Euglena and Chlamydomonas possess chloroplasts and make their food by photosynthesis They often referred to as unicellular 'plants'

Activity: 5. 14 Making hay infusion and observing protozoan (paramecium)

Materials: A hand full of hay, a large beaker, pond water, some milk

Method:

1. Take a hand full of dried grass or hay (free from pesticides or herbicides) and cut the grass into smaller pieces
2. Place the cut grass into the beaker and about 0.5-1 liter of water.
3. Add 1-2 drops of milk. The water will turn slightly turbid. The milk is food for the bacteria and they will start to reproduce. The ciliates feed on the bacteria and will also reproduce.
4. Let the beaker stand open for several days, protected from direct sunlight as this may result in overheating and the heat will reduce the oxygen concentration. Do make sure that the beaker receives sufficient light, though. Photosynthetic algae present in the pond water will produce oxygen.
5. Keep adding 1-2 drops of milk when the turbidity disappears. Bubble some air through the water at regular intervals (using an air-pump from an aquarium) or agitate the water a bit to enrich it with oxygen.
6. Replace the evaporated water.
7. Take some sample from the surface of the water (where there is oxygen) for microscopic investigation. If the water is agitated, then the microorganisms are (of course) not able to collect beneath the water surface.
8. Observe paramecium using microscope and draw the

5.2.4. Kingdom Monera

Major characteristics of Monera

The Kingdom Monera consists of prokaryotic unicellular organisms. No nuclear membrane or membrane-bound organelles such as chloroplasts, Golgi complex, mitochondria or endoplasmic reticulum are present. Monera have a cell wall of protein plus polysaccharide compound, but not cellulose. They reproduce asexually by binary fission. Important examples of Monera include blue green algae and Bacteria.

Major groups of monera and their habitats (Blue Green algae & Bacteria)

Blue-green algae, is prokaryotic singled celled photosynthetic organisms containing a blue pigment in addition to chlorophyll they also called **cyanobacteria.**

Predominantly occur singly or in colonies in diverse habitats in freshwater or a terrestrial environment

They are microscopic but can be seen when they are in a colony, or bloom. Cyanobacteria contain only one form of chlorophyll (a green pigment.) In addition, they contain various yellowish carotenoids, the blue pigment phycobilin and, in some species, the red pigment phycoerythrin. The combination of phycobilin and chlorophyll produces the characteristic blue-green color from which these organisms derive their popular name. Cyanobacteria are the first organisms known to have produced oxygen as a byproduct of photosynthesis

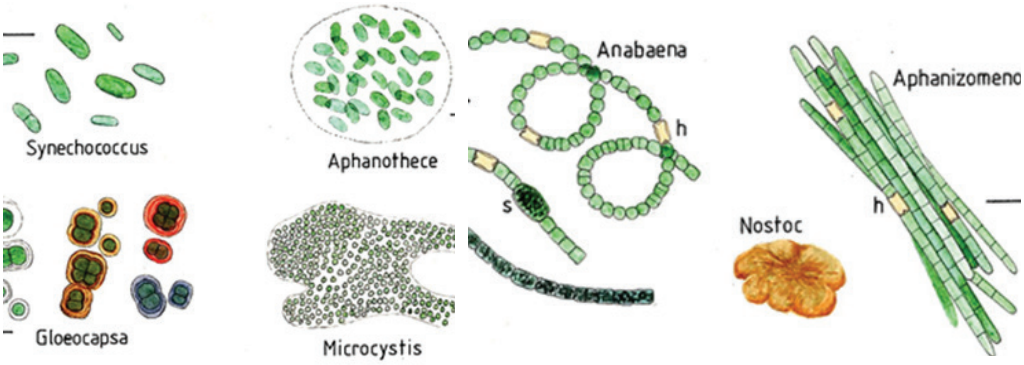


Figure.5:12 Colony and filamentous form of blue green algae

Bacteria: are very small organisms consisting of single cells they lack organized nucleus and chlorophyll pigments

Their cell walls are made, not of cellulose, but of a complex mixture of proteins, sugars and lipids (peptidoglycan).They can be found in various shapes and sizes, may be spherical, rod-shaped or spiral and some have filaments, called **flagella,** projecting from them and serve for movement. The genetic material DNA is contained in the cytoplasm called nucleoid. Bacteria are found everywhere and are the most numerous organisms on Earth.

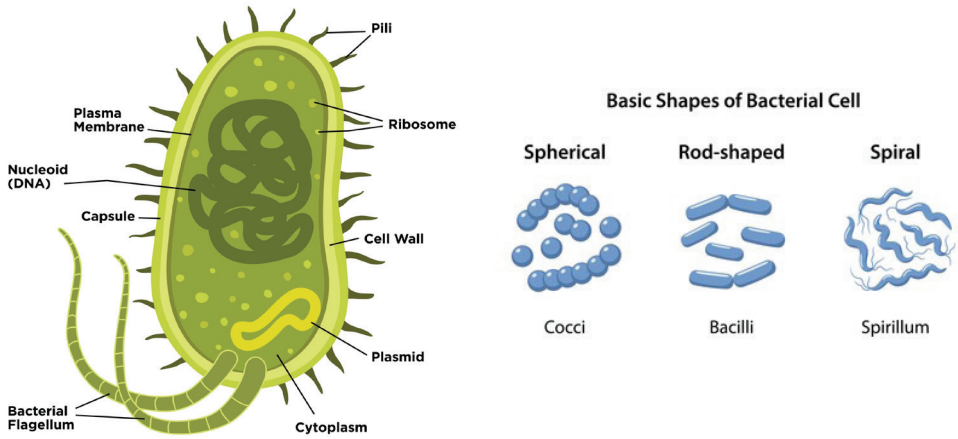


Figure: 5.13 structures and shapes of bacteria

Exercise 5.8

1. Define prokaryotes and eukaryotes cell.
2. Explain the difference between prokaryote and eukaryotes.

5.2.5 Kingdom fungi

Major characteristics of fungi

Fungi are eukaryotic (have membrane bounded nucleus) organisms that can be multicellular or unicellular. Mushrooms and molds are examples of multicellular fungi and yeast is an example of a unicellular fungi. All fungi have a cell wall made of chitin.

They are non-motile (not capable of movement) and consist of threads called hyphae. Fungi are heterotrophic organisms which mean they require organic compounds of carbon and nitrogen for nourishment. They are important as decomposer (saprophytes) and can be parasitic. They store carbon as glycogen, not in the form of starch. Fungi reproduce sexually and asexually by spore formation.

Mushroom **Yeast** **Mold**

Figure 5:14 different types of fungi

Yeast: single-celled fungi. It converts complex carbohydrates into alcohol and carbon-dioxide. Used for a variety of commercial purposes like baking ('injera' or bread) and in the production of alcohol. **Mould:** fungi which grow on decayed bread, cheese, fruit or other food. Many of the mould fungi live in the soil or in dead wood.Another important example of a useful fungus is *Penicillium* (a fungus which was used to make penicillin, one of the most powerful antibiotics ever created). mushroom is the flush, spore bearing fruiting body of fungus typically produced above ground, on soil, or on its food sources. Edible mushrooms are the flush and edible fruiting bodies .includes many fungal species that are harvested, wild or cultivated.

Activity: 5.15 collecting and looking at fungi
Materials you require: microscopes, hand lenses, scalpel blades, forceps, microscope slides and cover slips.
Procedure: 1. In groups search around the school for mushroom in the school compound or if necessary you can grow your own fungi (moulds) on a little moist injera or by letting a piece of fruit go rotten. Some of them may be quite big like (mushroom) but they may want to use a microscope to look at some of them. Look at different structures of fungi (fruiting body or mycelium.Draw several different types of fungus.

Unity summary

- The seven characteristics of living things are movement, respiration, sensitivity, growth, reproduction, excretion and nutrition
- A species is a group of organisms that can reproduce to produce fertile offspring.
- The binomial system is an internationally agreed system in which the scientific name of an organism is made up of two parts showing the genus and the species.
- Classification is a way of sorting organisms into a meaningful order
- The artificial classification based on one or a few easily observed characteristics
- Natural or biological classification system tries to use natural relationships between organisms
- The smallest natural group of organisms is the species
- The five kingdom system is the most common way of grouping living things based on simple distinctive characteristics.
- The sequences of classification are Kingdom, Phylum, Class, Order, Family, Genus, and Species
- Living things can be classified into five major kingdoms: Kingdom Animalia, Plantae, Fungi, Protista, and Monera (Bacteria
- Animal kingdom is eukaryotic and multicellular but has no cell wall or photosynthetic pigments. They are mostly motile and they are heterotrophic,
- Insect has three parts: head, thorax and abdomen regions with three pairs of jointed legs, compound eyes and, typically, two pairs of wings
- Plant kingdom is eukaryotic and multicellular organisms. They have a distinct cell wall made of cellulose.
- Bryophytes are non-vascular plants that lack true leaves, stem, and roots that can help them transport water and nutrients.
- Pteridophytes are vascular plants that have well-developed xylem and phloem that reproduce by spores.
- Gymnosperms are vascular plants that bearing cone.
- Angiosperms are flowering plants that reproduce by seeds which are formed in flowers.

- Protista are eukaryotic and can be unicellular or simple multicellular. They reproduce sexually or asexually.
- Monera are of prokaryotic unicellular organisms that lack the membrane bounded organelle and reported by binary fusion
- Bacteria are very small organisms consisting of single cells they lack organized nucleus and chlorophyll pigments
- Fungi are eukaryotic (have membrane bounded nucleus) organisms that can be multicellular or unicellular and have chitin cell wall.

Review Exercise

I. Write "true" if the statements are correct and "false" if the statements are incorrect

1. Blue green algae belongs to the kingdom-Protista
2. Bacteria are found everywhere and are the most numerous organisms on Earth.
3. The Kingdom Monera consists of prokaryotic multicellular organisms.
4. Protoctists are single-celled organisms containing a nucleus.
5. Amphibians can breathe in air or in water.
6. Mammals have fur, they suckle their young and the young develop inside the mother.
7. Mosses have well-developed stems, leaves, roots and reproduce by spores.
8. Fungi are made up of thread-like hyphae and an autotroph organism.
9. Insects mostly live on land and have wings and two pairs of legs.

II. Match correct terms/meaning given in column 'B' with their correct levels given in column 'A'

'A'	'B'
1. Plants	A. made up of thread-like hyphae & reproduce by spores.
2. Fungi	B. makes their food by photosynthesis.
3. Protists	C. gets their food by eating plants or other animals.
4. Animal	D. single-celled organisms containing a nucleus.
5. Monera	E. single celled organism lacking nucleus

III. Fill the following question by the appropriates terms

1. The type naming organism using the Latin name indicating genus and species is called_____
2. _____group of organisms which are able to interbreed and produce fertile offspring.
3. _____is the scientific name of our human race.
4. _____the type of asexual reproduction in the Kingdom Monera.
5. Highest grouping in a classification system_____

IV. Choose the best answer from the following suggested option

1. Which one of the following is not a defining characteristic of all organisms?
   A. Growth B. respiration C. locomotion D, Reproduction
2. The smallest natural group of organisms is the_____.
   A. Kingdom B. Species C. phylum D. class
3. If two organisms belong to the same order, then they must also belong to the same
   A. Genes B. class C. family D. species
4. An organism's scientific name is based on how it is classified. Which of the following levels of classification determine the name?
   A. Genus and species B. Phylum and class
   C. Order and family D. kingdom and spies

5. The science of identifying, classifying, and naming living things is called?
   A. System B. nomenclature C. Taxonomy D. hierarchy
6. The level below the kingdom is_____.
   A. Genus B. order C. phylum D. class
7. The sequence of scientific category from general to specific:
   A. Species---- genera---family---order --- class---phylum kingdom
   B. Genera--- family---order ---class---phylum kingdom---genus
   C. order---class---phylum---kingdom---family---species---genus
   D. kingdom----phylum---class---order---family---genus---species
8. Protists are;
   A. All autotrophs cannot make their own food
   B. All autotrophs can make their own food
   C. They can be either autotrophs and heterotrophs
   D. They are not eukaryotic (do not have nucleus
9. Which kingdom best fit a mushroom?
   A. Animal B. Fungi C. plant D. protist
10. The main function of the vascular bundles found in many plants is to ...
   A. taps energy or food making
   B. carry out photosynthesis
   C. assist plant in reproduction
   D. transport substances around the plant
11. The scientific (and common) names of three animals are: Canis familiaris (dog), Canis lupis (wolf) These three animals all belong to the same ...
   A. Class but different genera
   B. Genus but different species
   C. Species different genera
   D. Class different species
12. In a five-kingdom system of classification, bacteria are members of
   A. kingdom B.fungi C. plants D.monera E.protests

V. Give short answer for the following question

1. Explain the importance of classification hierarchy.
2. Describe the relationship of hierarchical levels in classification.
3. Explain the major characteristics kingdom animalia and fungi.
4. Writes the unique or distinguishing characteristics of kingdom plantia, Protista and monera.
5. Sort the following list of organisms with respet to the kingdom they belong

Lizard	amoeba	paramecium	blue green	algae
Papaya	spider	termites	mushroom	yeast
Sunflower	pigeon	rat	maize	salmonella
Trypanosome	pea	euglena	mold	Sorghum
bacteria				

Key words

Living things	Genus	Gymnosperms
Classification	Pteridophytes	Protista
Specie	Class	Monera
Binomial system	Family	Vascular
Kingdom	Eukaryotic	Angiosperms
Sorting	Multicellular	Spore
Phylum	Unicellular	Prokaryotic
Order	Bryophytes	Poikilothermic

Unit 6 EARTH IN SPACE

Learning Outcome

At the end of this unit, you will be able to:

- describe the shape of the Earth;
- identify evidences supporting the shape of the Earth;
- list local and global ideas about the shape of the Earth;
- name dimensions (circumferences, diameters, and angular distances) of the Earth;
- recognize all parts of the Earth;
- describe the organization and contents of the different parts of the Earth;
- explain different observations about the Earth in terms of the nature and behaviors of the different parts of the Earth;
- demonstrate movements of the Earth (revolution and rotation);
- explain the effects of motions of the Earth;
- construct the model of Earth and use it to explain phenomena related to its motion;
- identify atmospheric and lithospheric systems;
- explain their cycle effects of the Earth;
- describe the measuring techniques for too big (Earth) and to small (continental drift) quantities measurement and estimation.

Contents	
Section	**Learning competencies**
6.1. Shape & dimensions	• Describe the shape of the Earth • Identify evidences supporting the shape of the Earth • List local and global ideas about the shape of the Earth • Name dimensions (circumferences, diameters, and angular distances) of the Earth
6.2. Parts of the Earth (Body & Atmosphere)	• Recognize all parts of the Earth • Describe the organization and contents of the different parts of the Earth • Explain different observations about the Earth in terms of the nature and behaviors of the different parts of the Earth.
6.3. Movements of the Earth	• Demonstrate movements of the Earth (revolution and rotation) • Explain the effects of motions of the Earth. • Construct the model of Earth and use it to explain phenomena related to its motion
6.4. Atmospheric and lithospheric Systems & Cycles, (effects, measurement ideas/ estimation)	• Identify atmospheric and lithospheric systems. • Explain their cycle effects of the Earth • Describe the measuring techniques for too big (Earth) and to small (continental drift) quantities measurement and estimation

6.1 Shape and Dimensions of The Earth

At the end of this section you will be able to:

- describe the shape of the Earth;
- identify evidences supporting the shape of the Earth;
- list local and global ideas about the shape of the Earth.

Introduction

This unit introduces the concept of Earth in Space, parts of the Earth and shape and dimensions of the Earth. The unit also presents atmospheric and Lithospheric Systems & Cycles, (effects, measurement ideas/estimation).

Activity 6.1

Form a group and perform the following activities.

I. By referring internet explorer or other reference materials discuss about the shape of the earth

II. Identify the local and global assumptions and evidences about its shape.

What Is Earth?

Our Planet Earth is one of the eight planets in our Solar System and the only known planet to support life. It is our home planet that everything is just right for life to exist. Geologists study the materials, processes, products, physical nature, and history of the Earth.

Shape of the Earth

What are the local ideas about the shape of the Earth?

In ancient times, mankind has different questions and assumptions about the geometrical shape of our planet earth. Locally most cultures describe the Earth as flat such as:

- The early ancient Greeks, Sumerians, Babylonians, Egyptians and Vikings all believed that the Earth was a flat disc or plane surrounded by water. This was based on the evidence of what they saw around them.
- The ancient Chinese believed that the Earth was a flat square shape surrounded by heavens that were a round ***egg shape.*** This was based on their belief in a heaven that was above the Earth.
- Members of the Flat Earth Society claim to believe the Earth was flat. A Flat Earth model depicting Antarctica as an ice wall surrounding a disc-shaped Earth. Walking around on the planet's surface, it *looks* and feels *flat.*

Project work 6.1: Model of the earth
Form a group and build model of the earth including its body parts and make discussion about shape with your class mates then present your work to the rest of the class.

Figure 6.1. Flat shape assumption of the earth

The primary reason why ancient people believed the Earth was flat because it look flat from our vantage point on the ground. The misconception arises simply because the Earth is big.

Another assumption about the shape of the were Earth was round. Most ancient Greeks believed the Earth was round not flat. The two main reasons for their assumption were:

- **Lunar eclipses:** during a lunar eclipse the shadow of the Earth always had a round profile.
- **Star patterns:** The second observation was the change in the pattern of stars as you move from north to south.

Around 500 B.C. Pythagoras first proposed a spherical Earth shape. His idea lies, mainly on aesthetic grounds rather than on any physical evidence. Based on actual physical evidence Aristotle (384-322 B.C.) listed the following arguments for a spherical Earth.

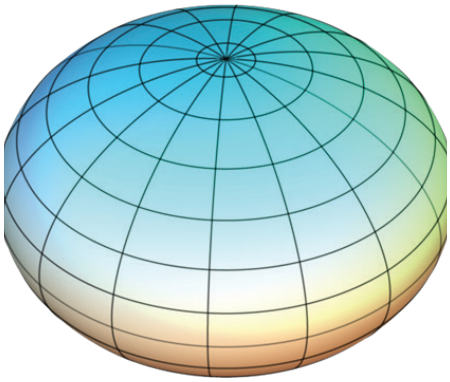
- Ships disappear hull first when they sail over the horizon,
- Different constellations are visible at different latitudes.

Figure 6.2. if earth were flat , its shadow during an eclipse would be straight on the moon, not covered.

Today scientists argue that the shape of Earth is "Oblate spheroid" where 'Oblate' refers to a slight elliptical appearance and 'spheroid' means almost a sphere but not actually a sphere. This describes the true shape of the Earth, which means flatten at the poles and bulges in the middle.

Evidences about the true Shape of the Earth

There are many ways to prove that the earth is *spherical.* The following are some of them:



1. Circumnavigation of the earth: If you travel across the world along a straight path in a fast flying plane without stopping anywhere, you would come back to the same place from where you started. This is called circumnavigation.

Figure: 6.3. Oblate spheroid
Shape of the Earth

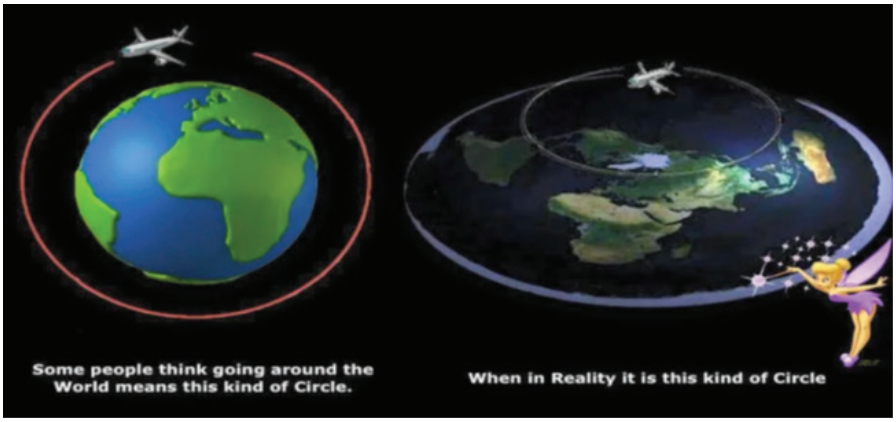


Figure 6.4 . different assumptions about the earth when flaying plane

2. Earth's curved horizon – The earth's horizon when seen from a ship, a plane, or a high cliff appears curved. The curved horizon widens as the observers altitude increases until it becomes circular. If the earth were not spherical, there would be no circular horizon. The curvature of the horizon is influenced by the curvature of the earth's surface.

3. Ship's visibility:

When two ships on the same line of observation are coming towards the observer while maintaining a considerable distance, the front Ship will be seen before the ship at the back. If the earth's surface were flat, both ships could be seen at the same time. This Provides the earth is spherical in shape.

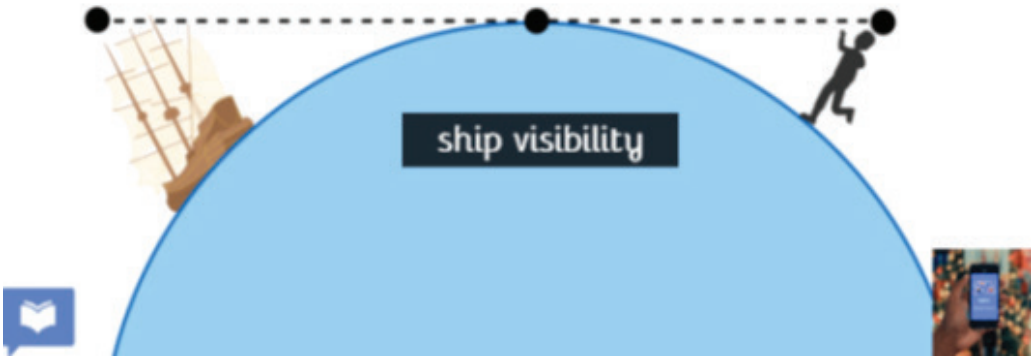


Figure 6.5. Ship's visibility

4. Sun rise and sun set: The sun rises and sets at different times in different places. As the earth rotates from west to east, places in the east see the sun earlier than those in the west. If the earth were flat, the whole world would have sunrise and sunset at the same time. But we know this is not happen because of spherical the shape of the earth.

5. The lunar eclipse: The shadow cast by the earth on the moon during a lunar eclipse is always circular. It takes the outline of an arc of a circle. Only a sphere can cast such a circular shadow as shown in figure 6.6

Figure 6.6. The position of the Sun, moon and earth, in an eclipse

6. Space photographs: Pictures taken from high altitudes by rockets and satellites show clearly the curved edge of the earth. This is perhaps the most convincing and the most up-to-date proof of the earth's sphericity.

7. The changing altitude of the sun - Even though the sun is at constant position in the sky. In the morning and evening the sun observed to be at low level while at noon the sun observed to be at a high level.

Exercise 6.1

I. Choose the best answer from the given alternatives.

1. The Earth's actual shape is most correctly described as
   A. a perfect sphere C. an oblate sphere
   B. a circle D. an eccentric ellipse
2. Which object best represents a true scale model of the shape of the Earth?
   A. a Ping-Pong ball C. a football
   B. an egg D. globe
3. The shape of the earth is_____
   A. egg shaped
   B. spherical
   C. Angular
   D. Flat

II. Short answer questions.

1. Write evidences that support the true shape of the earth.
2. List local and global ideas about the shape of the earth

6.1.2 Dimensions of the Earth

At the end of this section you will be able to:

- name dimensions (circumferences, diameters, and angular distances) of the Earth.

Activity 6.2

Form a group and discuss about circumference, diameters and angular distances of the earth. Then present your discussion to the whole class.

How big is Earth?

The dimension of the earth can be expressed in terms of radius, diameter, circumference, mass, and time.

Earth, is one of the eight planets and its average distance to the sun is **1 AU (1.496 × 10^8 km).** Its diameter (the distance from one side to the other through Earth's center) is about 12,756 kilometers), a polar diameter of 12,714 km, and its mass is **5.974 × 10^{24} kg.**

Earth's circumference (around the equator) is 40,075 kilometers; however, from pole to pole the meridional circumference of Earth is only 40,008 km. The Earth's diameter is also wider at the Equator, creating a phenomenon called an **equatorial bulge.**

The orbital and rotational period of planet Earth is 365.256 days and 23.9345 hours, respectively.

Angular distance is (also known as angular separation) is the angle between the two sightlines, or between two point objects as viewed from an observer.

Note

- ***Equatorial diameter* -** is longer than polar diameter (12,756 km).
- **Polar diameter -** is shorter than equatorial diameter (12,714 km).
- **Equatorial circumference -** is longer than polar circumference (40,075 km).
- **Polar circumference -** is shorter than equatorial circumference (40,008 km).
- **AU- astronomical unit**

Exercise 6.2

i. Fill the blank space

1. The Earth's diameter is wider at the Equator, creating a phenomenon called an _____.
2. _____ is the angle between the two sight lines.

ii. Short answer questions

1. What is the Precise measurements of the Earth
   a. polar diameter?
   b. Polar circumference?
   c. Equatorial diameter?
   d. Equatorial circumference?

6.2 Parts of the Earth (Body and Atmosphere)

At the end of this section you will be able to

- recognize all parts of the Earth;
- describe the organization and contents of the different parts of the Earth;
- explain different observations about the Earth in terms of the nature and behaviors of the different parts of the Earth.

Structure of the Earth

Activity 6.3. Group discussion

Form a group and discus the following activities and present your discussion to the class.

1. What are the layers of the atmosphere?
2. 2. In which layer a) we live on? b) and air plane flay?

Structure of the Earth is the composition of the atmospheric layer and solid or mineral part of the Earth. The structure of the earth consists of

1. External structure (Outer zone) and
2. Internal structure (Inner zone)

6.2.1 External Structure of the Earth (Outer zone)

External structure of the earth consists of different layers of the Atmosphere.

Atmosphere

Atmosphere is the thin layer of gases held on the earth by gravitation attraction. Earth's atmosphere is so much more than the air we breathe. It is composed by abiotic (non-living matter) and biotic matter (living organism). Non-living matter found in the atmosphere includes mixture of gases, water vapor and dust particles. Atmosphere consists of different gases such as carbon dioxide(0.03%), oxygen(21%), nitrogen(78%) and other gases(0.97%). The living organism includes the smallest or microscopic organisms like bacteria.

Vertical stracture of an atmosphere is the run of pressure, temperature, density, chemical composition with distance from the center of the Earth.

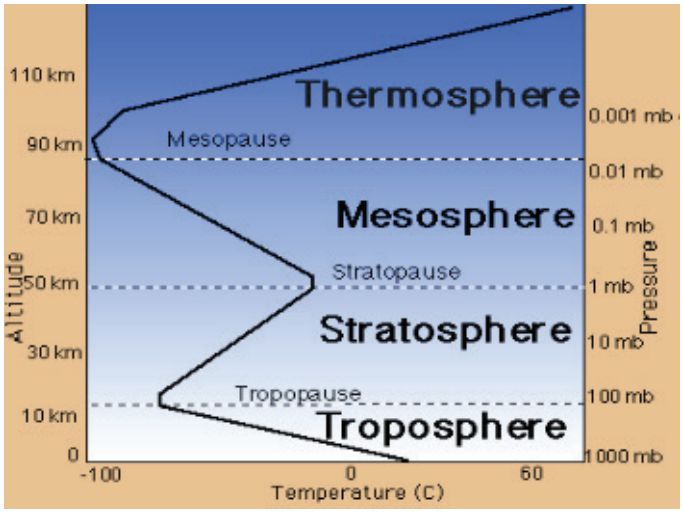


Figure 6.7 vertical structure of the atmosphere

I. Troposphere

The troposphere is the lowest layer of Earth's atmosphere - the part we live in. It contains most of our weather such as clouds, rain, and snow. It contains about 75% of all of the air in the atmosphere, and almost all of the water vapor (which forms clouds and rain). The decrease in temperature with height is a result of the decreasing pressure. Thus, air higher up is cooler than air lower down.

The lowest part of the troposphere is called the boundary layer. The top of the troposphere is called the **tropopause.** This is lowest at the poles, where it is about 7 - 10 km above the Earth's surface. It is highest (about 17 - 18 km) near the equator. Air is warmest at the bottom of the troposphere near ground level. Air gets colder as one rises through the troposphere. That is why the peaks of tall mountains can be snow-covered even in the summer time.

II. Stratosphere

Stratosphere is the second layer of the atmosphere as you go upward. This extends upwards about 50 km above the surface of the Earth. It contains much of the ozone in the atmosphere. The increase in temperature with height occurs because of absorption of ultraviolet (UV) radiation from the sun by this ozone. Temperatures in the stratosphere are highest over the summer pole, and lowest over the winter pole. By absorbing dangerous UV radiation, the ozone in the stratosphere protects us from skin cancer and other health damage. The Lower boundary of the stratosphere is called the tropopause; the upper boundary is called the ***stratopause.***

III. Mesosphere

The mesosphere is the third layer of Earth's atmosphere. The mesosphere is directly above the Stratosphere and below the thermosphere. It extends from about 50 to 85 km above our planet. Temperature decreases with height throughout the mesosphere. The coldest temperatures in Earth's atmosphere, about -90° C at the "mesopause."

The boundary between the mesosphere and the thermosphere above it is called the **Mesopause.**

IV. Thermosphere

The thermosphere lies above the mesopause, it extends from about 90 km to between 500 and 1,000 km above our planet. It is a region in which temperatures again increase with height. This temperature increase is caused by the absorption of energetic ultraviolet and X-Ray radiation from the sun. The temperature of the thermosphere varies between night and day and between the seasons.

The boundary between the thermosphere and the exosphere above it is called the **Thermopause.**

V. Exosphere

Located between about 700 and 10,000 kilometers above Earth's surface, the exosphere is the highest layer of Earth's atmosphere.

Exercise 6.3

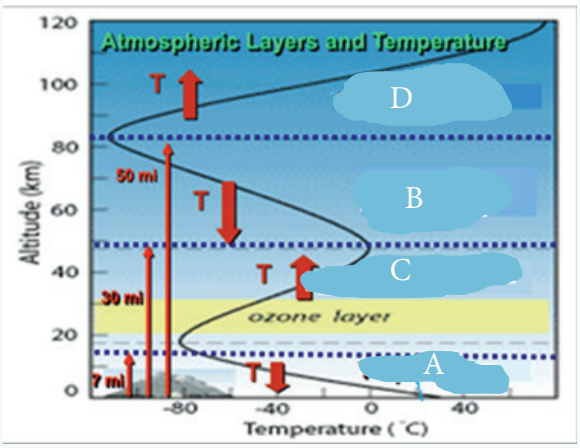
I. Choose the best answer from the given alternatives.

1. The _____ is the outer most layer of Earth's atmosphere.
   A. troposphere C. stratosphere
   B. exosphere D. thermosphere
2. What makes up nearly 78 percent of the Earth's atmosphere?
   A. Oxygen C. Nitrogen
   B. Carbon dioxide D. Hydrogen
3. What percentage of the earth's atmosphere does oxygen comprise?
   A. 75% B. 50% C. 21% D. 32%
4. In which layer of our atmosphere weather occurs?
   A. Troposphere C. Stratosphere
   B. Exosphere D. Thermosphere

II. Short answer questions

1. Referring to diagram write the name of each layer in the Earth's atmosphere

   A. _____
   B. _____
   C. _____
   D. _____



6.2.2 Internal Structure of the Earth (Inner zone)

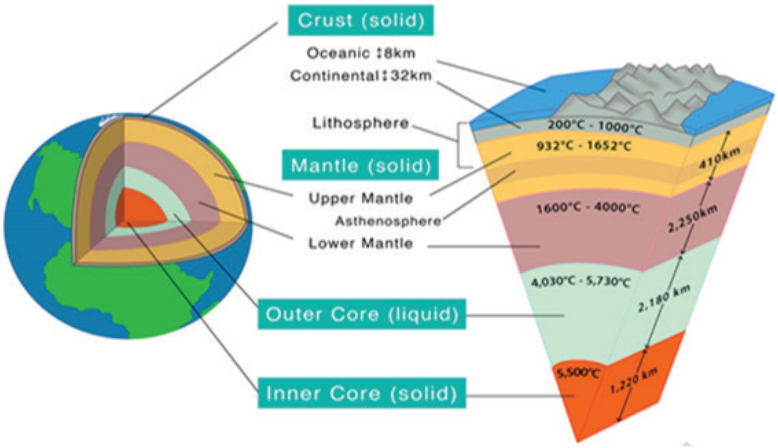
Project work 6.2: Making a model of solid Earth layers

Using 5 colors of modeling clay and waxed dental floss make in group a model of solid Earth layers and show the model to your teacher.

Directions

- Waxed dental floss
- Form a ball to represent the inner core. We choose red to represent the intense heat of the inner core.
- Next, roll out a circle of another color and wrap around the ball and roll gently. This next layer represents the outer core.
- Each subsequent color will need more modeling clay than the last. You will need a color to represent the lower mantle, another for the upper mantle, and the outer layer for the crust.
- Once your ball of 5 layers of modeling clay is complete, use a piece of waxed dental floss to cut the ball down the middle, revealing all the layers underneath.

*Be sure not to press the layers too firmly together so that the colors don't mix.



i. Crust
ii. Mantle
iii. Core

Figure : 6.8. Internal structure of the earth

I. The Crust

The Earth crust is the outside and coldest layer of the earth and is made of solid rock, mostly basalt and granite. It is the thinnest layer and forms the outer shell on which life exists. There are two types of crust; oceanic and continental. Oceanic crust is denser, 5 to 10 kilometers thick and mainly composed of basalt. However, Continental crust is less dense, thicker, and mainly composed of granite, and the upper layer of the earth crust.

II. The mantle

The mantle is the thickest layer, lies between Earth's dense, super-heated core and its thin outer layer, the crust. It is composed of atomic elements include oxygen, silicon and magnesium. It is a solid layer but acts like a viscous liquid due to temperatures being close to the melting point of key minerals in this layer. Below the crust is the upper mantle, with the upper-most portion referred to as the **asthenosphere.** The upper mantle is liquid rock, and very hot. The lower mantle is the lower liquid portion of the mantle.

Note: The lithosphere is the solid, outer part of the Earth, including the brittle upper portion of the mantle and the crust.

III. The Core

Earth's core is the very hot, very dense center of our planet. It is composed mainly of an iron and nickel alloy. The core is divided into inner and outer core.

The outer core is a liquid because the temperatures there are adequate to melt the iron-nickel alloy. However, the inner core is a solid even though its temperature is higher than the outer core.

Exercise 6.4

I. Fill the blank space

1. The boundary between the mesosphere and the thermosphere is called _____.
2. _____is located at the bottom of the mesosphere and the boundary between the mesosphere and the stratosphere below.
3. Lower boundary of the stratosphere is called the_____.
4. Upper boundary of the stratosphere is called the _____.

II.Choose the best answer from the given alternatives.

1. Which layer of the earth is solid ?
   A. Inner mantle and inner core
   B. Crust and outer core.
   C. Crust and the inner core
2. Which layer of the earth is liquid ?
   A. The outer core
   B. Mantle crust and the inner core
   C. Crust and outer core.
3. Which layer of the earth is both solid and liquid?
   A. Mantle B. Crust C. Core
4. What do we call the center of the Earth?
   A. Mantle B. Crust C. Core

5. What is the outermost layer of the Earth that consists of the continents and the oceans?
   A. Mantle B. Crust C. Core
6. What is the layer of the Earth between the outer core and crust?
   A. Core B. Mantle C. Crust
7. What do we call the crust that makes up the continents (land)?
   A. Tectonic plates C. Continental crust
   B. Oceanic crust
8. The hottest layer of the Earth is:
   A. Mantle C. Outer core
   B. Inner core

6.3. Earth's Movements

At the end of this section you will be able to:

- demonstrate movements of the Earth (revolution and rotation);
- explain the effects of motions of the Earth;
- construct the model of Earth and use it to explain phenomena related to its motion.

The Earth is in motion all the time. People cannot feel this motion because they move with it.

There are two types of movements of the earth, namely:

1. The rotation of the Earth on its own axis
2. The revolution of the Earth around the Sun

Activity 6.4

Form a group and discuss the following phenomena.

i. Why do not we feel when the Earth move?

ii. What would happen if the Earth didn't rotate?

Share your views with the rest of the class

Nicolaus Copernicus was a Polish astronomer known as the father of modern astronomy. He was the first modern European scientist to propose that Earth and other planets revolve around the sun, or the Heliocentric Theory of the universe.

6.3.1 Rotation of Earth

Earth Rotation is the motion of the earth around its axis of rotation. An Imaginary a line passing through the center of Earth that goes through both the North Pole and the South Pole is called an axis. The movement of Earth around its axis, from west to east is rotation..

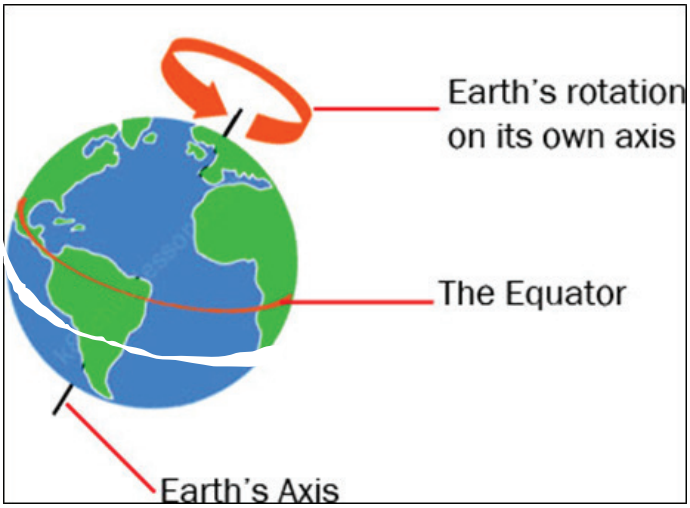


Figure 6.9 rotation of the earth

Earth requires 23 hours, 56 minutes, and 4 seconds to make one complete rotation on its axis. Hence the length of a day on Earth is actually 24 hours.

Effects of motions of the Earth

The rotation of the Earth has the following effects:

- The spinning of the earth on its axis causes days turn into nights and vice versa.
- A difference of one hour is created between two meridians which are 15 degrees apart.
- A change in the direction of wind and ocean currents. Winds and ocean currents deflect to the right in the Northern Hemisphere and to the left in the Southern Hemisphere as a result of rotation.
- The rise and fall of tide every day.

6.3.2 Revolution of the Earth

Earth Revolution is the motion of the earth around the sun. For Earth to make one complete revolution around the Sun takes 365.24 days. This amount of time is the definition of one year.

The closest Earth gets to the Sun each year is at perihelion (147 million km) on about January 3rd and the furthest is at aphelion (152 million km) on July 4th. During one revolution around the Sun, Earth travels at an average distance of about 150 million km. Earth revolves around the Sun at an average speed of about 27 km per second, but the speed is not constant. The planet moves slower when it is at aphelion and faster when it is at perihelion. The reason the Earth has seasons is that Earth is tilted 23 ½ degree on its axis. During the Northern Hemisphere summer the North Pole points toward the Sun and in the Northern Hemisphere winter the North Pole is tilted away from the Sun.

Note:

1. The farthest (maximum distance) position from the sun in orbit of the earth is called aphelion while the nearest position of the earth to the Sun is known as perihelion

2. Meridian is acircle on the surface of the earth passing through the poles.

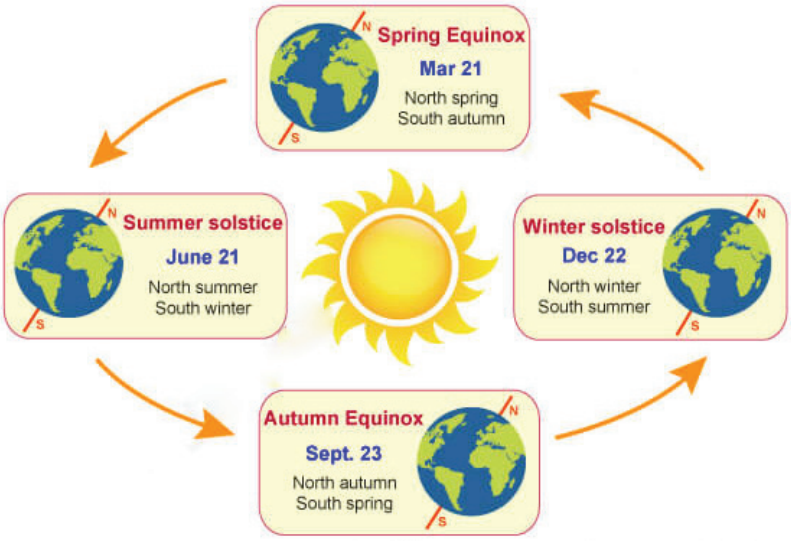


Figure 6.10 Revolution of the Earth showing solstice and Equinox .

Earth's seasons changes because its axis tilted by 23.50 relative to Earth's orbit around the sun. Thus, the sun's rays hit different parts of the planet more directly depending on the time of year.

- From June to August, the sun's rays hit the Northern Hemisphere more directly than the Southern Hemisphere. The result is warm (summer) weather in the Northern Hemisphere and cold (winter) weather in the Southern Hemisphere.
- From December to February, the sun's rays hit the Northern Hemisphere less directly than the Southern Hemisphere. The result is cold (winter) weather in the Northern Hemisphere and warm (summer) weather in the Southern Hemisphere.
- From September to November, the sun shines equally on both hemispheres. The result is fall in the Northern Hemisphere and spring in the Southern Hemisphere.

- The sun also shines equally on both hemispheres from March to May. The result is spring in the Northern Hemisphere and fall in the Southern Hemisphere.

The table 6.1 given below provides the basic differences between rotation and revolution.

Rotation	**Revolution**
Rotation is the movment of Earth on its axis.	evolution is the movement of the Earth around the Sun.
The Earth takes 24 hours to complete a rotation.	The Earth takes 365 days for one complete revolution around the Sun
The imaginary line that passes through the center of the Earth and the poles is called an axis.	The path of the Earth moving around the Sun is called an orbit. The Earth's orbit is elliptical.
The rotation of earth on its axis causes day and night.	The revolution of earth around the sun causes seasonal change.

Exercise 6.5

I. Fill the blank space

1. Earth _____ around the sun.
2. Day and night are the result of Earth's _____
3. Seasons are the result of Earth's _____ around the sun.

II. Choose the best answer from the given alternatives.

1. Which of the following is NOT a factor affecting Earth's seasons or climate?
   A. Earth's orbital revolution around the sun.
   B. Earth's axis tilt.
   C. Earth's distance from the sun.
   D. Earth's wind and ocean current patterns.
2. How long does the Earth take to complete one rotation on its axis?
   A. 24 hours B. 12 hours C. 1 hour D. 6 hours

6.4 Systems and Cycles (effects, measurement ideas/estimation)

At the end of this section you will be able to:

- identify atmospheric and lithospheric systems;
- explain their cycle effects of the Earth;
- describe the measuring techniques for too big (Earth) and to small (continental drift) quantities measurement and estimation.

Activity 6.5

Go to your library and read about four earth systems. the definitions of each earth system. Then Form a group discuss your findings and present the following systems to the class.

a. Atmosphere
b. Lithosphere
c. Hydrosphere
d. Biosphere

6.4.1. Earth's Systems

Earth consists of land, air, water and life. The land contains mountains, valleys and flat areas. The air is made up of different gases, mainly nitrogen and oxygen. The water includes oceans, lakes, rivers, streams, rain, snow and ice.Life consists of people, animals and plants. There are millions of species, or kinds of life, on Earth. Their sizes range from very tiny to very large.

Earth's parts once were seen as largely separate from each other. Now they are viewed together as the "Earth system." Each part connects to and affects each of the other parts.

For example:

- Clouds in the air drop rain and snow on land.
- Water gives life to plants and animals.
- Volcanoes on land send gas and dust into the air.
- People breathe air and drink water.

The Main Components of the Earth System

The earth system is an integrated system, but it can be subdivided into four main components, sub-systems or spheres. Thess are geosphere, atmosphere, hydrosphere and biosphere. These components are also systems in their own right and they are tightly interconnected.

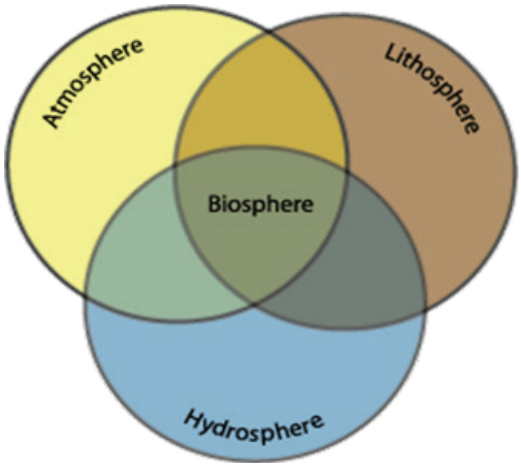


Figure 6.11 Earth Systems interactions

The four main components of the earth system are:

1. The geosphere (lithosphere): - this is the part of the Earth composed of rock and minerals. It includes the solid crust, the molten mantle and the liquid and solid parts of the earth's core. In many places, the geosphere develops a layer of soil in which nutrients become available to living organisms.

The surface of the geosphere is subject to processes of erosion, weathering and transport, as well as to tectonic forces and volcanic activity, which result in the formation of landforms such as mountains, hills

2.The atmosphere - this is the gaseous layer surrounding the earth and held to its surface by gravity. The atmosphere receives energy from solar radiation which warms the earth's surface and is re-emitted and conducted to the atmosphere.

The atmosphere also absorbs water from the earth's surface via the process of evaporation; it then acts to redistribute heat and moisture across the earth's surface. In addition, the atmosphere contains substances that are essential for life, including carbon, nitrogen, oxygen and hydrogen.

3. The hydrosphere - this consists of those parts of the earth system which composed of water in its liquid, gaseous (vapour) and solid (ice) phases.

4. The biosphere - this contains all living organisms and it is intimately related to the other three spheres: most living organisms require gases from the atmosphere, water from the hydrosphere and nutrients and minerals from the geosphere.

Living organisms also require a medium for life, and are adapted to inhabit one or more of the other three spheres.

These are described in the table 6.2 below

Earth spheres	Descriptions
1. Geosphere (lithosphere)	all of the rocks and "hard parts" of the Earth
2. Hydrosphere	all of Earth's water
3. Biosphere	all living things
4. Atmosphere	the blanket of gases surrounding the planet

6.4.2. Cycles of the Earth System and Effects

Project work6.3

Make a group and refer encyclopedia, internet, in reference books, or in other resource material about water cycle, carbon cycle and nitrogen cycle. Write a report about your findings and present to a class.

What is cycle?

Cycles are sequences of events that repeat themselves in the same order. Cycles that exchange materials among living and nonliving components of the Earth are known as biogeochemical cycles. The key materials that cycle through the major biogeochemical cycles are water cycle ,carbon cycle, nitrogen cycle, of which are essential for life.

Water Cycle

The water cycle involves three main phases, related to the three states of water: solid, liquid, and gas. Ice, or solid water, is most common near the poles and at high altitudes. Ice sheets and glaciers hold the most solid water. There are four main stages in the water cycle. These are evaporation, condensation, precipitation and transpiration.

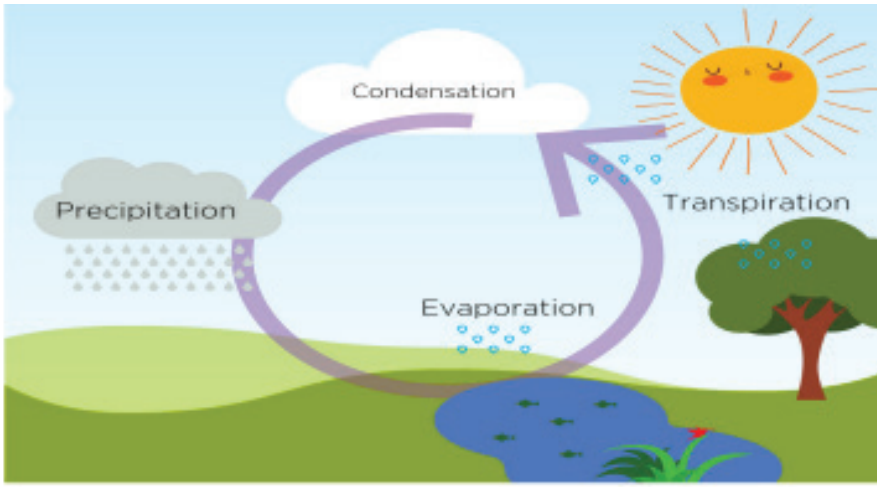


Figure 6.12. Water Cycle

Evaporation:- is when warmth from the sun causes water from oceans, lakes, streams, ice and soils to rise into the air and turn into water vapor (gas). Water vapor droplets join together to make clouds.
Condensation: - is when water vapor in the air cools down and turns back into liquid water.
Precipitation:- is when water (in the form of rain, snow, hail or sleet) falls from clouds in the sky.
Transpiration : is the process of water movement through a plants.

The carbon cycle

The carbon cycle describes the process in which carbon atoms continually travel from the atmosphere to the Earth and then back into the atmosphere.

Steps of carbon cycle

1. CO_2 Used by Plants for Photosynthesis
2. Consumption by Animals
3. Ocean Intake
4. Decay and Decompose
5. Formation of Fossil Fuels
6. Use of Fuels for Industrial Purposes
7. Carbon Emissions
8. Respiration by Plants and Animals

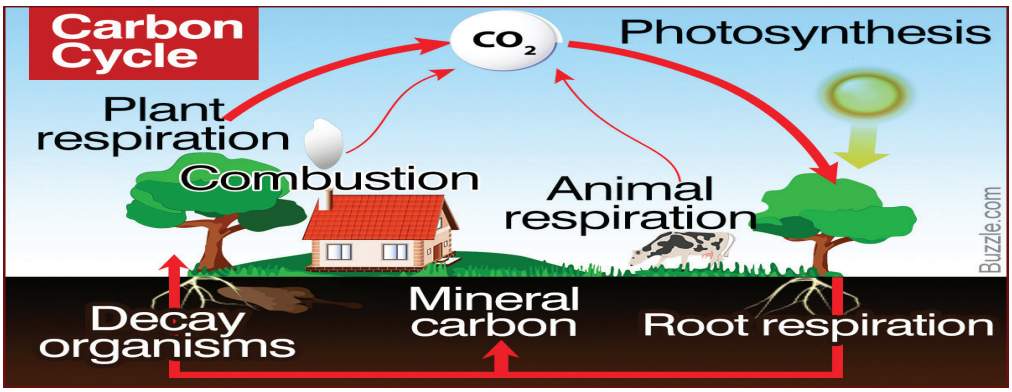


Figure 6.13. Carbon cycle

The Nitrogen Cycle

The process by which the balance of nitrogen is always maintained by the exchange of nitrogen between living world and the environment is known as nitrogen cycle. **Steps of nitrogen cycle:** Nitrogen cycle has five steps, such as:

a) Nitrogen fixation b) Nitrification c)Assimilation
d) Ammonification e) De-Nitrification

Except assimilation all the four other types are completed by bacteria.

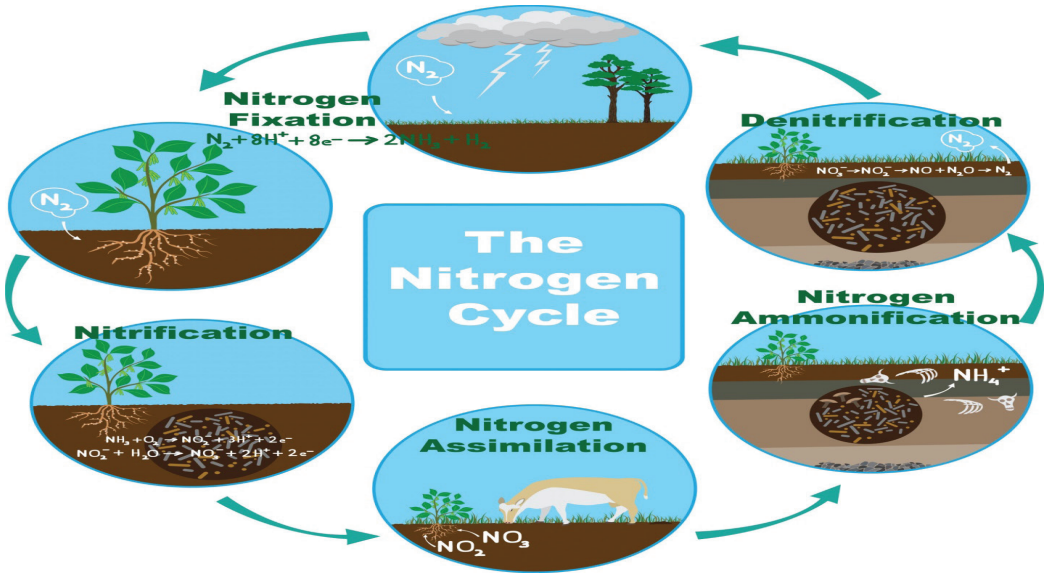


Figure 6.14. Nitrogen cycle

Measuring techniques of Earth and Continental drift

Plate motions are best measured by satellite-based methods. A number of techniques are used to measure the size of Earth and continental drifts. One of these methods is using Global Positioning System **(GPS).**

GPS is a network of satellite and receiving devices used to determine the location of something on Earth. If you could have two such receiving devices, at two different points on Earth for example the two ends of continental drift, you can easily determine the distance between them using **GPS.**

Another device used for measuring size of Earth and continental drifts is known as **Laser ranging (laser radar).** It is an optical sensing technology

Exercise 6.6

I. Choose the best answer from the given alternatives.

1. All living things on earth are part of the_____.
   A. Biosphere C. Geosphere
   B. Hydrosphere D. Cryosphere
2. The Geosphere is
   A. all liquid water on Earth. C. 100km thick layer of gases.
   B. frozen water on earth. D. Earth's entire solid body.
3. Fresh water is important for life on Earth because:
   A. without fresh water sources people will get sick and die.
   B. Plants need water to grow.
   C. the biosphere needs water.
   D. All of the above.
4. This photo is mainly showing an example of the:

   A. Hydrosphere
   B. Geosphere
   C. Atmosphere
   D. Biosphere

5. An example of a connection between Atmosphere and Hydrosphere is:
   A. rivers B. rain C. lakes D. rocks

Key Terms of a Unit

- Earth science	- Atmosphere	- Geosphere	- Biosphere
- Earth	- Troposphere	- lithosphere	- Hydrosphere
- Inner core	- Stratosphere	- revolution	- Core
- Geologists - AU	- Mesosphere - Meridian	- rotation	

- Lunar eclipses	- Exosphere	- Crust	
- Oblate spheroid	- Thermosphere	- Mantle	

Unit Summary

- Earth is the only planet in the universe to support life. It is about 150 million kilometers far from the sun. This distance, called an astronomical unit (AU), is a standard unit of measurement in astronomy.
- Earth is an oblate spheroid. This means it is spherical in shape, but not perfectly round. The geoid describes the model shape of Earth, and is used to calculate precise surface locations.
- The equatorial circumference of the Earth is 40,075 km. This is the distance around the equator of the Earth. If you measure the circumference of the Earth, while passing through the poles, the distance is only 40,008 km. This is because the Earth is not a perfect sphere. It's rotating rapidly, which causes the equator to bulge out.
- The equatorial diameter of the Earth is 12,756 km. This is the diameter of the Earth measured from one side of the Earth, passing through the center. If you go from pole to pole through the center, the distance is only 12,714 km.
- The 4 components of the Earth subsystems are called "spheres." Specifically, they are the "lithosphere" (land), "hydrosphere" (water), "biosphere" (living things), and "atmosphere" (air).
- Earth's interior is divided into three major layers: the crust, the mantle, and the core. Each layer has a unique chemical composition, physical state, and can impact life on Earth's surface.
- Rotation and Revolution are two motions of the earth. When earth spins or rotates around its axis, that movement of spinning is called ***Rotation of Earth.*** Besides, when Earth spins or revolves around the sun, that movement is called Revolution of Earth.
- The motions of the earth have its own effects. These are: the main effects of the Earth's rotation are a diurnal cycle of light and darkness,

i.e. day and night, rise and fall of the sea level twice a day, sunrise in the east and sunset in the west. Effects of Earth's revolution include change of the seasons and variation in the length of days and nights.

REVIEW EXERCISE

I. Write True if the statement is Correct and False if the statement if incorrect.

1. Atmosphere is the thin layer of gases held on the earth by gravitation attraction.
2. Earth rotates around the sun.
3. Day and night are the results of Earth's rotation.
4. We live on the core of the Earth.

II. Match the following earth spheres with their related meaning

Column 'A'	**Column 'B'**
1. Lithosphere	A. all of Earth's water
2. Hydrosphere	B. all living things
3. Atmosphere	C. all of the rocks and "hard parts" of the Earth
4. Biosphere	D. the blanket of gases surrounding the planet

III. Fill the blank space

Use these words to fill in the blanks next to the sentences below.

Words	365.25 days Season	Revolution Rotation	24 hours Axis

1._____is the amount of time for Earth to make a complete rotation.

2._____is the process of Earth spinning on its axis.

3._____ is the amount of time it takes Earth to completely orbit the sun.

4._____is the process of Earth orbiting the sun.

5._____is an imaginary line that runs through the center of Earth from the North Pole to the South Pole.

6. _____is term used to describe a certain time of year.

IV. Choose the best answer from the given alternatives

1. Which of the following is an example of Biosphere connecting to Atmosphere?
   A. plants produce oxygen C. animals live in caves
   B. animals eat plants D. animals drink water
2. What is Earth's outermost system?
   A. Atmosphere C. Cryosphere
   B. Hydrosphere D. Geosphere
3. The earth's four systems are:
   A. independent from one another C. all part of the atmosphere
   B. all connected D. not important for life on earth
4. Photosynthesis is an example of an interaction between the biosphere end the:
   A. Atmosphere C. Cryoshpere
   B. Geosphere D. hydrosphere
5. What is Earth's largest system?
   A. Biosphere C . Hydrosphere
   B. Geosphere D. Atmosphere

6. The Earth is slightly flattened from a perfect spherical shape because of

A. its rotation C. the pull of the sun and moon

B. storms on the sun's surface. D. its molten core

7. As altitude within the troposphere increases, the amount of water vapor generally

A. decreases, only C. remains the same

B. increases, only D. decreases, then increases

8. Oxygen is the most abundant element by volume in Earth's

A. Hydrosphere C. mantle

B. crust D. inner core

9. An observer watching a sailing ship at sea notes that the ship appears to be "sinking" as it moves away. Which statement best explains this observation?

A. The Earth is revolving.

B. The Earth is rotating.

C. The Earth has a curved surface.

D. The surface of the ocean has depressions.

10. Which one of the following Nitrogen cycle Not completed by bacteria?

A. Nitrogen fixation B. Nitrification

C. Assimilation D. Ammonification

11. What is the approximate elevation of the stratopause?

A. 10 km B. 80 km C. 30 km D. 50 km

12. The best evidence that the Earth has a spherical shape is provided by

A. photographs of the Earth taken from space satellites.

B. the amount of daylight received at the North Pole on June 21.

C. the changing orbital speed of the Earth in its orbit around the Sun.

D. the cyclic change of seasons.

13. One of the following is Not the stapes of water cycle?

A. evaporation C. de-nitrification

B. precipitation D. transpiration.

V. Short answer question

1. Write the Earth's atmosphere from lowest to highest in altitude.
2. Name and explain the three layers of Earth?
3. What are the two movements of the Earth?
4. What are effects of the Earth's rotation?
5. What is the role of the carbon cycle?
6. What are the effects of revolution of Earth?

Unit 7 MOTION, FORCE, ENERGY AND ENERGY RESOURCES

Learning Outcome

At the end of this unit, you will be able to:

- describe the term motion;
- identify types of motion (motion on straight line, circular motion, rotary motion and curvilinear motion);
- show those types of motion in the class;
- explain the term force;
- demonstrate the pulling/pushing activity of force;
- explain gravitational force;
- List all effects of force;
- demonstrate some effects of force;
- relate effects of force with their daily life experience;
- name measuring device of force;
- identify different measuring scales on measuring device of force;
- explain parts of measuring device of force;
- define energy as a property of matter that can be converted;
- list all forms of energy;
- explain which energy converted to other forms of energy;
- list sources of energy;
- distinguish between renewable and non-renewable forms of energy;
- describe how energy is used wisely;
- list the strategies of conservation of energy;
- explain resource depletion and environmental degradation.

Contents

Section	Learning competencies
7.1 Definition and types of motion	• Define motion as the change of position with time. • Describe the types of motion. • Give examples for each type of motion
7.2 Definition of force and gravitational force	• Explain the term force. • Demonstrate the pulling/pushing activity of force. • Explain gravitational force.
7.3 Effects of force	• List all effects of force • Demonstrate some effects of force. • Relate effects of force with their daily life experience
7.4 Measuring forces	• Name measuring device of force • Identify different measuring scales on measuring device of force • Explain parts of measuring device of force
7.5 Definition of Energy (Property of matter can be converted)	• Define energy as a property of matter that can be converted
7.6 Forms and Conversion of Energy	• List all forms of energy • Explain which energy converted to other forms of energy.
7.7 Energy Sources (sun, fuel, hydroelectric, wind, nuclear)	• List sources of energy. • Distinguish between renewable and non-renewable forms of energy.
7.8 Wise use & Conservation	• Describe how energy is used wisely. • ist the strategies of conservation of energy
7.9 Resource depletion & environmental degradation	• Explain resource depletion and environmental degradation

7.1. Definition and types of Motion

At the end of this section you will be able to:

- define motion as the change of position with time;
- describe the types of motion;
- give examples for each type of motion.

Introduction

In this unit, you will be introduced to the basic concepts of motion, force, energy, forms and conservation of energy, wise use and conservation of energy, resource depletion and environmental degradation.

7.1.1. Definition of Motion

When you go to your school, your journey begins from home. Your home is your original position and your school is your final position. While you are going from home to school, you are increasing the gap between your present position and your home.

This continuous change of position is known as a motion. Notice that your change of position is observed by considering the distance from your school to home. Your home is taken as a reference frame.

Activity 7.1

Form a group and perform the following task; present your finding to the class.

1. Define motion by your own word and give examples.
2. When you move in bus, describe are you at rest or in motion with respect to
   i. the bus seat
   ii.the ground

Now let us understand motion clearly with the help of a few Examples:

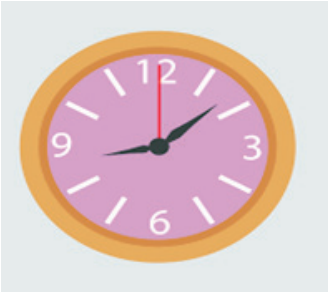
- Our daily activities, like walking, running, closing the door, etc. involve motion. There is a change of position of the object involved in these activities.
- When an airplane travels from Addis Ababa to Dire Dawa is an example of motion.

- The automobiles that carry passengers from one place to another possess motion. In this case, the position of passengers is changed from one place to another.

Activity 7.2

Form a group and perform the following tasks

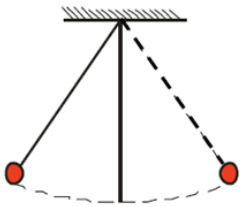
i. Observe the motions indicated in Fig 7.1.
ii. Have you noticed any difference between the motions in Fig 7.1 (a-e)? Describe them.
iii. Group these motions based on their path.



a. Wrists watch.

b. A car moving on a straight line

C. Roundabout



d. Simple pendulum

e. Motion of kids swing

Figure 7.1 different types of motion

In Fig 7.1 (a) you observe that the motion of second or minute hand of a wrist watch about an axis. In Fig 7.1 (b) you observe that a car is moving on a straight road. Its path is a straight line. Fig 7.1(c) shows that the path of the moving car is a curved line. While Fig 7.1 (c and e) show the 'to and fro' motions of an object.

According to the nature of the movement. or based on the path followed, motion is classified into four types as follows:

1. Rectilinear Motion
2. Curvilinear Motion
3. Oscillatory Motion
4. Rotary Motion

1**. Rectilinear Motion:** Motion in a straight line is called rectilinear motion. In other words, when an object moves along a straight line path, it is called **rectilinear motion.**

Example:

- A boy walking on astraight road.
- A car moving in a straight road.
- A falling ball from a certain height.
- A boy pulling a toy towards him. etc.

2. Curvilinear motion

The motion of a particle or object moving along a curved path is called curvilinear motion.

Examples: Motion of a car around a circular path,

- The motion of a ball thrown horizontally from a certain height.
- The motion of the moon around the earth.
- Motion of a basketball into the basket, etc.

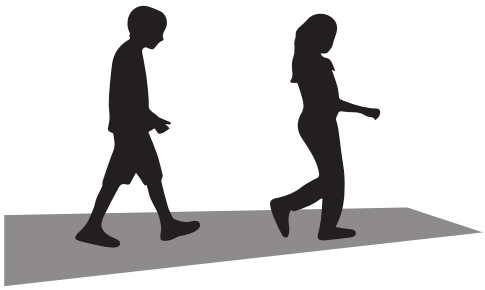


Figure 7.2 a boy and girl walking on a straight road

Figure 7.3 Curvilinear Motion of a basketball into the basket

Note: Circular motion is a special case of curvilinear motion, in which the body moves along a circular path.

Examples of circular motion are:

- A bicycle or a car moving on a circular track of park.
- The motion of the moon around the earth etc.

3. Rotary Motion:

A type of circular motion where an object spins on its own axis, it is called **rotational motion.**

Example:

- rolling ball,
- Spinning top and
- the motion of the second or minute hand of a wrist watches
- Movement of the earth on its axis.

4. Oscillatory Motion

A repeating motion in which an object continuously repeats in the same motion again and again is called Oscillatory Motion. It is also a to and fro, back and forth or up and down motion. Some of the common examples of oscillatory motion are:

- A swinging swing
- The motion of a pendulum
- A boat tossing up and down on a river

Figure 7.4 rotating wooden spinning top

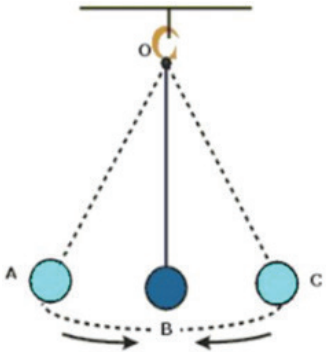


Figure 7.5 motion of a pendulum

Exercise 7.1

I. Choose the best answer from the given alternatives.

1. Which of the following is a type of motion?
   A. rectilinear motion B. rotary motion
   C. oscillatory motion D. all of these
2. Which motion type describes the motion of pendulum of a clock?
   A. linear motion C. rotary motion
   B, oscillatory motion D. all of these
3. The spinning of the body about its axis is_____
   A. rotatory C. translational
   B. circular D. vibratory

4. Which type of motion is “the pendulum of a wall clock moves at regular intervals”?
   A. Rectilinear C. Rotatory
   B. Vibratory D. B and C
5. Which type of motion is “a train moving on a straight track “?
   A. Circular C. Rotary
   B. Rectilinear D. none of the above
6. The act, process or state of the change in place or position of a body with respect to time and relative to the observer is said to be_____.
   A. Rest C. Motion
   B. Stationary D. none of the above

II. Short answer questions

1. What is a motion?
2. State at least four types of motion, and give practical examples for each type

7.2. Definition of Force and Gravitational Force

Definitions of force

At the end of this section you will be able to:

- explain the term force;
- demonstrate the pulling/pushing activity of force;
- explain gravitational force.

Activity 7.3

Form a group and perform the following activities Share your opinion to the whole class.

i. What is a force?
ii. Mention some examples of forces from your daily activities.
iii. Explain the following actions.

- A push you exert on a wall,
- A pull exerted to drag a box on a table.

All of us are familiar with the word force as we use it in our everyday life. Let us use to describe interactions between different bodies in nature.

For example when you kick a ball, tear a paper, bend a wire, hold a bag, walk on the floor, close and open a door, you apply a force. A **force** is a push or pull upon an object resulting from the object's interaction with another object. Whenever there is an interaction between two objects, there is a force upon each of the objects. Forces influence objects that are at rest or that are already in motion. It can also be defined as an external agent which can change the state of rest or motion of a body.

Figure 7.6(a) When a horse pulling a cart.

Figure 7.6(b) When a wagon pushed.

Forces are classified into two broad categories . These are

(i). Contact forces and

(ii). Non-contact forces

i. Contact forces are forces exerted when two objects are in touch or contact.

For example; Spring force, frictional force, collision force etc

ii. Non-contact forces are forces exerted without body contact. They are forces acting at a distance. Gravitational force, magnetic force and electrical forces.

Gravitational force

ACTIVITY7.4
Form a group and discuss the following ideas. Present your discussion to the class.
Throw a ball vertically upward and observe its motion. What will happen to the ball? Will it continue to move upward ?

We live on Earth. It is difficult to get away from earth. If you jump upwards, you fall back down again. The earth's gravity pulls you down wards. The earth's gravity causes a force that pulls any object down wards. This force is called weight (gravitational force). Gravity always pulls you towards the center of the Earth. It doesn't matter where you are on the surface of the earth. Since, a freely falling body in the air moves down irrespective of its mass. This is due to force of gravity. Example:

- Falling of fruits from trees due to Earth's gravitational pull
- The Earth's gravitational pull keeps us all stationary; otherwise, we all would be flying now.
- Revolution of the Earth around the Sun.
- Revolution of the moon around the Earth.

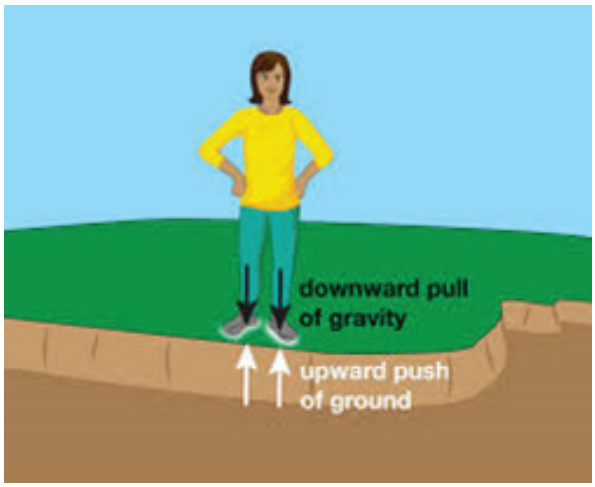
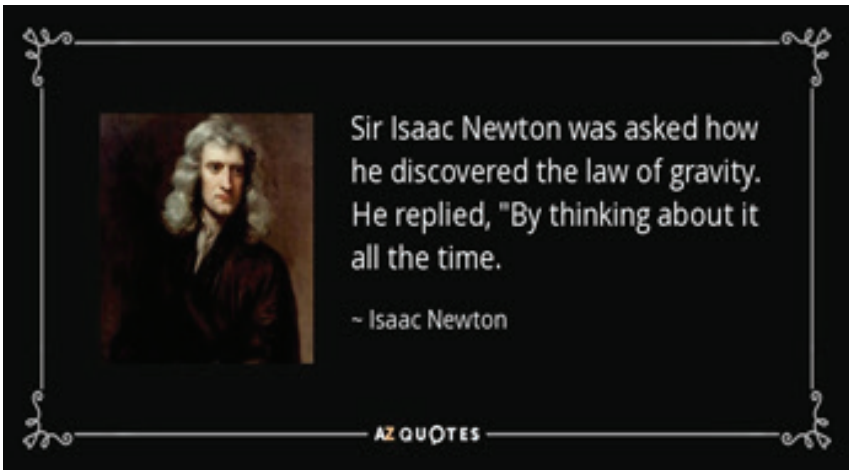


Figure 7.7. Our weight is caused by the pull of the earth's gravity

Sir Isaac Newton was an English mathematician and physicist who lived from 1642-1727. The legend is that Newton discovered Gravity when he saw a falling apple while thinking about the forces of nature. Whatever really happened, Newton realized that some force must be acting on falling objects like apples because otherwise they would not start moving from rest. Newton called this force "gravity" and determined that gravitational forces exist between all objects.



Note

- Force is a push or pull of an object.
- Gravity is attractive force between objects with mass.
- Gravity always a pull never a push.
- Gravitational force is the force of attraction among all masses in the universe; especially the attraction of the earth's mass for bodies near its surface.

Exercise 7.2

I. Fill in the blank spaces with the appropriate word(s).

1._____is the force of attraction that acts between all objects in the universe, without exception.

2. Push or pull of an object in a certain direction is known _____.

II. Choose the best answer from the givenalternatives.

1. The force of gravity on a person or object at the surface of a planet is called

   A. Mass B. Gravity C. weight D. motion

2. The force that pulls objects toward Earth is called _____.

   A. mass B. gravity C. air D. wind

3. Which one of the following best describes a gravitational force?

   A. A repulsive force between any two objects with mass.

   B. A force of attraction between two objects with mass.

   C. A force between any two objects, whether or not they have mass.

   D. All of the above.

III short answer questions

1. Define the term Force as a science with appropriate examples from your daily life.

2. Explain gravitational force.

7.3. Effects of Force

At the end of this section you will be able to:

- list all effects of force;
- demonstrate some effects of force;
- relate effects of force with their daily life experience.

Activity 7.5

Observe the activities shown in Figure 7.8

a. Explain the effects of forces in each activity.

b. List and show other effects of force in the class in front of the students.

a, When force is applied on flour dough, it changes its shap

b, The force of brakes can stop a moving car

c, The soccer kicks ball

d, A goal keeper changes the direction of motion of a ball

Figure 7. 8. Different effects of forces

Force has the following main effects, when it is exerted on an object:

- A force can move a stationary object

  Example: when a force is applied to a stationary tennis ball, it will make the ball continue its motion in the direction of the applied force.
- A force can change the speed of a moving object.

  Example: When we keep on pedaling the pedal of the bicycle, the speed of the bicycle increases. And when we apply the brake, the speed of the bicycle decreases.
- A force can either stop or slow down the moving object.

  Example: The force of brakes can stop a moving car.
- A force can change the direction of a moving object.

  Example: the direction of the moving football can be changed by applying force at an angle.
- A force can change the shape of an object.

  Example: When force is applied on flour dough, it changes itsshape shape.

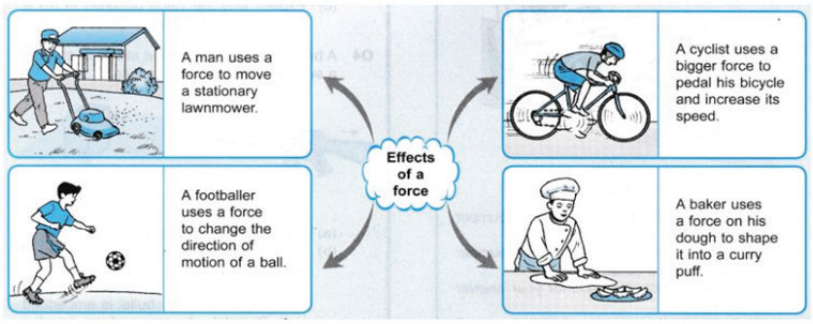


Figure 7.9 . Some Effect of force on daily activities

How effect of force is related with the daily life?

We are very familiar with the various effects of force in our everyday life. Like, we walk, we run, we play, and we sit, and we stand. In all of these activities, we have all seen the changing of the motion of the state.

Some additional examples are:

- Pushing a wheel barrow.
- Opening or closing a door.
- Squeezing wet clothes.
- Gravitational force.
- Brakes applied to stop a moving vehicle.
- A football kicked
- Rubbing a glass rod with a silk.
- Running and Pushing etc.

Exercise 7.3

I. Choose the best answer from the given alternatives.

1. Which one of the following is the effects of force, when it is exerted on an object?
   A. Change the speed of a moving object.
   B. Move a stationary object.
   C. Change the shape of an object.
   D. All of the above.
2. If two forces from opposite direction are applied on an object and the object is malleable, it will_____.
   A. Change in color C. Change in direction
   B. Change in motion D. Change in shape
3. To squeeze toothpaste or press a lift button we apply_____.
   A. heat C. force
   B. signal D. stationary
4. A force can act on a stationary object and can cause it to____.
   A. Move C. develop
   B. Grow D. stationary

II. short answer questions

1. List and describe some effects of a force.

7.4. Measuring Forces

At the end of this section you will be able to:

- name measuring device of force;
- identify different measuring scales on measuring device of force ;
- explain parts of measuring device of force.

Activity 7.6

Discuss the following activity in your group and present your discussion to the class.

1. What is the instrument used to measure a force?
2. Explain parts of measuring device of force.
3. Can we use spring balance to measure mass when a shopkeeper measure banana, orange or others?

In science, if we want to know that one force is bigger than another we do not simply guess; we make measurements. How can we mea sure forces?

To measure the amount of force exerted on an object we use an instrument called Newton meter (force meter) also called spring balance. But **Newton mete**r is the scientific instrument used to measure a force.

Parts of Spring balance

- **Spring balance:** Scale made up of a hook attached to a spring that stretches in proportion to the weight of the object being weighed.
- Hook: Curved part on which the body to be weighed is hung.
- **Graduated scale:** The divisions of equal length that are marked on the spring balance and constitute the units of measurement.
- **Pointer:** Pointer connected to the spring that moves along a graduated scale to indicate the weight of the body being weighed

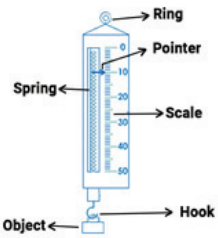


Figure 7.10. Spring balance

Figure 7.11 shows how to measure the force needed to pull a block of wood along the bench. The following steps are used in measuring a force.

- Check that the force meter reads zero before you start.
- Attach the hook of the force meter to the block.
- Hold the ring at the other end of the force meter and pull the block.
- Read the value of the force from the scale

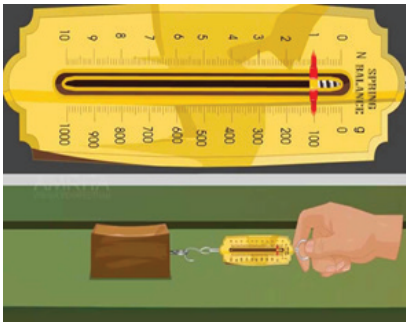


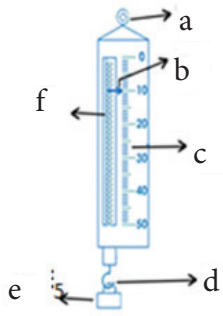
Figure 7.11 Measuring pull of force

The spring stretches when a force is applied to the hook. The bigger the force applied, the longer the spring stretches and the bigger the reading.

Exercise 7.4

I. Fill in the blank spaces with the appropriate word(s).

1. A force is measured using an instrument called _____.
2. _____is a curved part of force meter on which the body to be weighed is hung.
3. The diagram shows parts of newton meter; Name the parts of the arrow that are represented by letter.



a. _____
b. _____
c. _____
d. _____
e. _____
f. _____

7.5. Definition of Energy (Property of matter can be converted)

At the end of this section you will be able to:

- define energy as a property of matter that can be converted.

Project work 7.1

By using internet or other reference materials perform the following tasks in group and present your findings to the class.

i. Define energy as a property of matter.

ii. Explain with examples how energy and matter are related.

In unit two of this book, you learned important concepts about matter. In this section, you will learn the concepts of energy. Matter is any thing that have mass and occupies space..

Energy is a very common word frequently used in our day-to-day life. **Energy** is the property of matter, and it comes in many forms, such as heat, sound, light, and motion. It can be transferred between objects, and converted in form. It cannot be created or destroyed.

Examples of energy and matter

- A raindrop falling from the sky is made of matter (water), and it has potential, kinetic, and thermal energy.
- A light bulb is made of matter, and it emits energy in the form of heat and light.
- The wind consists of matter (gases in air, dust, pollen), and it has kinetic and thermal energy.

Exercise 7.5

I. Fill in the blank spaces with the appropriate word(s).

1._____ is the property of mater, and it comes in many forms.
2._____is any thing that have mass and occupy space.
3. Alit light bulb is made of matter, plus it emits energy in the form _____ and _____

7.6. Forms and Conversion of Energy

At the end of this section you will be able to:
- list all forms of energy;
- explain which energy converted to other forms of energy.

Forms of energy

The world we live in provides us with many different forms of energy. **Examples** of these are: light energy, heat energy, mechanical energy, gravitational energy, kinetic energy, potential energy, thermal energy, and electrical energy, sound energy, chemical energy, nuclear or atomic energy, elastic potential energy and so on. We can think energy coming in different forms, some for storing and some for transferring.

Activity 7.7

Discuss the following questions in a group and present your opinion to the whole class.

i. List the different forms of energy.
ii. Define conversion of energy and give examples which energy is converted to other forms of energy.

The table 7.1 shows some different forms of energy and their descriptions.

Form of energy	Description
Chemical energy	Energy of a chemical substance.
Elastic energy	Energy of a stretched or squashed object
Electrical energy	Energy carried by electricity
Gravitational potential energy	Energy of an object that has been lifted
Heat energy	Energy spreading out from a hot object
Kinetic energy	Energy of a moving object
Light energy	Energy spreading out from a bright object
Sound energy	Energy coming from a vibrating source

Energy conversions

Energy can be changed from one form to another. The process of changing energy from one form to another form is called energy conversion. A very common energy conversion is a change from gravitational potential energy to kinetic energy. This occurs whenever an object falls due to the force of gravity. Each form can be converted or changed into the other forms. The notion of energy is that energy is changed from one form into different forms using transducers. Transducer is a device used to transform energy from one form to another.

For example:

1. Battery converts chemical energy into electrical energy.
2. A generator converts mechanical energy into electrical energy.
3. A motor converts electrical energy into mechanical energy.

The Table7.2 Summarizing Energy Conversion from one form to another

Original energy	**Transducer**	**Energy transformed**
Chemical energy	Battery	Electrical energy
Mechanical energy	Generator	Electrical energy
Solar energy	Solar panel	Electrical energy
Chemical energy	Motor car	Mechanical energy
Electrical energy	Microphone	Sound energy
Electrical energy	Heater	Heat energy

Exercise 7.6

i. Choose the best answer from the following alternatives

1. One of the following is not a form of energy?

   A. Light B. sound C. Kinetic D. weight

2. The process of changing energy from one form to another form is called

   A. energy conversion C. energy depilation

   B. energy conservation D. energy degradation

3. Which One of the following is a form of energy?

   A. Chemical B. Solar C. Electrical D. all of the these

ii. Complete the table

1. Energy can be transferred in different ways. Copy the table and use words from the list to complete the first column.

   Chemical energy, Electrical energy, Mechanical energy

Original energy	Transducer	Energy transformed
	Microphone	Sound energy
	Generator	Electrical energy
	Battery	Electrical energy

7.7. Energy Source (sun, fuel, hydroelectric, wind, nuclear)

At the end of this section you will be able to:

- list sources of energy;
- distinguish between renewable and non-renewable forms of energy.

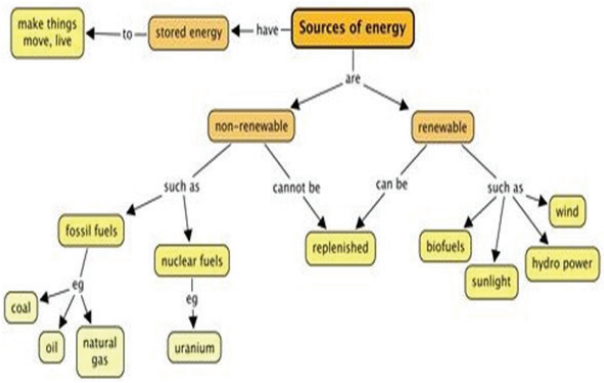


Figure 7.12 Source of energy

Activity 7.8
Discuss the following activities within a group and present your discussion to the class.
1. List out any five activities from your daily life in which different forms of energy are involved.
2. Differentiate between renewable and non-renewable sources of energy.

In simple terms we can say that anything out of which usable energy can be extracted is a source of energy. There is a variety of sources that provide us energy for different purposes. Some of them are coal, petrol, diesel, kerosene, natural gas, hydroelectric power, wind mills, solar panels, and biomass etc.

The energy sources can be replenished in a short period of time are referred to as "renewable" energy sources, whereas the energy sources that we are using up and cannot be generated in a short period of time are called non-renewable energy sources. Thus, all the sources of energy can be divided into two categories: renewable sources and non-renewable sources of energy.

Renewable energy sources: - are the energy sources, which can be turned into use again after being used. It comes from natural sources and continually regenerates themselves, which makes them nearly inexhaustible. These energy sources are plentiful, sustainable, naturally replenished and good to the environment.

The major types or sources of renewable energy are:

- Solar energy from the sun
- Wind energy
- Geothermal energy from the heat inside the earth
- Hydropower from flowing water
- Biomass from plants
- Ocean energy in the form of wave, tidal, current energy and ocean thermal energy.

Non- renewable sources: are the energy sources, which cannot be turned into use again. It is a finite resource. It is a natural substance that is not replenished with the speed at which it is consumed. These are formed over thousands of years from the buried remains of ancient sea plants and animals that lived millions of years ago. Most of these energy sources are "dirty" fossil fuels, which are generally dingier for the environment.The major types or sources of non-renewable energy are: petroleum, natural gas, coal, nuclear energy and hydrocarbon gas liquids

Table 7.3 major differences between renewable and non-renewable resources

Renewable Resources	**Non-renewable Resources**
Depletion	
Renewable resources cannot be depleted over time	Non-renewable resources deplete over time
Renewable resources include sunlight, water, wind and also geothermal sources such as hot springs and fumaroles	Non-renewable energy includes fossil fuels such as coal and petroleum.
Enviromental Impact	
Most renewable resources have low carbon emissions and low carbon footprint	Non-renewable energy has a comparatively higher carbon footprint and carbon

Cost	
The upfront cost of renewable energy is high. – For instance, Generating electricity using technologies running on hydro power is costlier than generating it with fossil fuels	Non-renewable energy has a comparatively lower upfront cost.

Renewable and non-renewable resources have many similarities They both are resources and they both have to do with the environment, they both grow on Earth, as well. Thus, we must use them wisely.

Exercise 7.7

I. Choose the best answer from the given alternatives

1. Which of the following is a nonrenewable energy resource?
   A. Solar B. hydroelectric C. wind D. coal
2. What type of energy is derived from heated groundwater?
   A. solar energy C. geothermal energy
   B. hydroelectric energy D. nuclear energy
3. Which of the following is a renewable energy resource?
   A. Solar B. Biomass C. Geothermal D. All

II. Short answer questions

1. List sources of energy.
2. What is the difference between renewable and non-renewable forms of energy and give four examples for each.

7.8. Wise use and Conservation of Energy

At the end of this section you will be able to:

- describe how energy is used wisely;
- list the strategies of conservation of energy.

Energy conservation

The key for resolving the country's energy crisis lies on us with citizens. Among things we can do is the conservation of our energy sources. It is said that energy saved is as good as energy generated.

Therefore, we should not only judiciously use energy sources but save energy as much as we can. You can start conservation of energy from your home.

Energy conservation is the practice of using less energy in order to lower costs and reduce environmental impact. energy can be conserved by

- Reducing wastage and losses,
- Improving efficiency through technological upgrades and
- Improved operation and maintenance

the good rule to follow for conservation of energy use the following three (3) R's:

- **Reduce:** Reduce the amount of trash you produce and the amount of energy that are consumed. This is the best way to conserve natural resources and reduce pollution.
- **Reuse:** By products that you can use more than once. Try to avoid disposable items that use up natural resources and produce extra trash.
- **Recycle:** Recycling is the process that reuses and changes used materials into things that can be of use. Although it requires energy to recycle things, overall, recycling saves energy as well as landfill space and reduces our need for more natural resources. Lots of things can be recycled, Like plastic, metal, glass, paper, and

Project work 7.2

By referring internet or other reference materials explain how energy is used wisely and list the strategies of conservation of energy. Present your finding to the class.

Strategies of conservation of energy

The steps that you can and should take for saving energy at home or in the office are:

- Switch off lights, fans and other appliances when not in use.
- Water taps should not be left open.

- While cooking vegetables the vessel should remain covered.
- For cooking, only the required quantity of water should be used.
- Soak pulses in water for some time before cooking,
- Use more efficient appliances.
- Use public transport in place of your own vehicle to save fuel.
- Share automobiles rides to office, instead of driving alone to office.

Exercise 7.8

Give short answers.

1. What is conservation of energy?
2. Explain how energy is used wisely.
3. List Strategies of conservation of energy

7.9. Resource Depletion and Environmental Degradation

At the end of this section you will be able to:

• explain resource depletion and environmental degradation.

Activity 7.10

Perform the following tasks in groups and present your conclusion to the class.

i. Explain resource depletion and environmental degradation
ii. Explain causes and effects of resource depletion and environmental degradation

Resource depletion

Resource depletion is the exhaustion of raw materials within a region. Resources are commonly divided between renewable resources and non-renewable resources. Use of either of these forms of resources beyond their rate of replacement is considered to be resource depletion. There are different types of resource depletion. These are deforestation, mining, aquifer depletion, contamination of resources, slash-and-burn agriculture and overconsumption.

Causes for Resource Depletion	Effects of resource depletion	Solutions to the Resource Depletion Problem
• Waste • Farming • Overpopulation • Mining • Erosion • Pollution • Deforestation • Industrialization	• Water shortages • Oil shortages • Economic effects • Health effects • Air pollution • Loss of forests • Global warming • Extinction of animals and plants	• Avoid plastic • Reduce waste • Stop deforestation • Reduction in consumption • Save electricity • Recycle and reuse • Education

Environmental Degradation

The environmental degradation is the deterioration of the environment through depletion of resources. It includes all the biotic and abiotic element that form our surrounding that is air, water, soil, pant animals, and all other living and non-living element of the earth. The major factors of environmental degradation are:

- **human:** (modern urbanization, industrialization, overpopulation growth, deforestation, etc.);
- **natural:** (flood, typhoons, droughts, rising temperatures, fires, etc.
- Causes environmental pollution refers to the degradation of the quality and quantity of natural resources.

The major Effects of Environmental Degradation are: impact on human health, poverty, atmospheric changes, loss of biodiversity and scarcity of natural resources.

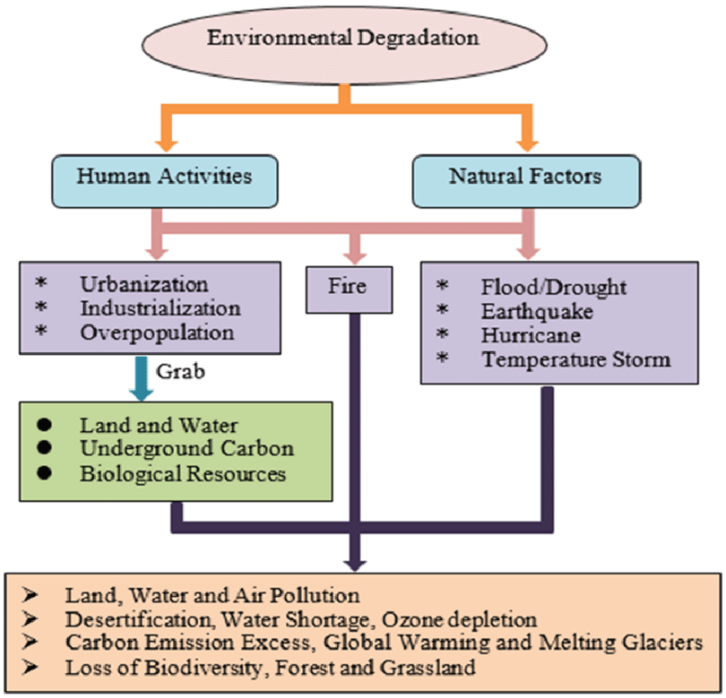


Figure7.14 Different causes of environmental degradation.

Key terms

- Motion
- Rectilinear motion
- Transducer
- Oscillatory motion
- Curvilinear motion
- environmental degradation
- Energy
- Resource depletion
- Force
- Gravity
- matter
- Gravitational force
- Weight
- Newton meter
- conservation of energy
- Spring balance
- Graduated scale
- Renewable energy
- Non-renewable energy
- Rotary motion

SUMMARY

In this unit you learnt that:

- Motion is a continuous change of position relative to a reference point. There are four types of motions. They are rectilinear, curvi linear, rotary and vibrational motion.
- A force is a push or pull upon an object resulting from the object's interaction with another object.
- Gravitational force is the force that is exerted by the Earth on every object, which is near or on its surface.
- Force acting on an object causes the object to change its shape or size, to start moving a stationary object, to stop a moving object, to speed up or to slow down a moving object.
- You can use a device called a force meter to measure the size of a force. It contains a spring connected to a metal hook. The spring stretches when a force is applied to the hook. The bigger the force applied, the longer the spring stretches and the bigger the reading. The unit of force is called the newton, and it has the symbol N.
- Energy is the property of mater, and it comes in many forms, such as heat, sound, light, and motion. It can be transferred between objects, and converted in form.
- Energy exists in different forms. Examples of these are: light energy, heat energy, mechanical energy, gravitational energy, electrical energy, sound energy, chemical energy, nuclear or atomic energy and so on. Each form can be converted or changed into the other forms.
- Sources of energy can be classified into: Renewable Sources and Non-renewable Sources.
- A renewable source is the natural resources that can be naturally replenished and are safe to the environment . Example: Solar energy, geothermal energy, Wind energy, biomass, Hydropower and tidal energy.
- Non-renewable sources of energy cause impact on nature and are a limited supply source. Non-renewable sources can be extracted from the earth, and will run out as time passes. Example: Natural gas, coal, petroleum, and Nuclear energy .

- Using energy more wisely can reduce air pollution and result in cleaner air. The power plants that supply energy release harmful greenhouse gases into the atmosphere.
- Resource depletion is the exhaustion of raw materials within a region. The depletion of natural resources is a big problem. It has several adverse effects on humanity as well as on the whole environmental system.
- Environmental degradation is the deterioration of the environment through depletion of resources such as quality of air, water and soil; the destruction of ecosystems; habitat destruction; the extinction of wildlife; and pollution.

Review Exercise

part I : Write True if the statement is correct and false if it is incorrect.

1. Energy can be created.
2. Geothermal energy is renewable
3. A force that acts an object falling from a tall building is called gravitational force.
4. A push or a pull of an object in a certain direction is known as motion.

Part II: Complete the following sentences by choosing alternatives from the bracket and fill in the blank space.

Coal, natural gas and oil are all examples of _____ (renewable/non-renewable) energy resources. When they are burned, they release _____ (energy/electricity). Coal, natural gas and oil are also known as _____ (nuclear fuels/fossil fuels). Wind and solar energy are examples of _____ (renewable/non-renewable) energy sources because they _____ (can/cannot) be replaced. _____ has to be applied to change the _____of a _____object. (moving, direction, force)

Part III: Choose the best answer from the given alternatives

1. Which type of motion repeats itself at regular intervals of time?
   A. Circular motion C. Rectilinear motion
   B. Vibratory motion D. none of the above
2. Rotation of the Earth is an example of_____.
   A. Periodic motion C. Circular motion
   B. Rectilinear motion D. Both A and B

3. The best definition of force is .
   A. a push or a pull of an object. C. stored energy
   B. energy in motion D. anything that takes up space
4. What type of energy does an oven produce?
   A. light B. heat C. light and heat D. none of the above
5. Which of the following is NOT the effect of force?
   A. Deformation C. To start motion
   B. To stop motion D. to change mass
6. Which instrument is used to measure force?
   A. Spring balance C. Beam balance
   B. Newton D. Thermometer
7. All of the following are examples of things that produce light energy EXCEPT:
   A. Candle B. Compass C. Desk Lamp D. Flashlight
8. Which of the following does NOT produce light energy?
   A. Car B. phone C. picture frame D. computer
9. What energy do you get from eating an apple?
   A. nuclear energy C. sound energy
   B. electromagnetic energy D. chemical energy
10. The energy source that does the least harm to the environment is
   A. Renewable B. Non-renewable
11. Nuclear energy is:
   A. Renewable B. Non-renewable
12. Which of the following is a source of energy?
   A. Sun B. Waves C. Wind D. All
13. Which of these are renewable energy sources?
   A. Coal B. natural gas C. sunlight D. None of these
14. Which type of renewable energy uses the movement of air to generate electricity?
   A. water B. sun C. wind D. Biomass
15. Most energy sources are used to give us
   A. Food and water C. Electricity and heat
   B. Heat and water D. Heat and fuel

16. Renewable energy is energy that:
    A. Can be reused over and over.
    B. Can be used up completely.
    C. Can be changed into a new energy
    D. Cannot be replaced.
17. Non-renewable energy is the energy that
    A. Can be used over and over.
    B. Can be used up completely.
    C. Cannot be changed into a new energy.
    D. Can be replaced
18. Resource depletion is commonly associated with
    A. Water usage C. trees and fishing
    B. Fossil fuel consumption D. All
19. Which one of the following is an examples of environmental degradation:
    A. Deforestation. C. Water Pollution.
    B. Depletion of the ozone layer D. All

Part IV: Give short answer questions

1. Explain and give one example for gravitational force.
2. Why does an object thrown up wards come down after reaching a point?
3. What is meant by weight?
4. Name the motion possessed by these objects- blades of an electric fan in motion, a spinning top, and hands of a clock, vehicle on a straight road, the earth around the sun and pendulum of a wall clock.

Addis Ababa City Administration Education Bureau